

The hunt for formamide in interstellar ices

A toolkit of laboratory infrared spectra in astronomically relevant ice mixtures and comparisons to ISO, *Spitzer*, and JWST observations

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Received 24 May 2023 / Accepted 30 June 2023

ABSTRACT

Context. Although solid-state pathways are expected to dominate the formation mechanisms of many complex organic molecules (COMs), very few COMs have been securely identified in interstellar ices, in stark contrast with the many COM detections in the gas phase. The launch of the *James Webb* Space Telescope (JWST) and its increase in sensitivity and spectral resolution opens the possibility of identifying more COMs in ices, but additional laboratory data are necessary. Formamide (NH₂CHO) is one such COM that is of great interstellar and prebiotic relevance where more laboratory data are needed in the hunt for its presence in interstellar ices.

Aims. This work aims to characterize the mid-IR spectra of formamide in its pure form as well as in mixtures of the most abundant interstellar ices via laboratory simulation of such ices, as well as to demonstrate how these laboratory spectra can be used to search for formamide in ice observations.

Methods. Mid-IR spectra (4000–500 cm⁻¹/2.5–20 μm) of formamide, both in its pure form as well as in binary and tertiary mixtures with H₂O, CO₂, CO, NH₃, CH₃OH, H₂O:CO₂, H₂O:NH₃, CO:NH₃, and CO:CH₃OH, were collected at temperatures ranging from 15–212 K.

Results. Apparent band strengths and positions of eight IR bands of pure amorphous and crystalline formamide at various temperatures are provided. Three of these bands are identified as potential formamide tracers in observational ice spectra: the overlapping C = O stretch and NH₂ scissor bands at 1700.3 and 1630.4 cm⁻¹ (5.881 and 6.133 μm), the CH bend at 1388.1 cm⁻¹ (7.204 μm), and the CN stretch at 1328.1 cm⁻¹ (7.529 μm). The relative apparent band strengths, positions, and full width half maxima (FWHM) of these features in mixtures at various temperatures were also determined. All of the laboratory spectra are available to the community on the Leiden Ice Database for Astrochemistry (LIDA) for use in the interpretation of both observations (e.g., from JWST) and laboratory spectroscopic data. Finally, the laboratory spectra are compared to observational spectra of a variety of low- and high-mass young stellar objects as well as prestellar cores observed with the Infrared Space Observatory, the *Spitzer* Space Telescope, and JWST. A comparison between the formamide CH bend in laboratory data and the 7.24 μm band in the observations tentatively indicates that, if formamide ice is contributing significantly to the observed absorption, it is more likely in a polar matrix. Upper limits ranging from 0.35–5.1% with respect to H₂O were calculated via scaling the formamide:H₂O laboratory spectrum to the observations. These upper limits are in agreement with gas-phase formamide abundances and take into account the effect of a H₂O matrix on formamide's band strengths.

Key words. astrochemistry – methods: laboratory: molecular – methods: laboratory: solid state – techniques: spectroscopic – ISM: molecules – infrared: ISM

1. Introduction

Of the >280 molecules that have been detected in interstellar environments (Endres et al. 2016), formamide (NH₂CHO, Fig. 1) has become one of the most widely and deeply investigated in observational, modeling, computational, and laboratory studies in the last decade. Containing all four of the most abundant biological elements (C, H, N, and O), formamide is the simplest molecule that contains the biologically essential amide bond and has been suggested as a plausible prebiotic precursor to various nucleobases (e.g., Saladino et al. 2003; Barks et al. 2010), the chemical building blocks of RNA and DNA. It has also been proposed as an alternative prebiotic solvent to promote condensation reactions, which form many vital biological molecules but are highly endergonic in purely aqueous solutions (e.g.,

phosphorylation), by lowering water activity (Gull et al. 2017; Pasek 2019; Lago et al. 2020).

Given this potential prebiotic relevance, the fact that formamide has been observed in numerous sources in the interstellar medium as well as on extraterrestrial bodies in our own Solar System has exciting implications for astrobiology. First detected in the interstellar medium in the gas phase by Rubin et al. (1971) in the Sagittarius B2 high-mass star-forming region, formamide has since been observed in over 30 massive young stellar objects (MYSOs) as well as low-mass YSOs (LYSOs) with hot corinos and protostellar shocks (López-Sepulcre et al. 2019 and references therein). Within our Solar System, gas-phase formamide has been found in the comae of the comets Lemmon, Lovejoy, and Hale-Bopp, with abundances ranging around 0.01–0.02% with respect to H₂O (Bockelée-Morvan et al. 2000;

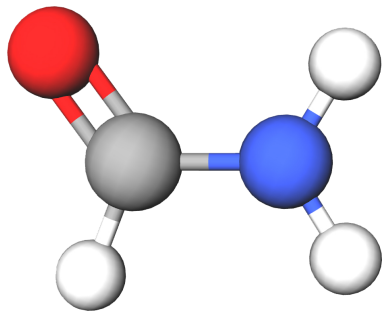


Fig. 1. Molecular structure of NH_2CHO (hydrogen, white; carbon, gray; nitrogen, blue; and oxygen, red).

Biver et al. 2014). It was also detected in situ by the Rosetta mission on comet 67P Churyumov-Gerasimenko, both on the surface by the Cometary Sampling and Composition experiment (COSAC) instrument on the Philae lander (Goesmann et al. 2015) and in the coma by the Double Focusing Mass Spectrometer (DFMS) on the Rosetta Orbiter Spectrometer for Ion and Neutral Analysis (ROSINA) instrument (Altwegg et al. 2017; Rubin et al. 2019), where the formamide abundance was found to be $\sim 0.004\%$ with respect to H_2O .

Notably, all of the interstellar sources in which gas-phase formamide has been securely detected have hot cores and corinos or shocked regions, where temperatures are high enough for formamide to thermally desorb from icy grains into the gas phase (López-Sepulcre et al. 2019). Additionally, in many of these sources, the formamide abundance correlates almost linearly with the abundance of isocyanic acid (HNCO ; Bisschop et al. 2007b; Mendoza et al. 2014; López-Sepulcre et al. 2015), and, in the case of the low-mass source IRAS 16293-2422, the two species are spatially correlated and have very similar deuteration ratios (Coutens et al. 2016).

These aspects of formamide observations could be considered evidence that formamide is formed in the solid state (i.e., via ice chemistry), possibly in a pathway chemically related to HNCO , and it is detected in the gas phase following desorption from icy grains. The ice formation and grain sublimation scenario is further supported by recent observational work investigating excitation temperatures of N-bearing complex organic molecules (COMs) in 37 MYSOs from the ALMA Evolutionary study of High Mass Protocluster Formation in the Galaxy (ALMAGAL) survey, where formamide had the highest excitation temperatures of all the studied N-bearing COMs (≥ 250 K; Nazari et al. 2022). These temperatures are consistent with thermal desorption experiments, in which formamide ice sublimates at high temperatures (typically >210 K) even when it is mixed with or deposited on top of more volatile species such as H_2O and CO , and at even higher temperatures (>250 K) when the experiments are performed on certain dust grain analog substrates (Dawley et al. 2014; Urso et al. 2017; Chaabouni et al. 2018; Corazzi et al. 2020).

Experimentally, solid-state formamide has been identified as a product of processing via a variety of energetic sources (e.g., electron, UV, X-ray, and ion irradiation) of a myriad of simple ice mixtures, including (but not limited to) $\text{CO}:\text{NH}_3$ (Demyk et al. 1998; Hudson & Moore 2000; Jones et al. 2011; Bredehöft et al. 2017; Martín-Doménech et al. 2020) and $\text{H}_2\text{O}:\text{CO}:\text{NH}_3$ (Hudson & Moore 2000; Ciaravella et al. 2019; Chuang et al. 2022), $\text{NO}:\text{H}_2\text{CO}:\text{H}$ and $\text{NO}:\text{CH}_3\text{OH}:\text{H}$ (Fedoseev et al. 2016), $\text{H}_2\text{O}:\text{HCN}$ (Gerakines et al. 2004), $\text{H}_2\text{O}:\text{CH}_4:\text{NH}_3$ and $\text{H}_2\text{O}:\text{CH}_4:\text{N}_2$ (Kaňuchová et al. 2016), and HNCO (Raunier

et al. 2004) and $\text{CH}_4:\text{HNCO}$ (Ligterink et al. 2018). Evidently, energetic processing of almost any ice mixture that contains H, N, C, and O is very likely to produce formamide. Such processing experiments mimic the radiation environments experienced by ices in protostellar envelopes and protoplanetary disks. Furthermore, recent experiments by Dulieu et al. (2019) demonstrate that hydrogenation of $\text{NO}:\text{H}_2\text{CO}$ can also produce formamide, providing a plausible nonenergetic formation pathway that is relevant to cold, dark clouds.

While a plethora of observational, experimental, and theoretical works (see Sect. 2) have significantly progressed our understanding of formamide's interstellar presence and its plausible chemical history, whether its formation occurs in the solid state, gas phase, or both remains unclear. A secure detection of formamide in ices would be immensely valuable to resolve this debate regarding its formation mechanism. Such a detection, if well resolved, could provide parameters such as formamide's solid-state abundance and its physico-chemical environment, which are essential to elucidating its formation pathway.

Previously, formamide has been tentatively detected in the solid state in the Infrared Space Observatory Short Wavelength Spectrometer (ISO-SWS) spectra of the MYSOs W33A and NGC 7538 IRS 9. In the case of W33A, an upper limit of 2.1% with respect to H_2O via the CH bend at $7.22 \mu\text{m}/1385 \text{ cm}^{-1}$ was derived, but the authors noted that the peak position in the observation ($7.24 \mu\text{m}$) was red-shifted relative to the formamide peak in their laboratory spectra (Schutte et al. 1999). For NGC 7538 IRS 9, no upper limit of formamide was provided – a laboratory spectrum of irradiated HNCO that showed IR evidence of formamide formation was qualitatively evaluated as a spectral fit to the observed $6 \mu\text{m}/1700 \text{ cm}^{-1}$ band (Raunier et al. 2004). In both of these cases, the bands attributed to formamide were overlaid on top of or blended with other strong ice features.

Typically, reference laboratory IR spectra are used to assign and fit astronomically observed IR features to specific species, and band strengths acquired via systematic laboratory experiments are used to quantify the column densities of these species. For COMs such as formamide that are expected to be present in the ice in very low concentrations ($\leq 5\%$), it is important to obtain these spectra and band strengths not only for pure ices, but also in chemical conditions that are more realistic for interstellar ices. Namely, the molecule of interest should be diluted in the more abundant simple ice species (e.g., H_2O , CO , and CO_2), as interactions with other species present in the ice matrix can significantly alter the positions, profiles, and apparent band strengths of a molecule's vibrational features. Morphological changes in the ice caused by thermal processing, such as transitions from amorphous to crystalline ice or matrix segregation, can also dramatically change an ice's spectral features, so spectra should be collected at a variety of temperatures as well. Considering such factors is not only important to accurately assign and quantify the molecule of interest, but it can also provide valuable information about the molecule's physico-chemical environment and history.

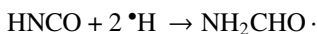
In previous IR characterization work, Brucato et al. (2006) derived the refractive index, density, and several band strengths of pure formamide, but integration ranges and errors were not provided for these band strengths, and no spectra of heated formamide or formamide in mixtures were collected. In order to tentatively assign the $7.24 \mu\text{m}$ band in W33A's spectrum to formamide, Schutte et al. (1999) collected spectra of formamide at 10 K in H_2O and $\text{H}_2\text{O}:\text{CH}_3\text{OH}$ matrices, but only one band was characterized from these spectra, and it is unclear for what phase of formamide the band strength used in the upper limit

calculation was derived. [Urso et al. \(2017\)](#) collected IR spectra of formamide in pure, H₂O-dominated, and CO-dominated ice matrices, but the band strengths, peak positions, and full width half maxima (FWHMs) of the formamide features in these mixtures are not given. [Sivaraman et al. \(2013\)](#) presented the peak positions of the bands of pure formamide in the 30–210 K temperature range, but no spectra of formamide in mixtures were collected.

Thus, in an effort to enable more secure assignments and accurate abundance and/or upper limit determinations of formamide in observed ice spectra, this work provides a comprehensive set of laboratory transmission IR spectra of pure formamide as well as formamide diluted in nine different astrophysically relevant ice mixtures of varying polarities. These spectra are provided at temperatures ranging from 15 to 212 K. Apparent band strengths were derived for eight integrated regions from the pure formamide spectra, and from these, three bands are evaluated as the most promising for future identification of formamide in observations. These bands are also fully characterized (i.e., peak positions, FWHMs, and relative band strengths are provided). Examples of how these spectra and values can be used in future analyses of ice observations are described, and new upper limits of formamide in a variety of objects (prestellar cores, low-mass protostars, and high-mass protostars) were calculated. Finally, all spectra are made publicly available on the Leiden Ice Database¹ ([Rocha et al. 2022](#)) for the community to use in fitting to their ice observations. This work is particularly timely given the recent launch of the *James Webb* Space Telescope (JWST), which may enable the detection of new COMs in interstellar ices due to its unprecedented sensitivity and spectral resolution.

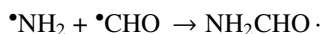
2. Formamide formation mechanism debate

A variety of pathways have been suggested to explain the observed solid-state formamide formation in laboratory ice experiments. One initially proposed mechanism was the hydrogenation of HNCO, an attractive premise given that it provided a direct chemical link between HNCO and formamide to explain their correlation in gas-phase observations:



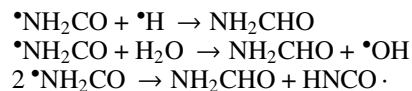
This pathway was first suggested by [Charnley \(1997\)](#) and was stated as a possible formation mechanism of formamide when it was observed in VUV irradiation experiments of pure HNCO ([Raunier et al. 2004](#)). However, hydrogenation experiments by [Noble et al. \(2015\)](#) via H bombardment of HNCO < 20 K did not produce detectable amounts of formamide, although the authors suggested that the reaction may be prevented in their experiments by the formation of very stable HNCO dimers or polymers, and that it could possibly proceed if HNCO is diluted in the matrix of an ice like H₂O. Indeed, subsequent experiments by [Haupa et al. \(2019\)](#) showed that, in a 3.3 K para-H₂ matrix, formamide can form from HNCO via a hydrogen addition-abstraction cycling mechanism, but in this reaction scheme, HNCO is still the favored product.

Another proposed formation pathway is the following radical-radical recombination:



This mechanism is technically barrierless and can proceed at low temperatures (~10 K) but produces higher yields at higher temperatures (~20–40 K) due to increased mobility allowing the radicals to orient in the proper reaction geometry ([Rimola et al. 2018](#); [Martín-Doménech et al. 2020](#)). In the laboratory, this mechanism requires some form of energetic processing to generate the NH₂ and CHO radicals, and its viability is supported by the presence of the CHO radical in the experimental spectra ([Jones et al. 2011](#); [Fedoseev et al. 2016](#); [Ciaravella et al. 2019](#); [Martín-Doménech et al. 2020](#); [Chuang et al. 2022](#)).

