

Infrared spectroscopy of HCOOH in interstellar ice analogues[★]

S. E. Bisschop¹, G. W. Fuchs¹, A. C. A. Boogert², E. F. van Dishoeck¹, and H. Linnartz¹

¹ Raymond and Beverly Sackler Laboratory for Astrophysics, Leiden Observatory, Leiden University, PO Box 9513, 2300 RA Leiden, The Netherlands

e-mail: bisschop@strw.leidenuniv.nl

² AURA/NOAO-South, Casilla 603, La Serena, Chili

Received 13 March 2007 / Accepted 24 April 2007

ABSTRACT

Context. HCOOH is one of the more common species in interstellar ices with abundances of 1–5% with respect to solid H₂O. With the launch of the Spitzer Space Telescope new infrared spectra have become available of interstellar ices in different environments. So far systematic laboratory studies on HCOOH-containing interstellar ice analogues are lacking.

Aims. This study aims at characterizing the HCOOH spectral features in astrophysically relevant ice mixtures in order to interpret astronomical data.

Methods. The ices are grown under high vacuum conditions and spectra are recorded in transmission using a Fourier transform infrared spectrometer. Pure HCOOH ices deposited at 15 K and 145 K are studied, as well as binary and tertiary mixtures containing H₂O, CO, CO₂ and CH₃OH. The mixture concentrations are varied from 50:50% to ~10:90% for HCOOH:H₂O. Binary mixtures of HCOOH:X and tertiary mixtures of HCOOH:H₂O:X with X = CO, CO₂, and CH₃OH, are studied for concentrations of ~10:90% and ~7:67:26%, respectively.

Results. Pure HCOOH ice spectra show broad bands which split around 120 K due to the conversion of a dimer to a chain-structure. Broad single component bands are found for mixtures with H₂O. Additional spectral components are present in mixtures with CO, CO₂ and CH₃OH. The resulting peak position, full width at half maximum and band strength depend strongly on ice structure, temperature, matrix constituents and the HCOOH concentration. Comparison of the solid HCOOH 5.9, 7.2, and 8.1 μm features with astronomical data toward the low mass source HH 46 and high mass source W 33A shows that spectra of binary mixtures do not reproduce the observed ice features. However, our tertiary mixtures especially with CH₃OH match the astronomical data very well. Thus interstellar HCOOH is most likely present in tertiary or more complex mixtures with H₂O, CH₃OH and potentially also CO or CO₂, providing constraints on its formation.

Key words. astrochemistry – line: profiles – molecular data – molecular processes – methods: laboratory – ISM: molecules

1. Introduction

Formic acid is one of the few molecules in star forming regions that is detected both in the solid state and in the gas phase (see e.g., Schutte et al. 1999; Ikeda et al. 2001). Infrared observations of ices in high mass Young Stellar Objects (YSOs) with the Infrared Space Observatory (ISO) contain absorption features that are assigned to HCOOH and that correspond to the C=O stretch, $\nu_S(\text{C}=\text{O})$, at 5.9 μm and the OH and CH bending modes, $\nu_B(\text{OH})$ and $\nu_B(\text{CH})$ at 7.2 μm (Schutte et al. 1997, 1999). The $\nu_S(\text{C}=\text{O})$ band is the strongest spectral feature of HCOOH ice, but suffers from partial overlap with the OH bending mode of solid H₂O at 6 μm, $\nu_B(\text{OH})$. The $\nu_B(\text{OH/CH})$ band at 7.2 μm is significantly weaker than the $\nu_S(\text{C}=\text{O})$ band, but is present in a relatively clean region and has previously been detected toward high mass star forming regions (Schutte et al. 1999). More recently, these features have also been observed toward low mass YSOs and possibly toward background stars (Boogert et al. in prep.; Knez et al. 2005). Typical abundances of HCOOH vary between individual sources from 1 to 5% with respect to solid H₂O. The aim of this paper is to study the $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH/CH})$ and other infrared features of HCOOH ices in

different astrophysically relevant laboratory matrices and to compare the results with astronomical spectra to infer the chemical environment of HCOOH in interstellar ices.

In the gas phase HCOOH is detected in envelopes surrounding high and low mass star forming regions, as well as in shocked regions toward the galactic center (van Dishoeck et al. 1995; Ikeda et al. 2001; Bottinelli et al. 2007; Requena-Torres et al. 2006; Bisschop et al. 2007). HCOOH is part of the surface chemistry network leading to the formation of complex organic molecules seen in hot cores (Tielens & Charnley 1997) and is therefore expected to be present in compact hot regions where it has evaporated from dust grains. In contrast, interferometric observations show that gas phase HCOOH emission can be extended beyond the hot core and is not coincident with other oxygen-bearing species such as CH₃OCHO (Liu et al. 2002; Hollis et al. 2003; Remijan & Hollis 2006). The observed gas phase HCOOH has low rotation temperatures and its abundance is a factor of 10⁻⁴ lower than that of solid HCOOH (Bisschop et al. 2007). It is very likely that gaseous and solid HCOOH co-exist in the same region. A low level of non-thermal desorption due to e.g., cosmic ray spot heating is sufficient to explain the observed gas phase abundances of HCOOH.

Previously, laboratory ice spectra have been studied for pure HCOOH and mixtures of HCOOH with H₂O and/or CH₃OH

[★] Appendix A is only available in electronic form at <http://www.aanda.org>

by Schutte et al. (1999) and Hudson & Moore (1999). These studies only focused on specific spectral features. The data clearly showed that the morphology and composition of the ice matrix strongly influence the HCOOH infrared spectrum. Recently, experiments have been reported for the layered ice system HCOOH/H₂O in which the interaction of HCOOH and H₂O in the interface is studied (see e.g., Cyriac & Pradeep 2005; Bahr et al. 2005; Souda 2006).

There is a clear need to understand the spectroscopy of the main ice constituents such as HCOOH, as new observational tools become available for the study of infrared features of interstellar ices, such as the Spitzer Space Telescope, the Stratospheric Observatory For Infrared Astronomy (SOFIA), and the James Webb Space Telescope (JWST). Solid HCOOH is likely mixed with species such as H₂O, CO, CO₂ and CH₃OH that are abundant in the ice and have related formation mechanisms. Infrared ice features of mixtures with these species have been studied in this paper. Pure HCOOH is measured as well for comparison. Quantitative and qualitative analysis of the 5.9 μm , 7.2 μm and other HCOOH ice features allows for a better determination of the amount of HCOOH present in interstellar ices as well as its ice environment.

The paper is organized as follows: Sect. 2 explains the experimental method, Sect. 3 the analysis techniques, Sect. 4 discusses the results of experiments with pure ices, as well as mixtures with H₂O, CH₃OH, CO and CO₂, in Sect. 5 the effects of ice matrix and temperature on the band strength are presented, Sect. 6 discusses the physical aspects of the interactions of HCOOH in the solid state with other species, Sect. 7 presents the astrophysical implications, and finally Sect. 8 summarizes the main conclusions of this paper.

2. Experimental procedure

The experiments are performed in a high vacuum (HV) set-up with a base pressure of $\sim 10^{-7}$ Torr, that has been previously described in detail by Gerakines et al. (1995). A CsI window mounted in close thermal contact with a closed cycle He cryostat is situated at the center of the chamber. It is used as a substrate holder and can be cooled down to 14 K. The sample temperature is controlled to better than ± 0.1 K using the cryostat cold finger, a resistive heating element and a Lakeshore 330 temperature control unit. The system temperature is monitored by two KP-type thermocouples (0.07% Au in Fe versus chromel), one mounted on the substrate face and the second close to the heater element. Ices are grown in situ onto the substrate, by exposing the cold surface to a steady flow of gas that is introduced into the chamber via an all metal flow control valve with a modified outlet directed at the substrate center along the surface normal. The ices are monitored by Fourier Transform Infrared (FTIR) absorption spectroscopy with 1 cm^{-1} spectral resolution covering the 4000–500 cm^{-1} range.

To test the effect of dilution of HCOOH (98% purity, J. T. Baker) by H₂O (deionized) four mixtures are prepared in a glass vacuum manifold. Both HCOOH and H₂O are further purified by subsequently freezing and thawing the samples in the vacuum manifold. The mixture ratios range from the relative abundances derived from observations of interstellar ices of 1 to 5% (see e.g., Gibb et al. 2004) to more HCOOH-rich ices with relative HCOOH:H₂O concentrations up to 50:50%. The astrophysical motivation for these higher mixture ratios is that other species such as CH₃OH are present in solid state environments with nearly equal amounts of H₂O and/or CO₂ (Ehrenfreund et al. 1998, 1999; Dartois et al. 1999). The

Table 1. Overview of ice morphologies, total ice exposure and deposition time.