Various mechanisms have also been suggested where formamide is produced from the NH₂CO radical, which could form by the radical-molecule association of NH₂ and CO or CN and H₂O ([Hudson & Moore 2000](#); [Bredehöft et al. 2017](#); [Rimola et al. 2018](#)):

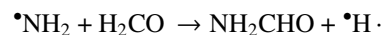


However, the formation of the NH₂CO radical via a pathway that does not involve hydrogen abstraction from already existing formamide, as seen in [Haupa et al. \(2019\)](#), has yet to be experimentally confirmed.

While these latter mechanisms do not provide an immediately obvious direct solid-state link between HNCO and NH₂CHO, some experimental studies have suggested alternative links consistent with these mechanisms. For example, once formed, formamide can decompose into HNCO via dehydrogenation and photolysis by H₂ loss ([Brucato et al. 2006](#); [Haupa et al. 2019](#); [Chuang et al. 2022](#)), so HNCO may be a product of NH₂CHO rather than the other way around. [Fedoseev et al. \(2016\)](#) proposed that the NH₂ radical can produce either HNCO or NH₂CHO depending on the degree of hydrogenation of the C- and O-containing molecule with which it reacts: the reaction of NH₂ with CO leads to HNCO, while NH₂ with HCO or H₂CO leads to formamide.

Thus, while formamide may not be a direct product of HNCO, the two species may be linked in a solid-state chemical network by common precursors. Astrochemical models using the rate constants from [Fedoseev et al. \(2016\)](#) further corroborate that, indeed, a direct chemical link between HNCO and NH₂CHO is not necessary to reproduce the observed linear correlation between them in models of various interstellar environments and suggest instead that their correlation could be explained by their similar responses to physical (i.e., thermal) environments ([Quénard et al. 2018](#)).

In addition to these solid-state mechanisms, the plausibility of the following gas-phase formation route has been extensively debated in computational and modeling works since its proposal in [Garrod et al. \(2008\)](#):



According to its first published electronic structure and kinetic calculations, this reaction is essentially barrierless at low temperatures and thus should proceed readily in interstellar environments ([Barone et al. 2015](#); [Vazart et al. 2016](#)). Furthermore, chemical models of the protostar IRAS 16293-2422 and the molecular shocks L1157-B1 and B2 utilizing the calculated rate coefficients of this reaction produce formamide abundances that are consistent with observed values ([Barone et al. 2015](#); [Codella et al. 2017](#)), and follow-up studies calculating rate coefficients of deuterated formamide formation via the same

¹ <https://icedb.strw.leidenuniv.nl>

reaction show that formamide's observed deuteration ratio does not necessarily exclude the possibility of gas-phase formation (Skouteris et al. 2017).

However, the accuracy of these calculated rate coefficients has been called into question given that they neglect the zero point energy (ZPE) of one of the transition states. When the ZPE of the transition state is included, the reaction barrier becomes large enough that the reaction rate is negligible at low temperatures (Song & Kästner 2016), although some argue that inclusion of the ZPE is not warranted for this transition state and results in overestimation of the reaction barrier (Skouteris et al. 2017). Recent gas-phase experiments attempting to perform this route did not confirm any formamide formation, and their detection upper limits are consistent with the reaction barrier that includes the transition state ZPE (Douglas et al. 2022).

3. Methodology

All of the measurements were collected in the Laboratory for Astrophysics at Leiden Observatory on the IRASIS (InfraRed Absorption Setup for Ice Spectroscopy) chamber. The setup was described in detail in Rachid et al. (2021) and Rachid et al. (2022), and it has since undergone several upgrades, including a decrease of its base pressure to $<1.0 \times 10^{-9}$ mbar by the addition of new pumps, an exchange of the laser used for interference measurements to one with a wavelength of 543 nm (as the formamide ice refractive index was measured by Brucato et al. 2006 at this wavelength), and the implementation of an independent tri-dosing leak valve system that can be calibrated with a quadrupole mass spectrometer (QMS) following the procedure described in Appendix C.

The optical layout of the chamber remains the same as that shown in Fig. 1 in Rachid et al. (2021): a Ge substrate sits at the center of the chamber and is cooled by a closed-cycle He cryostat to 15 K. Ices are grown on the substrate via background deposition of gases and vapors dosed into the chamber through leak valves. Infrared transmission spectra are collected through two ZnSe viewports that are parallel to the Ge substrate and normal to the IR light beam. During deposition, laser interference patterns used to determine ice thickness are measured on both sides of the Ge substrate (which is opaque and reflective in the visible light range) via photodiode detectors placed outside of viewports positioned 45° from the substrate normal. The patterns obtained from each side of the substrate during deposition show equal deposition rates on both sides. After deposition, the substrate can be heated to obtain IR spectra at different temperatures. In this work, 256 spectral scans with a 0.5 cm^{-1} resolution were collected and averaged while the substrate was heated at a rate of 25 K h^{-1} , resulting in a temperature uncertainty of $\pm 1.5 \text{ K}$ in each heated spectrum. Spectra were collected during heating until reaching the temperature at which the major matrix component desorbed. Before their analysis, all spectra were baseline-corrected using a cubic spline function.

The liquids and gases used in this work were formamide (Sigma Aldrich, $\geq 99.5\%$), water (Milli-Q, Type I), carbon dioxide (Linde, $\geq 99.995\%$), carbon monoxide (Linde, $\geq 99.997\%$), ammonia (PraxAir, $\geq 99.96\%$), and methanol (Sigma Aldrich, $\geq 99.9\%$). The mixing ratios calculated for all of the spectra via the method outlined in Appendix C are presented in Table 1. Uncertainties in the column densities used to calculate these ratios are estimated to be $\sim 21\%$ for the formamide column densities and $\sim 27\%$ for the matrix species column densities (see Appendix C). Prior to deposition, the liquid formamide

Table 1. Experimentally determined ratios of all mixtures presented in this work.

Mixture	Ratio
Pure NH_2CHO	–
$\text{NH}_2\text{CHO}:\text{H}_2\text{O}$	7:100
$\text{NH}_2\text{CHO}:\text{CO}_2$	7:100
$\text{NH}_2\text{CHO}:\text{CO}$	4:100
$\text{NH}_2\text{CHO}:\text{CH}_3\text{OH}$	8:100
$\text{NH}_2\text{CHO}:\text{NH}_3$	7:100
$\text{NH}_2\text{CHO}:\text{H}_2\text{O}:\text{CO}_2$	8:100:37
$\text{NH}_2\text{CHO}:\text{H}_2\text{O}:\text{NH}_3$	8:100:31
$\text{NH}_2\text{CHO}:\text{CO}:\text{CH}_3\text{OH}$	4:100:12
$\text{NH}_2\text{CHO}:\text{CO}:\text{NH}_3$	5:100:16

sample was heated to 60°C and pumped on directly with a turbomolecular pump in order to remove contaminants (primarily water).

The apparent band strengths of pure formamide are determined via depositing formamide onto the substrate held at 15 K while simultaneously collecting the transmission IR spectra and the laser interference pattern. The thickness d of the ice can be derived from the laser interference pattern via the following equation:

$$d = \frac{m\lambda}{2\sqrt{n^2 - \sin^2\theta}}, \quad (1)$$

where m is an integer number of constructive fringes, λ is the laser wavelength, n is the ice refractive index (1.361 for formamide at 543 nm, from Brucato et al. 2006), and θ is the angle of incidence.

Enough formamide is deposited so that four constructive fringes are acquired, the thickness of the ice at each fringe peak is calculated, and the integrated absorbances of eight spectral regions (see Table 2) are calculated from the spectra collected at the time that a fringe peak was reached. Then, the integrated absorbance for each spectral region is plotted as a function of ice thickness, and the slope of this line, $\Delta \int \text{abs}(\nu) d\nu / \Delta d$, is obtained via a least-squares fit. From this value, the apparent band strengths A' can be approximated with an equation based on the Beer-Lambert Law (e.g., Hudson et al. 2014; Gerakines et al. 2023):

$$A' = \frac{2.303 M}{\rho N_A} \times \frac{\Delta \int \text{abs}(\nu) d\nu}{\Delta d}, \quad (2)$$

where M is the molar mass of formamide ($45.041 \text{ g mol}^{-1}$), ρ is the density of formamide ice (0.937 g cm^{-3} , from Brucato et al. 2006), and N_A is Avogadro's number. Using change in integrated absorbance over change in thickness in this equation rather than the absolute values of both variables ensures that there is no contribution of any residue from previous experiments on the substrate to the calculated ice thickness. It also does not require a constant ice growth rate.

The apparent band strengths reported in Table 2 are the averages of three repeated measurements following this method. The experimental uncertainties derived from the standard deviation of these three measurements range from 3 to 8% for the eight band strengths. However, simply using the standard deviations from the repeated measurements as the band strengths

Table 2. Apparent IR band strengths (and their peak positions, assignments, and corresponding integrated regions) of the IR bands of amorphous NH₂CHO ice at 15 K derived in this work.

Peak positions (cm ⁻¹ μm)	Integrated region (cm ⁻¹)	Assignment ^(*)	A' (10 ⁻¹⁷ cm molec ⁻¹)	A' ^(**)
3371.2 2.966	3719–2980	NH ₂ antisymm. stretch (ν _{as} NH ₂)	15.6 ± 2	13.49
3176.4 3.148		NH ₂ symm. stretch (ν _s NH ₂)		
2881.9 3.470	2980–2842	CH stretch (ν CH)	0.56 ± 0.08	0.47
2797.7 3.574	2842–2600	CH bend overtone (2δ CH)	0.23 ± 0.03	–
1700.3 5.881	2050–1532	C = O stretch (ν CO)	6.4 ± 0.9 ^(***)	6.54
1630.4 6.133		NH ₂ scissor (δ NH ₂)		
1388.1 7.204	1532–1364	CH bend (δ CH)	0.9 ± 0.1 ^(***)	0.68
1328.1 7.529	1364–1258	CN stretch (ν CN)	0.8 ± 0.1 ^(***)	0.85
1108.1 9.024	1174–1030	NH ₂ rock (ρ NH ₂)	0.23 ± 0.03	–
1056.1 9.469		CH out-of-plane def. (γ CH)		
689.2 14.510	1030–575	NH ₂ wag (ω NH ₂)	2.3 ± 0.4	–
634.0 15.773		NH ₂ twist (2τ NH ₂)		

Notes. ^(*)Band assignments made according to the works of Räsänen (1983); Kwiatkowski & Leszczyński (1993); Torrie & Brown (1994); Lundell et al. (1998). Some band strengths contain multiple absorption modes in their corresponding integrated regions. ^(**)Values from Brucato et al. (2006). ^(***)Bands that are also characterized in mixtures in this work.

uncertainties neglects potential systemic sources of error such as uncertainties in the laser alignment geometry and the data analysis procedure. Thus, the uncertainties provided in Table 2 are calculated via error propagation of all of the experimental terms in Eq. (2), using the same estimated uncertainties as Rachid et al. (2022) for the ice thickness (4%) and integrated absorbance (10%) as well as the ice density (10%). This calculation yields an uncertainty of 15% for the reported band strength values.

From the pure formamide apparent band strengths, the apparent band strengths of formamide in the investigated mixtures, A'_i , are calculated using the formamide column densities N_{mix} (obtained from the methods described in Appendix C) via the following equation:

$$A'_i = 2.303 \times \frac{\int \text{abs}(\nu) d\nu}{N_{\text{mix}}}, \quad (3)$$

and the relative apparent band strengths, η , are subsequently found by:

$$\eta = \frac{A'_i}{A'}. \quad (4)$$

Following propagation of error from the pure apparent band strengths, integrated absorbances, and the formamide column densities in the mixtures (see Appendix C), the uncertainties of the relative apparent band strengths presented here are estimated to be ~28%.

4. Results

The spectra of pure amorphous and crystalline formamide are presented in Fig. 2, and the eight apparent band strengths calculated at 15 K are presented in Table 2. Peak positions and vibrational mode assignments are also provided. Some integrated regions contain multiple overlapping peaks; in these cases, the peak positions and assignments were provided for all peaks within the integrated region, but the peaks were not deconvolved to give an individual band strength for each peak. These band strengths have percent differences ranging from 1 to 35% compared to those given for the same peak values in

Brucato et al. (2006). As integration bounds were not provided by Brucato et al. (2006), any discrepancies in band strengths may be caused by differences in chosen integration regions.

The transition from amorphous to crystalline formamide is observed at 170 K, indicated by its bands becoming sharper and narrower and some peaks splitting. The amorphous nature of almost all of the pure and mixed ices collected at 15 K can be ascertained from their spectra, which have typical amorphous features that show evidence of matrix crystallization during the warm-up phase of the experiments. This excludes the mixtures containing CO, whose phase at 15 K in these experiments may be crystalline given recent investigations of CO ice structure ≥10 K (He et al. 2021; Gerakines et al. 2023; Rachid et al., in prep.).

Figure 3 presents the spectrum of pure formamide ice along with the spectra of the pure matrix components, all at 15 K. The formamide peaks indicated in the shaded areas were selected for full characterization (i.e., their peak positions, FWHMs, and relative band strengths are determined for mixtures): the overlapping C = O stretch and NH₂ scissor at 1700.3 cm⁻¹/5.881 μm and 1630.4 cm⁻¹/6.133 μm, respectively, and the slightly overlapping CH bend and CN stretch at 1388.2 cm⁻¹/7.204 μm and 1328.1 cm⁻¹/7.529 μm, respectively. These peaks were selected because they are strong, have sharp profiles, and overlap the least with the major peaks of the most common interstellar ices, making them the best candidates for identifying formamide in interstellar ice spectra. There is still some overlap between these formamide peaks and some minor peaks of the matrix components, namely the water OH bend at ~1600 cm⁻¹/6.25 μm, the methanol CH₃ and OH bends at ~1460 cm⁻¹/6.85 μm, and the ammonia NH scissoring at 1624 cm⁻¹/6.16 μm. However, with sufficiently high formamide concentrations, it may still be possible to identify formamide in these spectral regions, as these matrix bands are relatively weak and broad.

The matrix- and temperature-dependent changes in these selected formamide ice bands are discussed in the following subsections, and their peak positions, FWHMs, and relative band strengths in different mixtures at various temperatures are reported in Appendices A and B. The NH₂ stretching features at 3371.2 cm⁻¹/2.966 μm and 3176.4 cm⁻¹/3.148 μm and the NH₂ wagging and twisting features at 689.2 cm⁻¹/14.510 μm

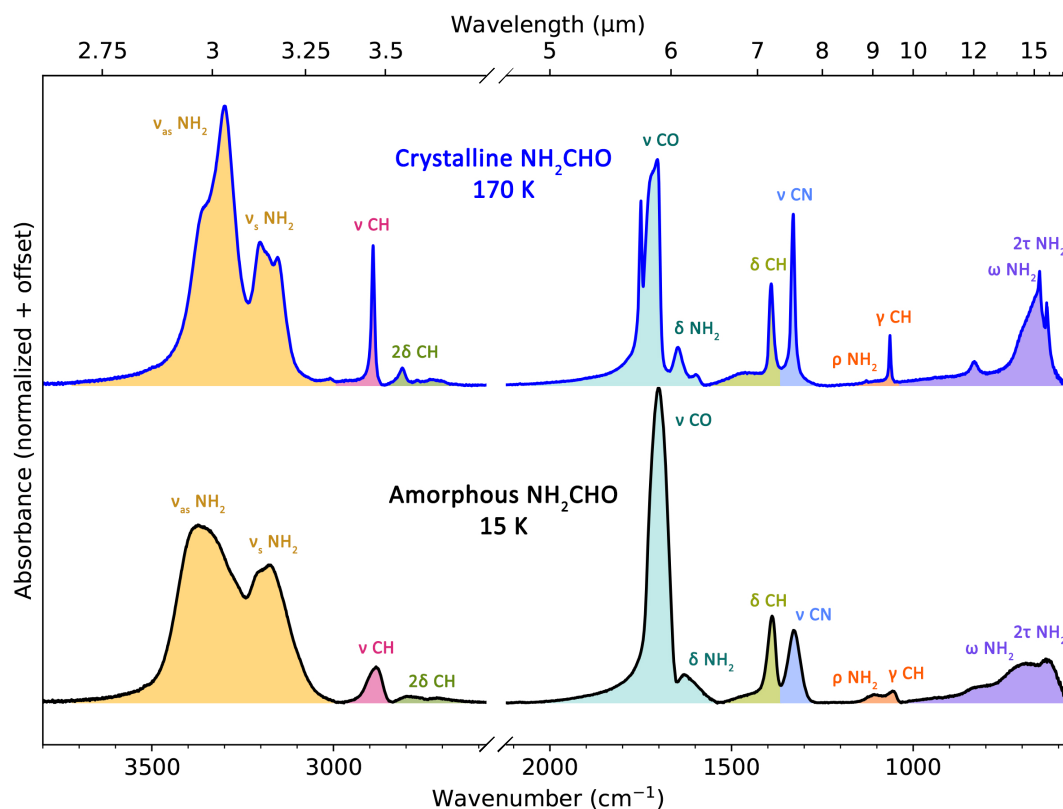


Fig. 2. Infrared spectra in the range 3800–575 cm^{-1} of both amorphous and crystalline formamide (NH_2CHO). The integrated regions used to calculate the band strengths in Table 2 are indicated via the shaded areas under the peaks.

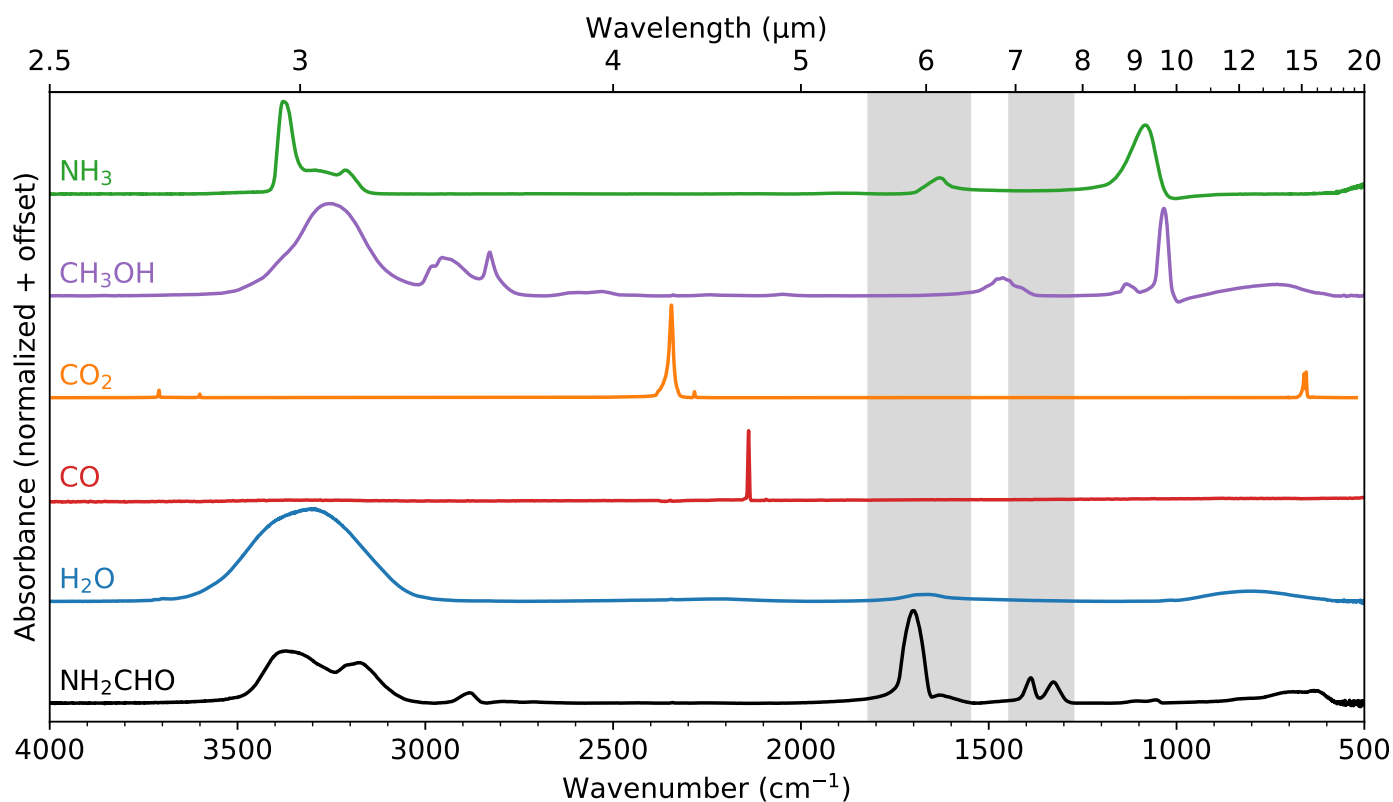


Fig. 3. Infrared spectra of pure formamide (NH_2CHO) and five molecules observed in high abundances in interstellar ices that are present in the mixtures analyzed in this work (H_2O , CO , CO_2 , CH_3OH , NH_3), all at 15 K. Peaks selected for further characterization in mixtures are indicated in the shaded areas.

and $634.0\text{ cm}^{-1}/15.773\text{ }\mu\text{m}$ were excluded from further characterization despite their relatively large band strengths due to their direct overlap with the two most intense water features, the OH stretch at $\sim 3279\text{ cm}^{-1}/3.05\text{ }\mu\text{m}$ and the H₂O libration at $\sim 780\text{ cm}^{-1}/12.8\text{ }\mu\text{m}$, respectively (Öberg et al. 2007). The remaining formamide bands, the CH stretch at $2881.9\text{ cm}^{-1}/3.470\text{ }\mu\text{m}$, the CH bend overtone at $2797.7\text{ cm}^{-1}/3.574\text{ }\mu\text{m}$, and the convolved NH₂ rock at $1108.1\text{ cm}^{-1}/9.024\text{ }\mu\text{m}$ and CH out-of-plane deformation at $1056.1\text{ cm}^{-1}/9.469\text{ }\mu\text{m}$, have low band strengths and directly overlap with various methanol features: the CH₃ stretches at $2950\text{ cm}^{-1}/3.389\text{ }\mu\text{m}$ and $2830\text{ cm}^{-1}/3.533\text{ }\mu\text{m}$, the CH₃ rock at $1126\text{ cm}^{-1}/8.881\text{ }\mu\text{m}$, and the C-O stretch at $1027\text{ cm}^{-1}/9.737\text{ }\mu\text{m}$ (Luna et al. 2018).

4.1. C = O stretching and NH₂ scissoring features (~ 1700 and 1630 cm^{-1})

Figure 4 shows how the profile of the C = O band ($1700.3\text{ cm}^{-1}/5.881\text{ }\mu\text{m}$) changes in different mixtures and temperatures and presents the peak positions and FWHMs of these spectra in a scatter plot. This type of scatter plot can help to narrow down the possible thermochemical environments of molecules identified in observations (see Sect. 5). The right bottom plot in the figure shows the strengths of the band in the different mixtures at 15 K relative to the value of the band strength of pure formamide. The integrated regions used to calculate these band strengths also include the NH₂ scissoring mode, which presents as a weak, broad feature overlapping with the red shoulder of the C = O stretch (see Fig. 4). The FWHM and relative band strengths of the formamide:NH₃ mixture are excluded from the bottom scatter plots in Fig. 4 and the tables in Appendix A due to the significant overlap of this band with ammonia's NH scissoring mode at $1624\text{ cm}^{-1}/6.16\text{ }\mu\text{m}$. The NH₃ peak is small enough in the NH₃-containing tertiary mixtures relative to the formamide C = O stretch to extract reliable peak positions and FWHMs, but relative band strengths were not calculated.

In pure amorphous formamide ($<170\text{ K}$), the C = O stretch appears as a single broad peak centered at $1704.2\text{ cm}^{-1}/5.868\text{ }\mu\text{m}$. Generally, being in a mixture causes the feature to sharpen, most dramatically so in apolar mixtures in which CO or CO₂ are the dominant species. For example, the FWHM of the feature in formamide:CO₂ at 15 K is 51.1 cm^{-1} , over three times narrower than that in pure formamide. Also, in the CO, CO:CH₃OH, and crystalline CO₂ matrices, some peak splitting occurs before the formamide crystallization temperature is reached. Such sharpening and splitting is typical when a polar molecule is highly diluted in an apolar matrix and is caused by the polar molecule being isolated in the matrix as a monomer or dimer, unable to form the hydrogen bonds with other polar molecules that tend to broaden and blend vibrational features (e.g., Ehrenfreund et al. 1996). Urso et al. (2017) also previously observed the formamide peaks splitting due to monomer and dimer formation in their very dilute 1:40 formamide:CO mixture. In the polar mixtures, however, as hydrogen bonding with the matrix is still possible, the feature remains broad. The feature is the most blue-shifted in the binary CO and CO₂ mixtures, where its peak values are 1717.2 and 1703.7 cm^{-1} , respectively, in the 15 K ices, while in polar mixtures it tends to red-shift, with the most red-shifted peak position being that of the tertiary H₂O:CO₂ mixture, 1694.0 cm^{-1} . Despite containing a high fraction of apolar CO, the tertiary mixtures with CO:CH₃OH and CO:NH₃ have peak positions similar to the polar mixtures. The

relative band strength of this formamide feature is >1 in all of the investigated matrices, with no observable trend related to polarity present in these values.

At formamide's crystalline phase transition temperature (170 K), the C = O peak blue-shifts and splits into multiple blended features. This is only observed in the pure formamide spectrum because all of the matrix molecules investigated here desorb below 170 K. An interesting trend to note is that, as the mixtures increase in temperature, the formamide C = O feature tends to broaden to have a FWHM value more similar to that of pure formamide. This trend can be easily identified in the scatter plot in Fig. 4, where the scatter points of several of the mixtures move closer to the points of the pure amorphous spectrum as temperature increases. It is also particularly noticeable in Fig. 4 in the spectra of mixtures containing H₂O, which have peak position and FWHM values at high temperatures ($>150\text{ K}$) that are close to those of the pure spectrum. Sudden broadening of the FWHM to a value closer to that of pure formamide also tends to occur at the matrix crystallization temperatures (for example, in the binary CO₂ mixture between ~ 30 and 40 K and in the H₂O-containing mixtures between ~ 130 and 150 K). These spectral changes indicate that formamide segregation is occurring in the matrix as the ice is heated and is particularly promoted when the ice undergoes a dramatic restructuring during matrix crystallization. The conclusion that solid-phase formamide diluted in a matrix is mobilized via heating is consistent with formamide thermal processing studies, in which formamide deposited on top of water ice diffused through the water during heating (Chaabouni et al. 2018).

4.2. CH bending and CN stretching features (~ 1388 and 1328 cm^{-1})

The shape and position of the CH bend ($1388.1\text{ cm}^{-1}/7.204\text{ }\mu\text{m}$) does not vary much depending on chemical environment or temperature, with peak positions only ranging from 1398.0 to 1387.2 cm^{-1} and FWHM values ranging from 11.1 to 27.5 cm^{-1} in the mixtures investigated here (see Figs. 5 and 6). As in the C = O stretch band, the binary apolar mixtures with CO and CO₂ have the most blue-shifted and narrow peaks; however, a trend of the mixture band shifting during heating to peak position and FWHM values closer to those of the pure band is not as clear. The band strength of the CH bend increases in all of the mixtures (e.g., $\eta = 1.63$ at 15 K in the formamide:H₂O mixture) except for the CO₂ mixture, in which the band strength decreases slightly ($\eta = 0.85$ at 15 K).

The CN stretching band ($1328.1\text{ cm}^{-1}/7.529\text{ }\mu\text{m}$) varies much more dramatically across different mixtures and temperatures (see Figs. 5 and 6), particularly in the binary apolar mixtures, in which it red-shifts by up to $\sim 50\text{ cm}^{-1}$ and splits into multiple convolved features. In the formamide:CO₂ spectrum, two peaks are present at 15 K at 1316.8 and 1277.0 cm^{-1} , with the peak at 1277.0 cm^{-1} having a greater intensity until 40 K, at which point the intensity of the 1316.8 cm^{-1} peak increases and that of the 1277.0 cm^{-1} peak decreases. The 1277.0 cm^{-1} peak intensity then continues to decrease during heating until CO₂ sublimates at 90 K (see Fig. 5). This trend is indicative of the 1277.0 cm^{-1} peak belonging to the formamide monomer and the 1316.8 cm^{-1} peak belonging to the formamide dimer, as it would be expected for the monomer peak to decrease and the dimer peak to increase if segregation occurs during heating, especially during a major ice structure rearrangement like matrix crystallization, which occurs for CO₂ at 40 K. Such assignments are consistent with the assignments in Mardyukov et al. (2007), who