H ₂ O (%)	HCOOH (%)	X (%)	Total exposure (L)	Deposition time (min)
–	100	–	900	15
100	–	–	900	15
50	50	–	900	15
66	34	–	900	15
80	20	–	900	15
91	9	–	900	15
–	11	89(CO)	1800	30
–	9	91(C ¹⁸ O ₂)	900	15
–	10	90(CH ₃ OH)	1800	30
62	8	30(CO)	1800	30
67	6	27(C ¹⁸ O ₂)	1800	30
68	6	26(CH ₃ OH)	1800	30

different HCOOH:H₂O mixtures that are used here are summarized in Table 1. A pure H₂O sample is prepared for comparison. Furthermore, some features of the solid CO band can be explained with mixtures or layers of CO with non-hydrogen-bonding molecules or with hydrogen-bonding species other than H₂O (e.g., Tielens et al. 1991; Pontoppidan et al. 2003; Fraser et al., in prep.). Accordingly, we have also studied a set of binary mixtures of HCOOH:X and tertiary mixtures of HCOOH:H₂O:X, where X stands for CO (99.997% purity, Praxair), C¹⁸O₂ (97% purity, Praxair) or CH₃OH (99.9% purity, Janssen Chimica) are prepared with concentrations of $\sim 10:90\%$ and $\sim 7:67:26\%$, respectively. These mixtures are chosen such that the matrix species determine the structure of the ice. Uncertainties on the concentrations of the ice constituents are $\sim 10\%$ due to inaccuracies in the pressure reading. The C¹⁸O₂ isotopologue is used to distinguish solid C¹⁸O₂ from regular atmospheric gas phase CO₂.

A typical experiment starts with taking a background spectrum at 15 K. Subsequently, the ices are deposited with a flow of $\sim 1.0 \times 10^{-5}$ Torr s⁻¹. An exposure of 1.0×10^{-6} Torr s⁻¹ corresponds to 1 Langmuir (L) or equivalently 1 monolayer (ML) s⁻¹ assuming that the molecular surface density is 10^{15} molecules cm⁻² and the sticking probability is 1 at 15 K. The deposition times and total exposures are summarized in Table 1. After deposition, FTIR spectra are taken at 15 K, and subsequently every 15 K for binary mixtures of HCOOH with H₂O or CO₂ and every 30 K for all other mixtures up until a temperature of 165 K is reached. At each temperature, the sample stabilizes for 20 min before a spectrum is recorded.

One pure HCOOH experiment is performed where HCOOH is deposited at 145 K. The aim of this experiment is to determine whether the HCOOH ice structure depends on the deposition temperature and whether potential phase changes are reversible. Subsequently, the experiment is cooled down to 15 K and from then on FTIR spectra are taken every 15 K during warm-up.

3. Data reduction and analysis

The spectral range from 4000–500 cm^{-1} is very rich in absorption features of complex organic species, and HCOOH is no exception. Consequently special care has to be taken in the data reduction. Depending on the mixture, different frequency ranges are used to fit third order polynomial baselines (see Table 2). An additional local baseline is subtracted in the 1800–1100 cm^{-1} frequency range. This is necessary as the HCOOH $\nu_5(\text{C}=\text{O})$ mode around ~ 1700 cm^{-1} (5.9 μm) overlaps

Table 2. Frequency ranges used for baseline subtraction.

Ice composition	Baseline range (cm ⁻¹)
HCOOH & HCOOH:H ₂ O	4000–3800, 1850–1800, 1050–1030, 540–500
HCOOH:CH ₃ OH & HCOOH:H ₂ O:X ^a	4000–3700, 1850–1800, 1580–1500, 950–900, 500–450
HCOOH:CO	4000–3800, 2440–2260, 2000–1800, 1550–1400, 600–500
HCOOH:CO ₂	4000–3900, 2100–1800, 1550–1430, 550–500

^a X = CO, C¹⁸O₂ or CH₃OH.

Table 3. Summary of absolute band strengths for HCOOH from the literature.