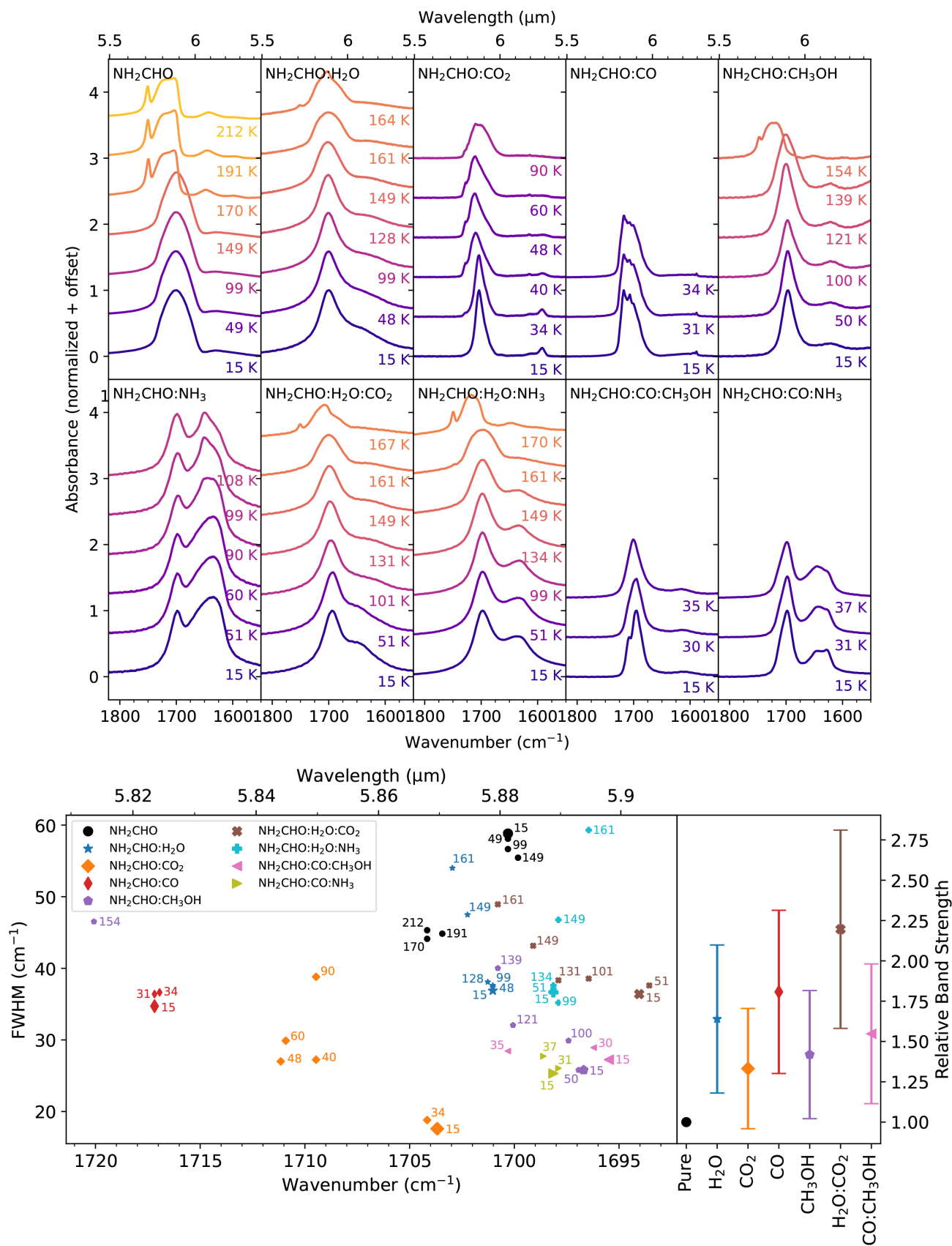


Fig. 4. Characterization of the C = O stretching and NH₂ scissoring bands of NH₂CHO. Top: the C = O stretching and NH₂ scissoring bands of NH₂CHO in pure and mixed forms at selected temperatures. Bottom left: the peak position and FWHM of the formamide C = O stretch in pure and mixed forms at selected temperatures. Bottom right: the band strengths of the formamide C = O stretch and NH₂ scissor in mixtures relative to the pure band strength (all at 15 K).

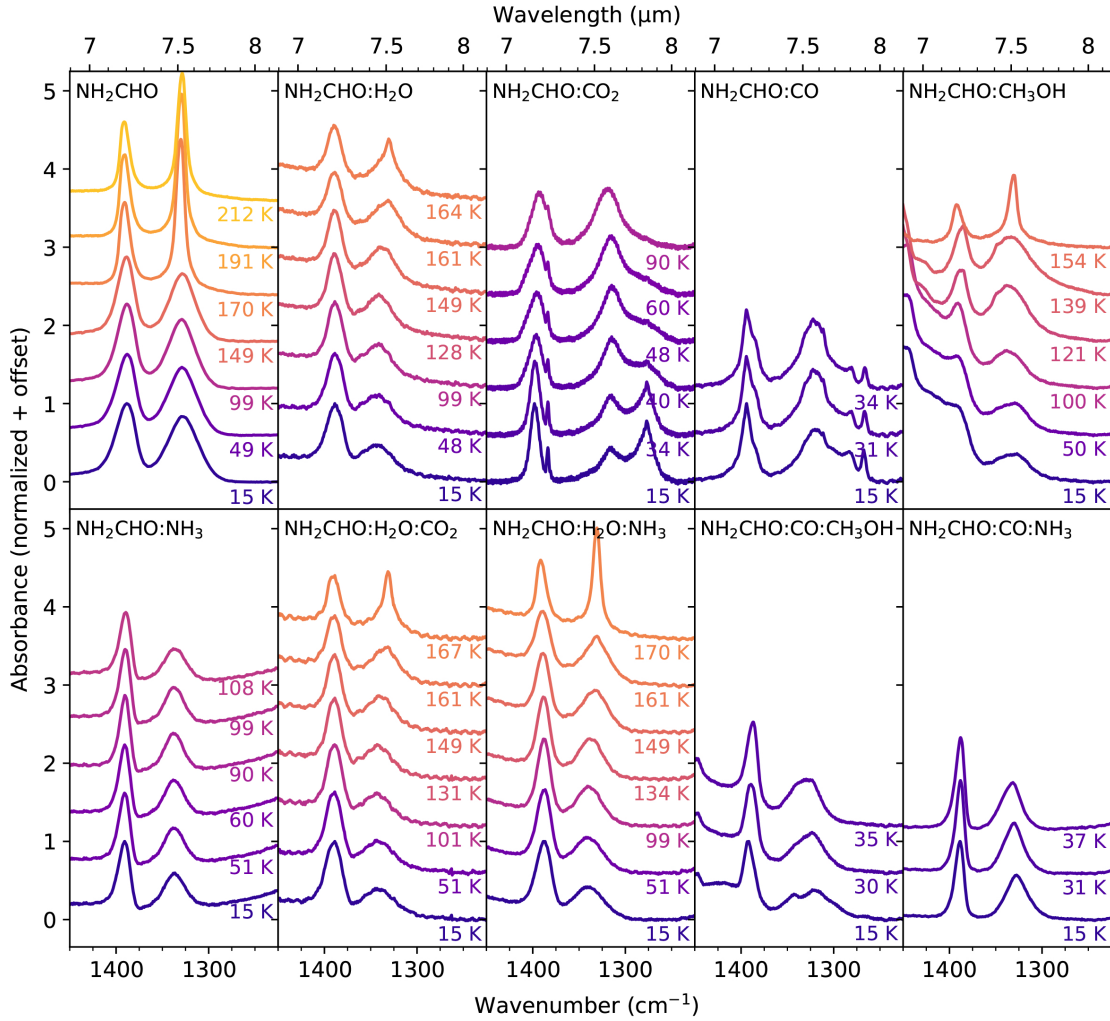


Fig. 5. CH bending and CN stretching bands of NH_2CHO in pure and mixed forms at selected temperatures.

observed the formamide monomer and dimer in a xenon matrix at 1267.2 and 1305.4 cm^{-1} , respectively, and supported their assignments with computations. The peak in the formamide:CO spectrum also has a red component that appears to decrease in intensity during heating, but the monomer and dimer peaks are not as clearly distinguishable as more than two peaks appear to be overlapping in that spectrum. In the mixtures containing other polar molecules, the band is generally blue-shifted, broadened, and decreases in intensity relative to the CH bend. The relative strength of the band is close to 1 in most of the characterized polar mixtures, except for the $\text{H}_2\text{O}:\text{CO}_2$ mixture, which has a relative band strength of 0.75 at 15 K. In contrast, the relative band strength is closer to 2 in all of the primarily apolar mixtures.

While the CN stretch clearly has more potential than the CH bend as a diagnostic of the chemical environment of formamide, it is also much broader and less intense in most of the mixture spectra than in the pure spectra. This diminishes the ability to identify this band in a spectral region where several other astronomically relevant COMs also have features (see Sect. 5).

5. Astronomical implications

The ability of formamide to form via both atom addition and energetic processing in a variety of ices containing C, H, N, and O means that its solid-state presence is plausible in many interstellar environments, ranging from dark interstellar clouds

to protoplanetary disks. However, in order to securely detect it, an absorption with a clear peak position and profile that is distinguishable from other ice features in the same spectral region must be identified.

The $\text{C}=\text{O}$ stretch is amorphous formamide's strongest and sharpest feature, but it overlaps with the blue wing of the strong and broad 6.0 μm feature present in most interstellar ice spectra. Water and ammonia, which have been securely identified in ices, as well as formic acid and formaldehyde, which have been tentatively identified, have features in this spectral region (Boogert et al. 2008, 2015). Additionally, many other carbonyl group-containing COMs that have been detected in the gas-phase and may be present in the solid state, like acetaldehyde, acetone, and methyl formate, also have strong absorptions in this wavelength region (van Scheltinga et al. 2018, 2021; Rachid et al. 2020). While this limits the potential of using formamide's $\text{C}=\text{O}$ band as its primary means of identification, the band can still be used for performing fits spanning a wider wavelength region in combination with other bands.

The CH bend and the CN stretch are medium-strength features that lie in the "COM-rich region" of interstellar ice spectra between 7 and 8 μm (Boogert et al. 2008). This region, where many organic functional groups have absorptions, sits on the tail of the strong 6.85 μm band (whose assignment remains uncertain but likely contains absorptions by methanol and the ammonium cation, Boogert et al. 2008, 2015). The methane

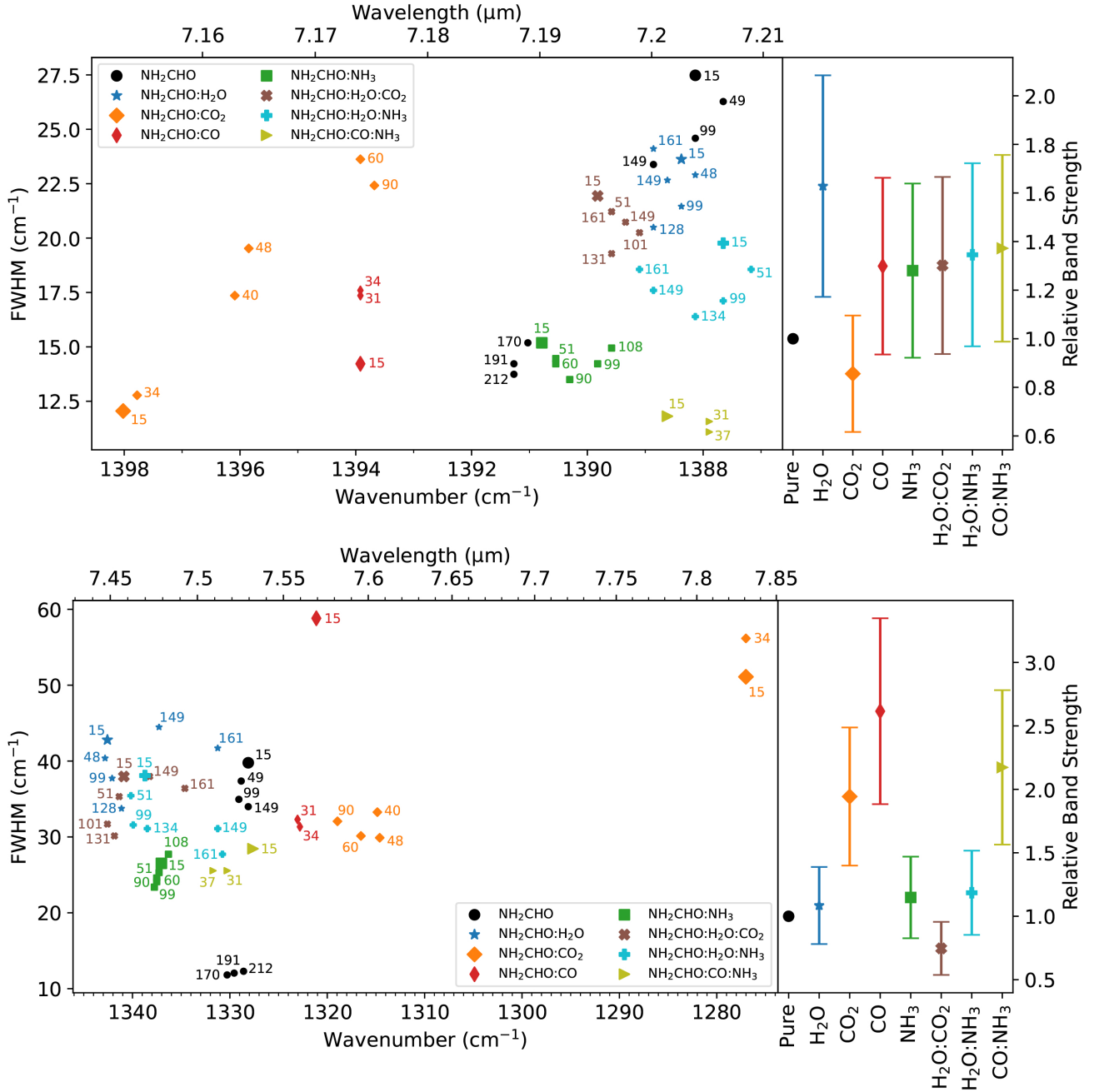


Fig. 6. Characterization of the CH bending and CN stretching bands of NH₂CHO. Left top: the peak position and FWHM of the formamide CH bend in pure and mixed forms at selected temperatures. Right top: the band strengths of the formamide CH bend in mixtures relative to the pure band strength (all at 15 K). Left bottom: the peak position and FWHM of the formamide CN stretch in pure and mixed forms at selected temperatures. Right bottom: the band strengths of the formamide CN stretch in mixtures relative to the pure band strength (all at 15 K).

CH bending band at 7.68 μm is the most clearly and frequently observed ice band in this region (Öberg et al. 2008), but additional weaker features at 7.03, 7.24, 7.41, and 8.01 μm are also consistently observed toward some sources (Fig. 7). Candidate carriers suggested for some of these absorptions include species like formic acid, ethanol, acetaldehyde, the formate anion, and, potentially, formamide (Schutte et al. 1999; Boogert et al. 2008; van Scheltinga et al. 2018).

As mentioned previously, Schutte et al. (1999) tentatively assigned formamide as a plausible contributor to the 7.24 μm band in W33A using a formamide:H₂O spectrum and calculated a formamide ice upper limit of 2.1% with respect to H₂O,

although they pointed out that in their lab data, the formamide peak position was blue-shifted by 0.02 μm relative to the observed band, and that an assignment to the CH bend of formic acid (HCOOH) may be more appropriate. Ethanol (CH₃CH₂OH) and the formate anion (HCOO⁻) have also been considered candidates for this band (Boogert et al. 2008; Öberg et al. 2011; van Scheltinga et al. 2018; Rocha et al., in prep). No distinct and consistently observed bands are located at the peak position of the formamide CN stretch at ~ 7.5 μm . However, in mixtures (particularly those with polar components), the intensity and sharpness of this band weaken (relative to the intensity and sharpness of the CH bend). Such a profile change

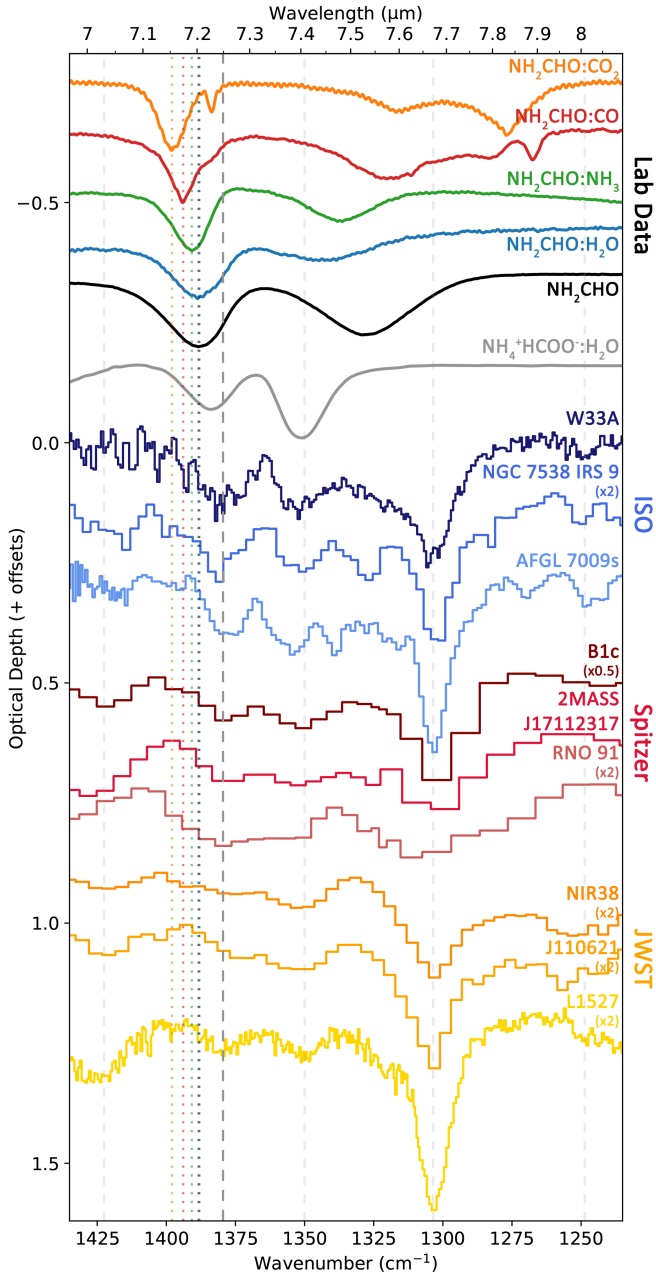


Fig. 7. Laboratory spectra of several formamide mixtures (all at 15 K) as well as an ammonium formate:water mixture from Galvez et al. (2010; at 150 K) versus local continuum-subtracted ISO spectra of three MYSOs (in blue), *Spitzer* spectra of three LYSOs (in red), and JWST spectra of both cold dense clouds in front of background stars and a LYSO (in gold). Optical depth scaling factors (if applied) are indicated by labels under the source names. Several of the band positions marked by gray dashed lines have been tentatively assigned to various COMs. The darkest gray dashed line indicates the peak position of the 7.24 μm band, and the colored dotted lines correspond to the peak positions of the formamide CH bend in the plotted formamide spectra. The wavelength calibration of the MIRI-LRS spectra, NIR38 and J110621, is still uncertain and was done for these spectra locally via the CH_4 band.

makes a distinction of the CN stretch from the continuum in this region less feasible if formamide is present at the low ice abundances expected for COMs, especially given that around this wavelength, many sources also show a broad and significant absorption commonly attributed to SO_2 ice (Boogert et al. 1997; Öberg et al. 2008). On the other hand, the CH bend remains

strong and sharp in all of the mixtures investigated here. All of the other absorption features of formamide either have profiles that are too broad or weak, or overlap directly with the strongest absorptions of the major ice components (see Fig. 3), and will therefore not be utilized in our hunt for formamide ice.

Thus, if formamide is indeed present in interstellar ices, the CH bend is likely its best tracer. We focus our subsequent analysis on the comparison of the formamide CH bend in mixtures to the observed 7.24 μm band in nine spectra collected toward a variety of sources by ISO, *Spitzer*, and the recently launched JWST (Fig. 7). The ISO (SWS) spectra include three massive young stellar objects (MYSOs), W33A, NGC 7538 IRS 9, and AFGL 7009s, and the *Spitzer* (IRS) spectra include three low-mass young stellar objects (LYSOs), B1c, 2MASS J17112317, and RNO 91. These archival spectra were selected due to their 7–8 μm regions having several deep and distinct features, indicating that they may be COM-rich, and because their profiles in this region slightly differ, demonstrating the variety of spectral features that have been observed here. In addition, three spectra recently collected by the JWST have been included: two pristine, high-extinction dark clouds toward background stars, NIR38 and J110621, observed with the Mid-InfraRed Instrument (MIRI) Low-Resolution Spectrometer (LRS; McClure et al. 2023) in the ERS program Ice Age (program 1309, PI: M. McClure), and a Class 0/I low-mass protostar, L1527, observed with the MIRI Medium-Resolution Spectrometer (MRS) in the GTO program JOYS (program I290, PI: E. F. van Dishoeck, van Dishoeck et al. 2023). These are some of the first spectra ever collected of such low-flux sources. While the resolution of the ISO data is comparable to that of the JWST data, the resolution of the *Spitzer* data is significantly lower ($R \sim 60\text{--}100$), limiting its use in the analysis of weak and narrow bands.

The 7.24 μm band is present to some extent in all of the sources, usually at an optical depth similar to the 7.41 μm band in the local continuum-subtracted spectra. The position and FWHM of the band were extracted from the spectra that have spectral resolutions high enough to clearly define the shape and position of the peak – that is, the ISO-SWS and JWST MIRI-MRS MYSO spectra – by fitting a Gaussian profile to the peak. Figure 8 shows these observed peak positions and FWHMs (indicated with star shapes) in a scatter plot with the peak positions and FWHMs of the CH bend extracted from the laboratory spectra of ethanol in a H_2O mixture (van Scheltinga et al. 2018), formic acid in a $\text{H}_2\text{O}:\text{CH}_3\text{OH}$ mixture (Bisschop et al. 2007a), and ammonium formate in a H_2O mixture at 150 K (Galvez et al. 2010) are also included in this figure (indicated with the letters E, F, and H respectively) to enable a comparison between formamide and the other commonly proposed carriers. From this plot, it is evident that, while the polar mixtures have the band position and profile closest to the observations, they are all still too blue-shifted (by $\sim 7\text{ cm}^{-1}/0.04\text{ }\mu\text{m}$) from the astronomical values for formamide to be the major carrier of this band. In contrast, ethanol, formic acid, and the formate anion in polar mixtures are much better candidates.

It is still possible that formamide could be contributing to the blue wing of this band. However, to result in non-negligible upper limits, the formamide must be present in a matrix containing other polar molecules, as the band is far too blue-shifted in the purely apolar mixtures to contribute significantly to the observed absorption. Therefore, we derived upper limits of formamide by fitting the CH bend in the laboratory spectrum of the formamide: H_2O mixture at 15 K to the 7.24 μm band in the local continuum-subtracted observed spectra (see example

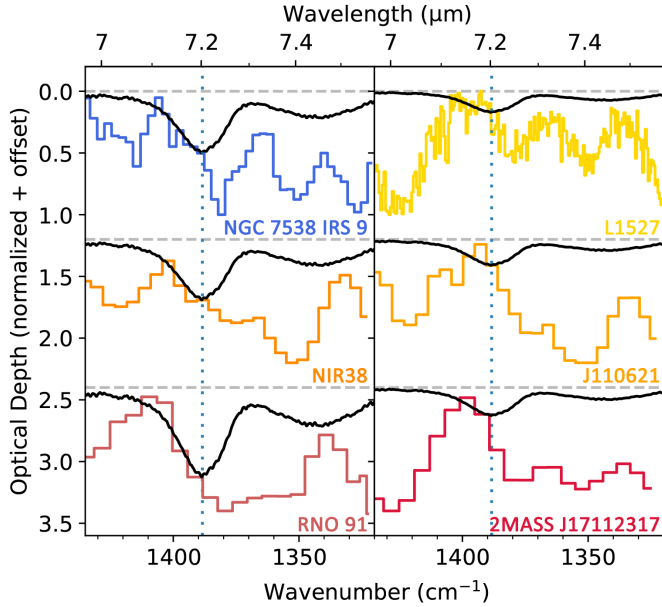


Fig. 9. Examples of fits of the CH bending band in the formamide:H₂O laboratory spectrum (in black, water contribution subtracted out) to the local-continuum subtracted observed spectra (in color) used to derive solid-state formamide upper limits. The peak position of the CH bend in the laboratory spectrum is marked with a dotted blue vertical line. These selected fits showcase the variation in the profile of the observed 7.24 μm band across the selected sources. In the observed spectra on the left, excess absorption on the blue wing of the observed 7.24 μm band allows for relatively high formamide upper limits ($\sim\leq 1.5\text{--}5.1\%$ with respect to H₂O), while the observed spectra on the right lack such a wing, resulting in lower formamide upper limits ($\sim\leq 0.35\text{--}0.68\%$ with respect to H₂O).

precision of this upper limit to be further refined by future high-resolution observations of B1c, planned to be observed by MIRI-MRS in the JOYS program.

A formamide upper limit of 2.1% with respect to H₂O was previously derived for W33A in Schutte et al. (1999) by assuming that the entire 7.24 μm band consisted of formamide and using a band strength of 3.2×10^{-18} cm molec⁻¹ attributed to Wexler (1967), where it is unclear for what phase of formamide this band strength was derived. Despite our very different approaches, we have fortuitously arrived at nearly the same upper limit value for W33A (2.2%).

In the higher resolution observational data of MYSOs explored here, the lack of a formamide CH bending feature distinct from other COM absorptions prevents a secure formamide ice detection. However, it is clear from the example upper limit fits shown in Fig. 9 that the profile of the 7.24 μm feature is not uniform across different sources, and several sources, such as NGC 7538 IRS 9, NIR38, and RNO 91, may have a blue wing on this band that spectrally overlaps with the CH bend of formamide. Therefore, it is possible that a more distinct absorption at the expected 7.20 μm will emerge more clearly in sources targeted by future JWST MIRI-MRS observations. The first ice spectra arriving now from MIRI-MRS illuminate a promising future. In the spectrum of the LYSO IRAS 15398-3359 acquired by the JWST CORINOS program (program 2151, PI: Y.-L. Yang, Yang et al. 2022), the COM features between 7 and 8 μm previously detected barely above 3σ levels in the spectra in Fig. 7 are beautifully resolved (although a distinct absorption centered at 7.20 μm is not present). More sources known to have strong COM absorptions in this spectral region have been specifically

Table 3. Upper limits of formamide with respect to H₂O compared to observed abundances of formamide in various objects.

Object	$N(\text{NH}_2\text{CHO})$ 10^{17} molec cm ⁻²	% H ₂ O
MYSOs		
W33A	≤ 2.8	≤ 2.2
NGC 7538 IRS 9	≤ 2.0	≤ 3.1
AFGL 7009s	≤ 1.4	≤ 1.2
LYSOs		
B1c	≤ 2.8	≤ 0.93
2MASS J17112317	≤ 1.3	≤ 0.68
RNO 91	≤ 2.2	≤ 5.1
L1527	≤ 1.0	≤ 0.50
Prestellar cores		
NIR38	≤ 1.0	≤ 1.5
J110621	≤ 0.47	≤ 0.35
LYSOs		
B1c (gas)	–	$\sim 0.05^{(a)}$
MYSOs		
ALMAGAL average (gas)	–	$\sim 0.05^{(b)}$
Comets		
67P (dust and gas)	–	$0.004^{(c)}$
Lemmon (gas)	–	$0.016^{(d)}$
Lovejoy (gas)	–	$0.021^{(d)}$
Hale-Bopp (gas)	–	$0.02^{(e)}$

Notes. ^(a)Value calculated via the beam dilution-corrected NH₂CHO/CH₃OH ratio from Nazari et al. (2023), assuming a CH₃OH/H₂O ratio of $\sim 5\%$. ^(b)Value calculated via the beam dilution-corrected NH₂CHO/CH₃OH ratio of 22 MYSO sources from Nazari et al. (2023), assuming a CH₃OH/H₂O ratio of $\sim 5\%$. ^(c)Value from Rubin et al. (2019). ^(d)Value from Biver et al. (2014). ^(e)Value from Bockelée-Morvan et al. (2000).

targeted by the JOYS program as well as the JWST proposals “It’s COMplicated” (program 1854, PI: M. McClure, McClure et al. 2021) and “Ice chemical complexity toward the Ophiuchus molecular cloud” (program 1959, PI: W. R. M. Rocha, Rocha et al. 2021), and these sources are scheduled to be observed throughout the remainder of this year. As demonstrated by the examples of spectral analysis of ices in this section, the laboratory spectra from this work can serve as a toolkit for formamide identification in such ice observations.

6. Conclusions

In an effort to facilitate the hunt for formamide in interstellar ices, laboratory spectra of pure formamide and formamide in various astronomically relevant ice mixtures ranging from temperatures of 15–212 K have been collected and made freely available to the astronomical community on the Leiden Ice Database for Astrochemistry (LIDA). The band strengths at 15 K for all pure formamide features between 4000 and 500 cm⁻¹/2.5–20 μm are presented, and the peak positions, FWHMs, and relative apparent band strengths of the three bands identified as the most promising for future formamide detection were

extracted from the pure and mixed formamide spectra. These spectra and extracted data were used to assess present and future detectability of ices in various interstellar objects. The primary conclusions drawn from this work are as follows:

- Out of the eight formamide features in the investigated IR spectral region, the C = O stretch ($1700.9\text{ cm}^{-1}/5.881\text{ }\mu\text{m}$), the CH bend ($1388.3\text{ cm}^{-1}/7.203\text{ }\mu\text{m}$), and the CN stretch ($1328.0\text{ cm}^{-1}/7.530\text{ }\mu\text{m}$) are likely to be the most useful for future formamide identification due to their strength, sharp profile, and low overlap with the strongest features of the major ice components, with the CH bending feature being the most promising. The NH_2 stretching features ($3371.2\text{ cm}^{-1}/2.966\text{ }\mu\text{m}$ and $3176.4\text{ cm}^{-1}/3.148\text{ }\mu\text{m}$) and the NH_2 wagging and twisting features ($689.2\text{ cm}^{-1}/14.510\text{ }\mu\text{m}$ and $634.0\text{ cm}^{-1}/15.773\text{ }\mu\text{m}$) directly overlap with strong water absorptions, while the CH stretch ($2881.9\text{ cm}^{-1}/3.470\text{ }\mu\text{m}$), the CH bend overtone ($2797.7\text{ cm}^{-1}/3.574\text{ }\mu\text{m}$), and the convolved NH_2 rock and CH out-of-plane deformation ($1108.1\text{ cm}^{-1}/9.024\text{ }\mu\text{m}$ and $1056.1\text{ cm}^{-1}/9.469\text{ }\mu\text{m}$) have both low band strengths and direct overlap with methanol absorptions, making them less suitable for formamide identification.
- In the mixtures investigated here, the CN stretch is the most affected by ice composition – its peak position varies by up to $\sim 68\text{ cm}^{-1}$ and its FWHM by up to $\sim 50\text{ cm}^{-1}$ across the mixtures investigated here, with peak splitting observed in the apolar mixtures. The C = O stretch can also change significantly, depending on the matrix, by up to $\sim 27\text{ cm}^{-1}$ in peak position and up to $\sim 40\text{ cm}^{-1}$ in FWHM, although peak splitting in the apolar mixtures is not as prominent as in the CN stretch. The CH bend is relatively unaffected by ice composition, with its peak position and FWHM only varying by $\sim 11\text{ cm}^{-1}$ and $\sim 15\text{ cm}^{-1}$, respectively, across the different mixtures. Relative to the pure spectrum, the band strength of the C = O stretch increases in all of the investigated mixtures. The CH bend band strength also increases in all of the mixtures except the binary CO_2 mixture, while a significant increase in the band strength of the CN stretch is only observed in the mixtures dominated by an apolar component.
- Although the polar formamide mixtures provide the closest match to the $7.24\text{ }\mu\text{m}$ band observed toward nine lines of sight (including dense clouds, LYSOs, and MYSOs) with three different space telescopes (ISO, *Spitzer*, and JWST), none provide a convincing fit, with all having their CH bend peak position approximately $7\text{ cm}^{-1}/0.04\text{ }\mu\text{m}$ too far to the blue from the clearly observed band at $1381\text{ cm}^{-1}/7.24\text{ }\mu\text{m}$. Instead, formic acid and ethanol mixtures containing H_2O provide a better fit. However, this does not exclude the possibility of formamide being present in these ices. The calculated formamide upper limits in these objects range from 0.35 to 5.1% with respect to H_2O , which are consistent with gas-phase abundances of formamide in several LYSOs, MYSOs, and comets. The upper limit value derived for W33A, 2.2% with respect to H_2O , is fortuitously in agreement with that derived by Schutte et al. (1999).

While a more secure formamide detection is not possible with the telescopic data explored in this work, the first ice observations arriving from JWST demonstrate an unprecedented sensitivity and spectral resolution that will enable us in the near future to broaden the search for formamide ice in both objects previously observed by *Spitzer*, whose analysis is limited by low

spectral resolution, as well as newly observed objects that were too dim to be observed by *Spitzer* or ISO.

Acknowledgements. This work is supported by funding from the European Research Council (ERC) under the European Union’s Horizon 2020 research and innovation programme (grant agreement No. 101019751 MOLDISK), the Netherlands Research School for Astronomy (NOVA), and the Danish National Research Foundation through the Center of Excellence “InterCat” (Grant agreement no.: DNR150). The authors acknowledge the Ice Age (program 1309, PI: M. McClure) and JOYS (program 1290, PI: E. F. van Dishoeck) observing programs for the JWST astronomical data used in this work. K.S. acknowledges Thanja Lamberts and Pooneh Nazari for helpful discussions about the formamide formation mechanism and Sergio Ioppolo for helpful discussions about the QMS calibration methodology.

References

- Altwegg, K., Balsiger, H., Berthelier, J.-J., et al. 2017, *MNRAS*, **469**, S130
- Barks, H. L., Buckley, R., Grieves, G. A., et al. 2010, *Chem. Bio.Chem.*, **11**, 1240
- Barone, V., Latouche, C., Skouteris, D., et al. 2015, *MNRAS*, **453**, L31
- Basford, J. A., Boeckmann, M. D., Ellefson, R. E., et al. 1993, *J. Vacuum Sci. Technol. A Vacuum Surf. Films*, **11**, A22
- Bisschop, S., Fuchs, G., Boogert, A., Van Dishoeck, E., & Linnartz, H. 2007a, *A&A*, **470**, 749
- Bisschop, S., Jørgensen, J., Van Dishoeck, E., & De Wachter, E. 2007b, *A&A*, **465**, 913
- Biver, N., Bockelée-Morvan, D., Debout, V., et al. 2014, *A&A*, **566**, L5
- Bockelée-Morvan, D., Lis, D., Wink, J., et al. 2000, *A&A*, **353**, 1101
- Boogert, A., Schutte, W., Helmich, F., Tielens, A., & Wooden, D. 1997, *A&A*, **317**, 929
- Boogert, A. C., Pontoppidan, K. M., Knez, C., et al. 2008, *ApJ*, **678**, 985
- Boogert, A. A., Gerakines, P. A., & Whittet, D. C. 2015, *ARA&A*, **53**, 541
- Bouilloud, M., Fray, N., Bénilan, Y., et al. 2015, *MNRAS*, **451**, 2145
- Bredehöft, J. H., Böhrer, E., Schmidt, F., Borrmann, T., & Swiderek, P. 2017, *ACS Earth Space Chem.*, **1**, 50
- Brucato, J. R., Baratta, G. A., & Strazzulla, G. 2006, *A&A*, **455**, 395
- Chaabouni, H., Diana, S., Nguyen, T., & Dulieu, F. 2018, *A&A*, **612**, A47
- Charnley, S. 1997, *Astronomical and Biochemical Origins and Search for Life in the Universe*, (Editrice Compositiori), 89
- Chuang, K.-J., Jäger, C., Krasnokutski, S., Fulvio, D., & Henning, T. 2022, *ApJ*, **933**, 107
- Ciaravella, A., Jiménez-Escobar, A., Cecchi-Pestellini, C., et al. 2019, *ApJ*, **879**, 21
- Codella, C., Ceccarelli, C., Caselli, P., et al. 2017, *A&A*, **605**, L3
- Corazzi, M. A., Fedele, D., Poggiali, G., & Brucato, J. R. 2020, *A&A*, **636**, A63
- Coutens, A., Jørgensen, J., Van der Wiel, M., et al. 2016, *A&A*, **590**, L6
- Dawley, M. M., Pirim, C., & Orlando, T. M. 2014, *J. Phys. Chem. A*, **118**, 1220
- Demyk, K., Dartois, E., d’Hendecourt, L., et al. 1998, *A&A*, **339**, 553
- Douglas, K. M., Lucas, D. I., Walsh, C., et al. 2022, *ApJ*, **937**, L16
- Dulieu, F., Nguyen, T., Congiu, E., Baouche, S., & Taquet, V. 2019, *MNRAS*, **484**, L119
- Ehrenfreund, P., Schutte, W. A., & Gerakines, P. 1996, *A&A*, **312**, 263
- Ellefson, R. E., Cain, D., & Lindsay, C. N. 1987, *J. Vacuum Sci. Technol. A Vacuum Surf. Films*, **5**, 134
- Endres, C. P., Schlemmer, S., Schilke, P., Stutzki, J., & Müller, H. S. 2016, *J. Mol. Spectr.*, **327**, 95
- Fedoseev, G., Chuang, K.-J., van Dishoeck, E. F., Ioppolo, S., & Linnartz, H. 2016, *MNRAS*, **460**, 4297
- Galvez, O., Maté, B., Herrero, V. J., & Escribano, R. 2010, *ApJ*, **724**, 539
- Garrod, R. T., Weaver, S. L. W., & Herbst, E. 2008, *ApJ*, **682**, 283
- Gerakines, P., Schutte, W., Greenberg, J., & van Dishoeck, E. F. 1995, *A&A*, **296**, 810
- Gerakines, P., Moore, M., & Hudson, R. 2004, *Icarus*, **170**, 202
- Gerakines, P. A., Materese, C. K., & Hudson, R. L. 2023, *MNRAS*, **522**, 3145
- Goesmann, F., Rosenbauer, H., Bredehöft, J. H., et al. 2015, *Science*, **349**, aab0689
- González Díaz, C., Carrascosa, H., Muñoz Caro, G. M., Satorre, M. Á., & Chen, Y. 2022, *MNRAS*, **517**, 5744
- Gull, M., Cafferty, B. J., Hud, N. V., & Pasek, M. A. 2017, *Life*, **7**, 29
- Haupa, K. A., Tarczay, G., & Lee, Y.-P. 2019, *J. Am. Chem. Soc.*, **141**, 11614
- He, J., Toriello, F. E., Emtiaz, S. M., Henning, T., & Vidali, G. 2021, *ApJ*, **915**, L23
- Hudson, R., & Moore, M. 2000, *A&A*, **357**, 787
- Hudson, R., Ferrante, R., & Moore, M. 2014, *Icarus*, **228**, 276
- Hudson, R. L., Gerakines, P. A., & Yarnall, Y. Y. 2022, *ApJ*, **925**, 156

- Jones, B. M., Bennett, C. J., & Kaiser, R. I. 2011, *ApJ*, 734, 78
- Kaňuchová, Z., Urso, R., Baratta, G., et al. 2016, *A&A*, 585, A155
- Kwiatkowski, J. S., & Leszczyński, J. 1993, *J. Mol. Struct.*, 297, 277
- Lago, J. L., Burcar, B. T., Hud, N. V., et al. 2020, *Origins Life Evol. Biosph.*, 50, 1
- Lieszkovszky, L., Filippelli, A., & Tilford, C. 1990, *J. Vacuum Sci. Technol. A Vacuum Surf. Films*, 8, 3838
- Ligterink, N., Terwisscha van Scheltinga, J., Taquet, V., et al. 2018, *MNRAS*, 480, 3628
- López-Sepulcre, A., Jaber, A. A., Mendoza, E., et al. 2015, *MNRAS*, 449, 2438
- López-Sepulcre, A., Balucani, N., Ceccarelli, C., et al. 2019, *ACS Earth Space Chem.*, 3, 2122
- Luna, R., Molpeceres, G., Ortigoso, J., et al. 2018, *A&A*, 617, A116
- Lundell, J., Krajewska, M., & Räsänen, M. 1998, *J. Phys. Chem. A*, 102, 6643
- Mardukov, A., Sánchez-García, E., Rodziewicz, P., Doltsinis, N. L., & Sander, W. 2007, *J. Phys. Chem. A*, 111, 10552
- Martín-Doménech, R., Öberg, K. I., & Rajappan, M. 2020, *ApJ*, 894, 98
- McClure, M., Bergin, E. A., Boogert, A. C., et al. 2021, *JWST Proposal. Cycle 1*, 1854
- McClure, M. K., Rocha, W., Pontoppidan, K., et al. 2023, *Nat. Astron.*, 7, 431
- Mendoza, E., Lefloch, B., López-Sepulcre, A., et al. 2014, *MNRAS*, 445, 151
- Nazari, P., Meijerhof, J., van Gelder, M., et al. 2022, *A&A*, 668, A109
- Nazari, P., Tabone, B., Rosotti, G., & van Dishoeck, E. 2023, *A&A*, submitted
- Nemanič, V., & Žumer, M. 2018, *AIP Adv.*, 8, 015128
- Noble, J., Theule, P., Congiu, E., et al. 2015, *A&A*, 576, A91
- Öberg, K. I., Fraser, H. J., Boogert, A. A., et al. 2007, *A&A*, 462, 1187
- Öberg, K. I., Boogert, A. A., Pontoppidan, K. M., et al. 2008, *ApJ*, 678, 1032
- Öberg, K. I., Boogert, A. A., Pontoppidan, K. M., et al. 2011, *ApJ*, 740, 109
- Pasek, M. A. 2019, *Chem. Rev.*, 120, 4690
- Quénard, D., Jiménez-Serra, I., Viti, S., Holdship, J., & Coutens, A. 2018, *MNRAS*, 474, 2796
- Rachid, M., van Scheltinga, J. T., Koletzki, D., & Linnartz, H. 2020, *A&A*, 639, A4
- Rachid, M., Brunken, N., De Boe, D., et al. 2021, *A&A*, 653, A116
- Rachid, M. G., Rocha, W., & Linnartz, H. 2022, *A&A*, 665, A89
- Räsänen, M. 1983, *J. Mol. Struct.*, 101, 275
- Raunier, S., Chiavassa, T., Duvernay, F., et al. 2004, *A&A*, 416, 165
- Rimola, A., Skouteris, D., Balucani, N., et al. 2018, *ACS Earth Space Chem.*, 2, 720
- Rocha, W., Jorgensen, J., Karska, A., Kristensen, L. E., & Perotti, G. 2021, *JWST Proposal. Cycle 1*, 1959
- Rocha, W., Rachid, M., Olsthoorn, B., et al. 2022, *A&A*, 668, A63
- Rubin, R., Swenson Jr, G., Benson, R., Tigelaar, H., & Flygare, W. 1971, *ApJ*, 169, L39
- Rubin, M., Altwegg, K., Balsiger, H., et al. 2019, *MNRAS*, 489, 594
- Saladino, R., Ciambecchini, U., Crestini, C., et al. 2003, *Chem. Bio. Chem.*, 4, 514
- Schutte, W., Boogert, A., Tielens, A., et al. 1999, *A&A*, 343, 966
- Sivaraman, B., Sekhar, B. R., Nair, B., Hatode, V., & Mason, N. 2013, *Spectrochim. Acta Part A Mol. Biomol. Spectr.*, 105, 238
- Skouteris, D., Vazart, F., Ceccarelli, C., et al. 2017, *MNRAS*, 468, L1
- Song, L., & Kästner, J. 2016, *Phys. Chem. Chem. Phys.*, 18, 29278
- Sun, W., Wu, C., Cheng, Y., et al. 2020, *Measurement*, 164, 108099
- Torrie, B., & Brown, B. 1994, *J. Raman Spectr.*, 25, 183
- Urso, R. G., Scirè, C., Baratta, G. A., et al. 2017, *Phys. Chem. Chem. Phys.*, 19, 21759
- van Dishoeck, E. F., Grant, S., Tabone, B., et al. 2023, *Faraday Discussions*, in press, <https://doi.org/10.1039/D3FD00010A>
- van Scheltinga, J. T., Ligterink, N., Boogert, A., Van Dishoeck, E., & Linnartz, H. 2018, *A&A*, 611, A35
- van Scheltinga, J. T., Marcandalli, G., McClure, M. K., Hogerheijde, M. R., & Linnartz, H. 2021, *A&A*, 651, A95
- Vazart, F., Calderini, D., Puzzarini, C., Skouteris, D., & Barone, V. 2016, *J. Chem. Theor Comput.*, 12, 5385
- Wexler, A. S. 1967, *Appl. Spectrosc. Rev.*, 1, 29
- Yang, Y.-L., Green, J. D., Pontoppidan, K. M., et al. 2022, *ApJ*, 941, L13
- Yarnall, Y. Y., & Hudson, R. L. 2022, *ApJ*, 931, L4
- Yongjun, C., Meng, D., Wenjun, S., et al. 2022, *Measure. Sci. Technol.*, 33, 065019

Appendix A: Peak positions and FWHMs of formamide in pure and mixed ices

This appendix contains the peak positions and FWHMs of the formamide features selected for complete IR characterization in this work. The values are listed for the formamide features in pure ice as well as in mixtures containing H₂O, CO₂, CO, CH₃OH, and NH₃. The peak position is the wavelength at which the absorption reaches its maximum, and the FWHM is the width of the peak between the half-maximum values on each side. A Savitzky-Golay filter with a second-order polynomial was applied to many of the mixture spectra here before extraction of the peak position and FWHM to eliminate shifts in these values caused by noise. The smoothing windows used ranged from 10-100 depending on level of noise present in each spectrum, and care was taken that these smoothing windows did not warp the shape of any features. Values were extracted until the temperatures at which the major matrix component desorbed were reached.

For formamide features in mixtures where there is direct overlap with weaker matrix component bands (e.g., the C = O stretch in the NH₂CHO:H₂O mixture), the spectrum of the matrix component without formamide collected using identical experimental parameters at the corresponding temperature was scaled to the formamide mixture spectrum via a feature without overlap with formamide features and subtracted prior to peak position and FWHM extraction. These cases are denoted with a ^M. For formamide features in mixtures where the formamide features lie on the tails of bands or on very wide bands without sharp features (e.g., the CH bend and CN stretch in the NH₂CHO:NH₃ mixture), a second-order polynomial was used to perform a local continuum subtraction. These cases are denoted with a ^P. For formamide features in mixtures where overlap with a strong matrix component band was very substantial and difficult to reliably subtract (e.g., the C = O stretch in the NH₂CHO:NH₃ mixture), only peak positions are given. These cases are denoted with a ^N. For formamide features that contain multiple peaks, all peak positions are given, and the FWHM of the strongest peak is given. However, if a weaker peak maximum occurs within the two half maximum values of the stronger peak (e.g., the CN stretch in the NH₂CHO:CO 15 K mixture), it is included in the FWHM. These cases are denoted with a ^B.

Table A.1: Peak positions and FWHMs of the C = O stretch and NH₂ scissor bands of pure formamide and formamide in various binary mixtures at temperatures from 15-212 K.

Matrix	T (K)	Peak		FWHM	
		(cm ⁻¹)	(μm)	(cm ⁻¹)	(μm)
Pure	15	1700.3	5.881	58.8	0.203
		1630.4	6.133	-	-
	49	1700.3	5.881	58.1	0.201
		1632.6	6.125	-	-
	99	1700.3	5.881	56.6	0.196
		1631.1	6.131	-	-
	149	1699.8	5.883	55.4	0.192
		1629.4	6.137	-	-
		1704.2	5.868	44.1	0.149
	170	1749.2	5.717	-	-
		1647.0	6.072	-	-
		1703.4	5.870	44.8	0.152
	191	1750.0	5.714	-	-
		1644.6	6.080	-	-
1704.2		5.868	45.3	0.154	
1750.4		5.713	-	-	
212	1642.2	6.089	-	-	
	1701.0	5.879	36.9	0.128	
	1701.0	5.879	37.4	0.129	
	1701.0	5.879	37.6	0.130	
H ₂ O ^M	129	1701.3	5.878	38.1	0.132
	149	1702.2	5.875	47.5	0.164
	161	1703.0	5.872	54.0	0.187
	15	1703.7	5.870	17.6	0.061
		1591.6	6.283	-	-
	34	1704.2	5.868	18.8	0.065
1591.6		6.283	-	-	
CO ₂	40	1709.5	5.850	27.2	0.093
		1591.6	6.283	-	-
	49	1711.2	5.844	27.0	0.093
		1615.2	6.191	-	-
	60	1710.9	5.845	29.9	0.103
		1614.0	6.196	-	-
90	1709.5	5.850	38.8	0.134	
CO	15	1717.2	5.824	34.7	0.119
		1588.9	6.294	-	-
	31	1717.2	5.824	36.4	0.125
		1588.7	6.294	-	-
	34	1716.9	5.824	36.6	0.126
		1588.7	6.294	-	-
CH ₃ OH ^M	15	1696.7	5.894	25.8	0.090
		1623.5	6.159	-	-
	50	1696.9	5.893	25.8	0.090
		1623.3	6.160	-	-
	100	1697.4	5.891	29.9	0.104
		1621.4	6.168	-	-
	121	1700.1	5.882	32.1	0.111
		1621.4	6.168	-	-
	139	1700.8	5.880	40.0	0.139
		1621.4	6.168	-	-
1720.1		5.814	46.5	0.156	
154	1646.9	6.072	-	-	
NH ₃ ^N	15	1698.4	5.888	-	-
	51	1698.1	5.889	-	-
	60	1697.9	5.890	-	-
	90	1696.9	5.893	-	-
	99	1697.7	5.890	-	-
	108	1698.9	5.886	-	-

Table A.2: Peak positions and FWHMs of the C = O stretch and NH₂ scissor bands of formamide in various tertiary mixtures at temperatures from 15-161 K.

Matrix	T (K)	Peak		FWHM	
		(cm ⁻¹)	(μm)	(cm ⁻¹)	(μm)
H ₂ O:CO ₂ ^M	15	1694.0	5.903	36.4	0.127
	51	1693.6	5.905	37.6	0.131
	101	1696.4	5.895	38.6	0.134
	131	1697.9	5.890	38.3	0.133
	149	1699.1	5.885	43.1	0.150
	161	1700.8	5.880	48.9	0.169
H ₂ O:NH ₃ ^M	15	1698.1	5.889	36.6	0.127
	51	1698.1	5.889	37.1	0.129
	99	1697.9	5.890	35.2	0.122
	134	1698.1	5.889	37.6	0.131
	149	1697.9	5.890	46.8	0.162
	161	1696.4	5.895	59.3	0.206
CO:CH ₃ OH ^P	15	1695.5	5.898	27.2	0.095
		1615.2	6.191	-	-
	30	1696.2	5.896	28.9	0.100
		1615.2	6.191	-	-
	35	1700.3	5.881	28.4	0.099
		1617.1	6.184	-	-
CO:NH ₃ ^M	15	1698.1	5.889	25.3	0.088
	30	1697.9	5.890	26.0	0.090
	37	1698.6	5.887	27.7	0.096

Table A.3: Peak positions and FWHMs of the CH bend band of pure formamide and formamide in various binary and tertiary mixtures at temperatures from 15-212 K.

Matrix	T (K)	Peak		FWHM	
		(cm ⁻¹)	(μm)	(cm ⁻¹)	(μm)
Pure	15	1388.1	7.204	27.5	0.142
	49	1387.7	7.206	26.3	0.136
	99	1388.1	7.204	24.6	0.127
	149	1388.9	7.200	23.4	0.121
	170	1391.0	7.189	15.2	0.079
	191	1391.3	7.188	14.2	0.074
	212	1391.3	7.188	13.7	0.071
H ₂ O ^M	15	1388.4	7.203	23.6	0.122
	49	1388.1	7.204	22.9	0.119
	99	1388.4	7.203	21.5	0.111
	129	1388.9	7.200	20.5	0.106
	149	1388.6	7.201	22.7	0.117
	161	1388.9	7.200	24.1	0.125
	CO ₂	15	1398.0	7.153	12.1
		1383.7	7.227	-	-
34		1397.8	7.154	12.8	0.065
		1383.7	7.227	-	-
40		1396.1	7.163	17.4	0.089
		1383.9	7.226	-	-
49		1395.8	7.164	19.5	0.100
		1383.9	7.226	-	-
60		1393.9	7.174	23.6	0.122
		1383.9	7.226	-	-
90	1393.7	7.175	22.4	0.116	
	1384.4	7.223	-	-	
CO	15	1393.9	7.174	14.2	0.073
	31	1393.9	7.174	17.4	0.090
	34	1393.9	7.174	17.6	0.091
NH ₃ ^P	15	1390.8	7.190	15.2	0.078
	51	1390.5	7.191	14.5	0.075
	60	1390.5	7.191	14.2	0.073
	90	1390.3	7.193	13.5	0.070
	99	1389.8	7.195	14.2	0.074
	108	1389.6	7.196	14.9	0.077
H ₂ O:CO ₂ ^M	15	1389.9	7.195	21.9	0.114
	51	1389.6	7.196	21.2	0.110
	101	1389.1	7.199	20.2	0.105
	131	1389.6	7.196	19.3	0.100
	149	1389.3	7.198	20.7	0.107
	161	1389.6	7.196	21.2	0.110
H ₂ O:NH ₃ ^P	15	1387.7	7.206	19.8	0.103
	51	1387.2	7.209	18.6	0.096
	99	1387.7	7.206	17.1	0.089
	134	1388.1	7.204	16.4	0.085
	149	1388.9	7.200	17.6	0.091
	161	1389.1	7.199	18.6	0.096
CO:NH ₃ ^M	15	1388.6	7.201	11.8	0.061
	30	1387.9	7.205	11.1	0.058
	37	1387.9	7.205	11.6	0.060

Table A.4: Peak positions and FWHMs of the CN stretch band of pure formamide and formamide in various binary and tertiary mixtures at temperatures from 15-212 K.

Matrix	T (K)	Peak		FWHM	
		(cm ⁻¹)	(μ m)	(cm ⁻¹)	(μ m)
Pure	15	1328.1	7.529	39.8	1.226
	49	1328.8	7.525	37.4	0.212
	99	1329.1	7.524	35.0	0.198
	149	1328.1	7.529	34.0	0.192
	170	1330.3	7.517	11.8	0.067
	191	1329.6	7.521	12.1	0.068
	212	1328.6	7.527	12.3	0.070
H ₂ O ^M	15	1342.6	7.448	42.8	0.237
	49	1342.8	7.447	40.4	0.223
	99	1342.1	7.451	37.7	0.209
	129	1341.1	7.456	33.7	0.188
	149	1337.3	7.478	44.5	0.248
	161	1331.2	7.512	41.7	0.234
CO ₂	15	1277.0	7.831	51.1	0.305
		1316.8 ^B	7.594	-	-
	34	1277.0	7.831	56.2	0.334
		1316.8 ^B	7.594	-	-
	40	1314.9	7.605	33.3	0.194
	49	1314.6	7.607	29.9	0.173
	90	1319.0	7.582	32.1	0.184
CO	15	1321.1	7.569	58.8	0.344
		1283.3 ^B	7.793	-	-
		1267.6	7.889	-	-
	31	1323.0	7.558	32.3	0.185
		1280.9	7.807	-	-
		1266.9	7.893	-	-
	34	1322.8	7.560	31.3	0.179
		1280.6	7.809	-	-
	1266.4	7.896	-	-	
NH ₃ ^P	15	1337.0	7.479	26.5	0.148
	51	1337.3	7.478	25.3	0.142
	60	1337.5	7.477	24.6	0.138
	90	1337.8	7.475	23.4	0.131
	99	1337.5	7.477	24.1	0.135
	108	1336.3	7.483	27.7	0.155
H ₂ O:CO ₂ ^M	15	1340.9	7.458	38.0	0.211
	51	1341.4	7.455	35.3	0.196
	101	1342.6	7.448	31.7	0.176
	131	1341.9	7.452	30.1	0.167
	149	1338.2	7.473	38.0	0.211
	161	1334.6	7.493	36.4	0.204
H ₂ O:NH ₃ ^P	15	1338.7	7.470	38.1	0.212
	51	1340.2	7.462	35.4	0.198
	99	1339.9	7.463	31.6	0.176
	134	1339.0	7.468	30.9	0.173
	149	1331.2	7.512	31.1	0.175
	161	1330.8	7.514	27.7	0.157
CO:NH ₃ ^M	15	1327.6	7.532	28.4	0.162
	30	1330.3	7.517	25.6	0.144
	37	1331.7	7.509	25.6	0.144

Appendix B: Relative apparent band strengths of formamide in pure and mixed ices

This appendix provides the relative apparent band strengths (η) of formamide, calculated via Equation 4, where the value of A' used in the calculations is the apparent band strength of the respective band in the pure amorphous formamide ice at 15 K (given in Table 2). Thus, the η value of each band in the pure ice at 15 K is unity. The integration ranges used to calculate the integrated absorbances are stated for each mixture individually, as the same integration ranges were not used for all mixtures due to shifting peak positions and FWHMs. Different integration ranges were used to calculate the integrated absorbances of the amorphous and crystalline pure formamide peaks for the same reason.

These η values can be used to calculate column densities or upper limits of formamide in a specific mixture and at a specific temperature by simply multiplying the corresponding relative apparent band strength by the appropriate apparent band strength in Table 2.

Table B.1: Relative apparent band strengths of the three selected bands of pure formamide and formamide in various binary and tertiary mixtures at temperatures from 15-212 K.

Matrix	T (K)	Rel. Band Strength C = O stretch + NH ₂ bend 1700.3 cm ⁻¹ (5.881 μm)	Rel. Band Strength CH bend 1388.1 cm ⁻¹ (7.204 μm)	Rel. Band Strength CN stretch 1700.3 cm ⁻¹ (7.529 μm)
Pure		Int. Range: 2050-1532 cm ⁻¹	Int. Range: 1532-1364 cm ⁻¹	Int. Range: 1364-1258 cm ⁻¹
	15	1.00	1.00	1.00
	49	0.99	1.00	0.97
	99	0.96	0.99	0.93
	149	0.95	0.97	0.91
		Int. Range: 2050-1570 cm ⁻¹	Int. Range: 1570-1364 cm ⁻¹	Int. Range: 1364-1258 cm ⁻¹
	170	0.70	1.01	1.015
191	0.70	1.01	1.014	
212	0.60	0.82	0.857	
H ₂ O		Int. Range: 1961-1553 cm ⁻¹	Int. Range: 1532-1368 cm ⁻¹	Int. Range: 1368-1258 cm ⁻¹
	15	1.64	1.63	1.08
	49	1.66	1.67	1.13
	99	1.71	1.61	1.16
	129	1.63	1.57	1.18
	149	1.73	2.22	2.01
161	1.32	2.10	2.00	
CO ₂		Int. Range: 1961-1532 cm ⁻¹	Int. Range: 1532-1364 cm ⁻¹	Int. Range: 1364-1240 cm ⁻¹
	15	1.33	0.85	1.94
	34	1.34	0.84	1.88
	40	1.31	0.78	1.85
	49	1.31	0.79	1.88
	60	1.33	0.83	1.92
90	1.36	0.93	1.91	
CO		Int. Range: 1900-1500 cm ⁻¹	Int. Range: 1532-1364 cm ⁻¹	Int. Range: 1364-1240 cm ⁻¹
	15	1.81	1.30	2.62
	31	1.81	1.34	2.71
34	1.80	1.38	2.72	
CH ₃ OH		Int. Range: 1800-1550 cm ⁻¹	-	-
	15	1.42	-	-
	50	1.41	-	-
	100	1.39	-	-
	121	1.75	-	-
	139	1.70	-	-
154	0.79	-	-	
NH ₃		-	Int. Range: 1484-1374 cm ⁻¹	Int. Range: 1374-1258 cm ⁻¹
	15	-	1.28	1.14
	51	-	1.28	1.11
	60	-	1.28	1.10
	90	-	1.24	1.05
	99	-	1.39	1.13
108	-	1.41	1.15	
H ₂ O:CO ₂		Int. Range: 1961-1553 cm ⁻¹	Int. Range: 1437-1368 cm ⁻¹	Int. Range: 1368-1279 cm ⁻¹
	15	2.20	1.30	0.75
	51	2.18	1.23	0.67
	101	1.89	1.10	0.51
	131	1.85	0.99	0.50
	149	1.88	1.55	1.24
161	1.57	1.42	1.33	
H ₂ O:NH ₃		-	Int. Range: 1437-1368 cm ⁻¹	Int. Range: 1368-1279 cm ⁻¹
	15	-	1.35	1.19
	51	-	1.36	1.21
	99	-	1.32	1.27
	134	-	1.19	1.27
	149	-	1.11	1.24
161	-	1.01	1.22	
CO:CH ₃ OH		Int. Range: 1810-1524 cm ⁻¹	-	-
	15	1.57	-	-
	30	1.64	-	-
35	1.68	-	-	
CO:NH ₃		-	Int. Range: 1460-1370 cm ⁻¹	Int. Range: 1370-1256 cm ⁻¹
	15	-	1.41	2.23
	30	-	1.49	2.14
37	-	1.54	1.85	

Appendix C: QMS calibration of an independent tri-dosing leak valve system and mixing ratio determination

Appendix C.1: Calibration procedure and mixing ratio determination

The new tri-dosing system mentioned in Section 3 allows for simultaneous but independent deposition of gases and vapors via three leak valves, each connected to a separate gas line. Compared to our previous method in which gases and vapors were premixed in the desired ice ratio in a gas bulb and then dosed into the chamber through a single valve, the new method allows for codepositing multiple gases and vapors without experimental errors in the ratio caused by mixing gases with different volatilities in a single bulb or dosing gases that may have different flow, pumping, and substrate deposition rates through the same valve. Subsequently, it greatly improves the ability to create mixtures with precisely determined ratios of molecules with low volatilities like formamide, which is challenging in traditional premixing procedures. The benefits of independent multidosing systems were also described for similar systems with two leak valves in Gerakines et al. (1995) and Yarnall & Hudson (2022).

There are several ways to calibrate such a system to ensure a certain ratio of ice components. One such method is calibrating the deposition rate on the substrate to a specific leak valve position with a specific pressure of the gas or vapor of choice in its manifold line. However, because formamide has a very low vapor pressure compared to liquids like H₂O and CH₃OH and tends to stick to and condense in various parts of the line, reproducing a specific line pressure throughout multiple experiments using this method is difficult. Therefore, to conduct a systematic and thorough IR characterization of formamide in a wide variety of ices with precisely constrained mixing ratios, a different method is necessary.

For this purpose, we calibrate molecules' ice deposition rates with the intensity of their mass signals during the deposition with a QMS. In this calibration procedure, a pure molecule is dosed at a constant rate into the chamber, with the substrate cooled to the desired deposition temperature and the IR spectrometer continuously collecting IR spectra, while the QMS continuously collects mass peak intensity values of selected mass-to-charge ratios (*m/z*) in the selected ion monitoring (SIM) mode. The IR spectrometer is used to measure the ice column density rather than the laser interference because the formamide deposition pressure does not remain stable over the long period of time necessary to generate multiple interference fringes (>18 hours), which is necessary to reliably extract a deposition rate. Conversely, a deposition rate can be extracted from integrated absorbance growth rates (obtained via a least-squares fit to the integrated absorbance over time) in ~30 mins, during which time the formamide deposition rate remains stable (as indicated by the linearity of the integrated absorbance increase over time). The integrated absorbance growth rate for that molecule can then be correlated to a specific mass peak's signal intensity (typically the molecule's base peak) in the QMS (obtained via averaging the mass peak's signal intensity values collected during the deposition and simultaneous IR data collection). The integrated absorbance growth rate can then be converted to the ice column density growth rate, dN/dt , via the following equation if the band strength of the pure molecule, A , is known:

$$\frac{dN}{dt} = \frac{2.303}{A} \times \frac{d \int abs(\nu) d\nu}{dt}. \quad (\text{C.1})$$

Table 1 provides the peak used for the calibration of each pure molecule and its corresponding band strength and reference.

Via this method, a calibration curve relating a mass peak's signal intensity in the QMS to its column density growth rate can be determined, with the slope of this curve referred to here as a molecule's *sensitivity* (see Figure C.1 for an example of such a calibration). When starting a deposition, the leak valve can then be opened accordingly so that the mass signal of the molecule in the QMS corresponds to the desired column density growth rate. In this work, such calibration curves were completed for all molecules used in these spectra with a Spectra Microvision Plus QMS. The relationship between column density growth rate and QMS signal intensity is linear for all molecules within the deposition pressure ranges used (R^2 values of the linear fits ranged from 0.9699-0.9999 with an average of 0.9936).

After the experiment, the mass signal data during the deposition can be converted via the equation from the calibration curve to a column density growth rate, which is then integrated over time to give the absolute column density of each species at the end of the deposition. However, in the case that some of the species in a given mixture share their strongest mass peaks and have no alternative strong peaks without overlap with the other mixture components (which is the case for several mixtures in this work), the individual column density growth rates must be extracted from the mass spectra by utilizing ratios of a given molecule's base peak to another mass peak that is not shared with any other molecules in a given mixture. For example, the mass spectrum of formamide contains a peak at 28, the base peak of CO. Thus, during the deposition of the formamide:CO mixture, the 28 *m/z* signal contains contribution from both formamide and CO. The contribution of formamide to the signal at 28 *m/z* was calculated by dividing the signal at 45 *m/z* (which, in this mixture, only formamide contributed to) by the ratio of the 45 and 28 *m/z* peaks during pure formamide deposition. This calculated contribution was then subtracted from the 28 *m/z* signal to yield the CO 28 *m/z* signal.

In order to estimate the error of the calculated column density of each component and, subsequently, the mixing ratios in each ice, multiple sources of error have been considered. These are discussed in the following subsections.

Appendix C.2: Ion interference effect

Ion interactions within the instrument, such as ion-molecule interactions or ions interacting with the QMS filament or rods, during the dosing of multiple species into the chamber can effect a molecule's sensitivity. Such interactions between two different species can cause their sensitivities to deviate from the values determined in the calibration of each species in pure form. This phenomenon is often referred to as the *ion interference effect*, and it complicates using a mass spectrometer to quantify gases or vapors in a mixture (Basford et al. 1993; Yongjun et al. 2022).

The magnitude of this effect is highly dependent on the species as well as the instrument. It increases with total pressure and decreases for a given species as its proportion in a mixture increases (Nemanič & Žumer 2018; Sun et al. 2020). Thus, the sensitivities that are most affected by this effect are those of species that are present in the lowest proportions in a mixture. Given that our formamide dosing pressure was in the range of a couple 10^{-9} mbar and that the intended ratio of formamide to matrix components was ~5:100 in the case of binary mixtures and 5:100:25 in the case of tertiary mixtures, we treated the interference effect of formamide on the matrix components

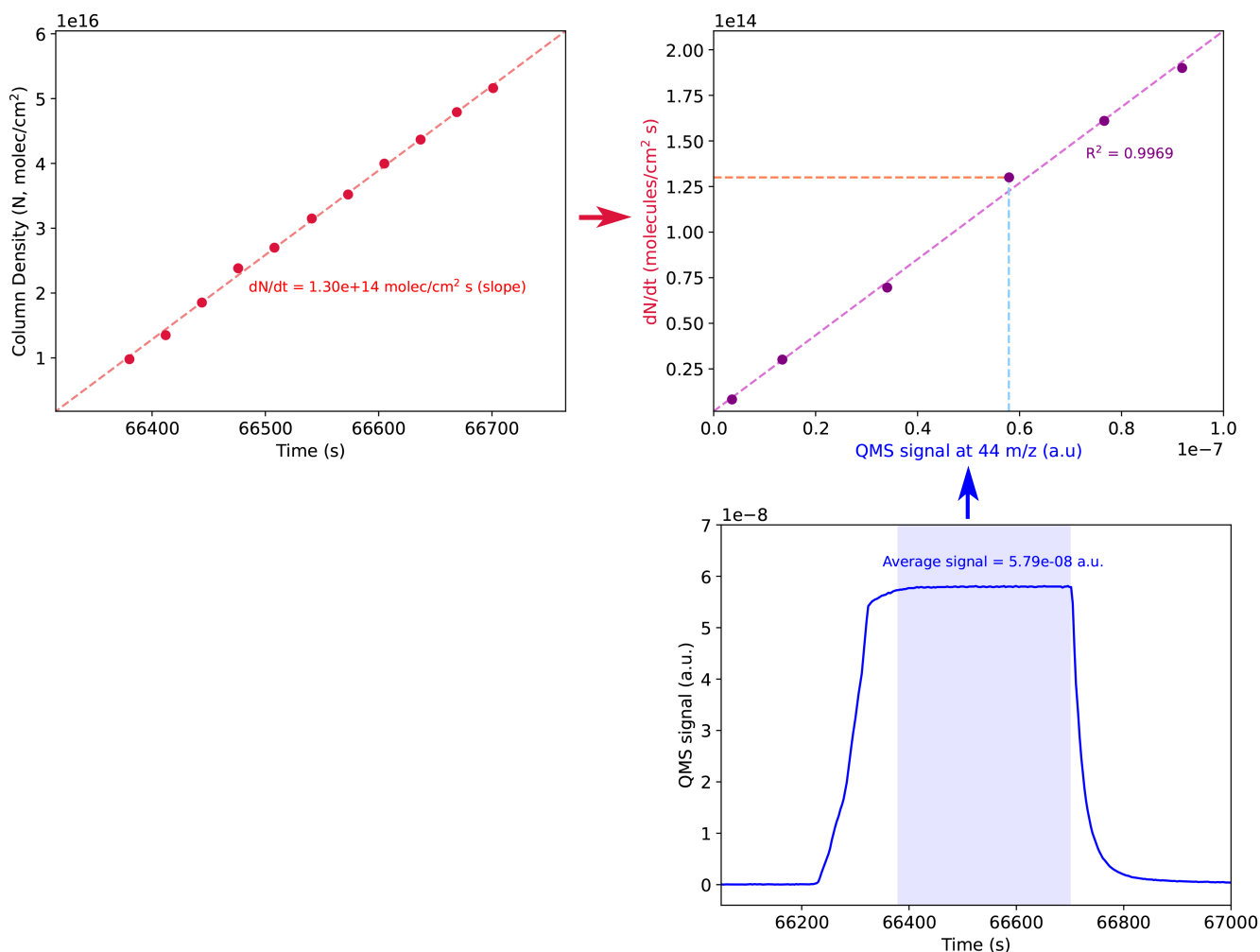


Fig. C.1: Example of how a calibration curve is obtained, shown here for CO. The rate of ice growth is found via monitoring the molecule's integrated absorbance of a selected peak, which can be converted to column densities via known pure band strengths, and relating them to the time of the spectrum collected (top left). Then the average value of a given mass signal (28 m/z for CO), measured by the QMS during the same time interval as the collection of the IR spectra, is calculated (bottom right). The calibration curve is obtained by correlating the two values to each other for a range of ice deposition rates (top right).

as negligible and accounted for ion interference only in the formamide signal. While the formamide absolute column densities are necessary to calculate its relative band strengths (see Section 3), the absolute column densities of the matrix components are not needed to find any values other than the mixing ratios.

In order to quantify the ion interference effect on formamide in each mixture, at the start of each deposition, formamide was first dosed alone, and its mass signal was given ~ 5 min to stabilize before the other matrix components were introduced into the chamber. Although this meant that each experiment started with a very brief deposition of pure formamide, the deposition rate of formamide was so slow in all of the experiments (on the order of tens of monolayers per hour) that this brief pure deposition was usually not even noticeable above the noise level in the IR spectra. Then, the ratio between formamide's signal before and after the matrix molecules were added to the chamber was used as a correction factor to remove the ion interference effect from formamide's signal. An example of this correction is shown in

Figure C.2 for the formamide:CH₃OH mixture, which had the highest correction factor of all the mixtures (1.11).

The ion interference effect on formamide was noticeable in all of the mixtures where the major matrix component was polar, while it was not detected above the noise level in the mixtures in which the major matrix component was apolar. In order to provide a conservative estimate of the error caused by the ion interference effect on the calculated column density of formamide, the percent difference of the formamide column density before and after the ion interference effect correction was obtained for all mixtures in which the effect was detected above the noise level. The average percent difference was $\sim 5\%$, with the highest percent difference being that of the formamide:CH₃OH mixture ($\sim 10\%$). To avoid underestimating the error in any of the mixtures, we use this maximum error, $\sim 10\%$, as the uncertainty in the column density caused by the ion interference effect.

The sensitivity of a QMS can also drift over time. However, such drift is typically only significant over timescales spanning several months to a couple years, and it is more significant for absolute sensitivities than for relative sensitivities (Ellefson et al.

Table C.1: Peaks and their band strengths used to obtain calibration curves of the pure molecules.

Molecule	Calibration peak (cm^{-1} μm)	Band strength used (10^{-17} cm molec $^{-1}$)
NH ₂ CHO	1701 5.88	6.4 (this work)
H ₂ O*	3280 3.05	20. (Bouilloud et al. 2015)
	1660 6.02	1.1 (Bouilloud et al. 2015)
CO ₂ *	2343 4.27	13. (Bouilloud et al. 2015)
	665 15.3	1.8 (Bouilloud et al. 2015)
CO	2139 4.67	0.87 (González Díaz et al. 2022)
CH ₃ OH*	1027 9.74	1.61 (Luna et al. 2018)
	1460 6.85	1.1 (Luna et al. 2018)
NH ₃	1071 9.34	1.95 (Hudson et al. 2022)

* If two calibration peaks are listed, then both peaks were utilized to create the calibration curve, with the peak with a lower band strength used to calculate the ice's column density at the higher deposition rates to avoid saturation effects in the calibration curve.

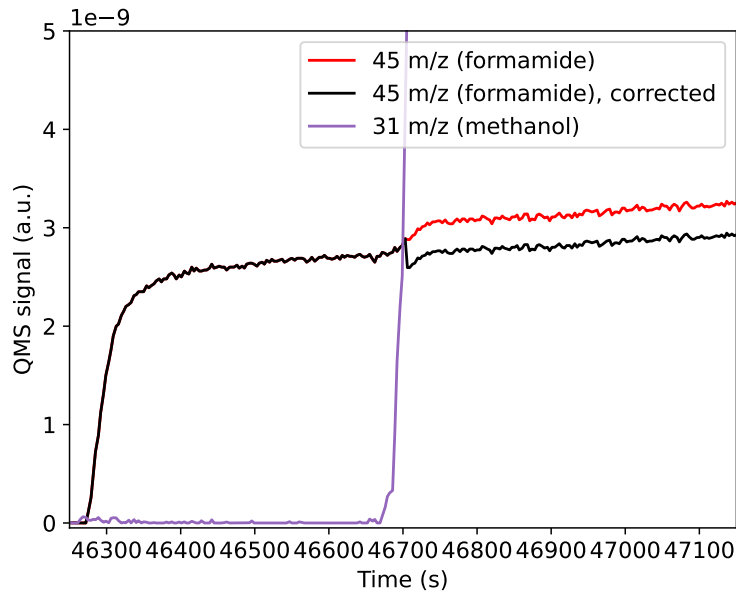


Fig. C.2: Correction of the ion interference effect observed in the formamide 45 m/z signal during CH₃OH codeposition.

1987; Lieszkowszky et al. 1990). The contribution of this drift to the method error was assumed to be negligible here given that all the formamide mixture spectra were collected within a span of two months, and that the calibration curves were usually either determined within a few days of their use to create an ice mixture or were frequently updated with new values that were consistent with the fits to the previous values.

Appendix C.3: Error calculation

In this method of determining ice column densities, multiple sources of error must be considered. First, there is the error in the method used to determine the ice column density growth rate (dN/dt). This error can be estimated by propagating uncertainties of the variables in Equation C.1. For all ices, the integrated absorbance growth rate uncertainty is estimated to be 10%, as mentioned in Section 3. For formamide, the uncertainty in the band strengths reported in this work is estimated to be 15 (also see Section 3). For the matrix components, literature band strength values were used (see Table C.1). However, in the

literature, variations between the band strengths reported in different publications can be large (e.g., Bouilloud et al. 2015). For this reason, we estimate a 25% uncertainty for the literature band strengths used.

Then, we determine the uncertainty resulting from converting the integrated QMS measurement to a column density experimentally by finding the difference between the column density calculated from the QMS signal and the column density calculated from the integrated IR absorbance at the end of a pure molecule's IR measurement. Comparing these differences in formamide, H₂O, and CO deposition experiments resulted in an average error $\sim 2.5\%$. To be conservative, we estimate the error from converting the QMS measurement to a column density to be 5%.

Propagating all of these uncertainties, along with the 10% uncertainty from the ion interference effect for the formamide measurements, results in an uncertainty of $\sim 21\%$ for the formamide column densities and $\sim 27\%$ for the matrix column densities in the ice mixtures.