Ice matrix	A[ν _S (C=O)] (cm molecule ⁻¹)	A[ν _S (C–O)] (cm molecule ⁻¹)	A[ν _B (OH/CH)] (cm molecules ⁻¹)
HCOOH ^a	6.7(–17)	–	2.8(–18)
HCOOH:H ₂ O 7:93% ^b	–	1.5(–17)	–
HCOOH–5H ₂ O ^c	8.5(–17)	5.0(–17)	1.1(–18)
HCOOH–6H ₂ O ^c	7.7(–17)	4.3(–17)	3.5(–18)

^a Laboratory gas phase data by Marechal (1987).

^b Laboratory solid state data by Hudson & Moore (1999).

^c Theoretical calculations by Park & Woon (2006).

with the ν_B(OH) mode of H₂O at 1655 cm⁻¹ (6.0 μm). This H₂O feature is much weaker than the HCOOH feature, even when only 10% of HCOOH is present, but it is responsible for a wing on the ν_S(C=O) band. Finally, a correction is made on the ν_B(OH/CH) feature around 1390 cm⁻¹ (7.2 μm) as it is close to a variable background feature. Spectra that comprise only the background feature are subtracted from the sample spectra for each temperature, but due to its variability, the subtraction does not fully remove the artifact.

The absolute band strengths are difficult to calibrate in our experiment, because the number of molecules that stick to the substrate sample is not exactly known. Instead, we focus in this paper on the relative band strengths that can be accurately measured. The relative values discussed here can be combined with previously measured values for the band strengths available from the literature (summarized in Table 3). The absolute band strengths for the vibrational modes depend on whether HCOOH is in the gas phase or solid state and consequently the numbers in the table do not perfectly match.

To determine the relative band strengths, the ν_S(C=O), ν_B(OH/CH), and ν_S(C–O) features of HCOOH as well as the ν_B(OH) and libration (ν(lib)) bands of H₂O are integrated over the frequency ranges given in Table 4. The integration boundaries for H₂O are identical to those used recently by Öberg et al. (2007). Since the ν_B(OH) band at 930 cm⁻¹ (10.8 μm) and the OCO bending mode (ν_B(OCO)) at 705 cm⁻¹ (14.2 μm) of HCOOH overlap with the H₂O ν(lib) band at 750 cm⁻¹ (13.3 μm), these two features are subtracted before integration.

The relative band strengths of the HCOOH and H₂O spectral features are calculated with respect to pure HCOOH and H₂O via:

$$\frac{A(\text{HCOOH:H}_2\text{O} = Y : 1)}{A_0(\text{HCOOH})} = Y \times \frac{\int_{\text{band}} I(\text{HCOOH:H}_2\text{O} = Y : 1)}{\int I_0(\text{HCOOH})}. \quad (1)$$

Here Y is the HCOOH ice fraction, $\int_{\text{band}} I(\text{HCOOH:H}_2\text{O} = Y : 1)$ the integrated absorbance of the band in the mixture, $A_0(\text{HCOOH})$ the band strength for pure HCOOH ice at 15 K, and $\int I_0(\text{HCOOH})$ the integrated absorbance for pure HCOOH ice at 15 K. A similar formula is used for H₂O. The resulting

Table 4. Frequency ranges used for integration bounds.

Vibrational mode	Wavenumber (cm ⁻¹)
HCOOH	
ν _S (C=O)	1800–1550
ν _B (OH/CH)	1450–1370
ν _S (C–O)	1300–1150
H ₂ O	
ν _S (OH)	3700–3000
ν(lib)	1000–500

relative band strengths are studied as a function of temperature (see Sect. 5). An extensive discussion of the uncertainties of the band strengths is given by Öberg et al. (2007). In short, the largest fraction of the uncertainties arises from inaccuracies on the mixing ratio of the ice and yields relative uncertainties that are below 10%. Additionally, baseline subtraction and deposition time inaccuracies are 1 to 2%. A conservative approach yields an overall uncertainty of ~12% in the resulting calculations for the relative band strengths. Furthermore, it is important to note that the relative uncertainty between different temperatures for a single experiment amounts to ~2% as only baseline inaccuracies contribute. Observed temperature trends are thus much more certain than concentration dependencies.

The ν_S(C=O), ν_B(OH/CH) and ν_S(C–O) bands of HCOOH are fitted with Gaussian line profiles for the astrophysically most relevant 1800–1100 cm⁻¹ range. To fit the asymmetric profile of the ν_S(C=O) stretch around 1700 cm⁻¹ two components are needed. It turned out to be difficult to disentangle the properties of the absorption bands of HCOOH in some mixtures, in particular those containing CO and CO₂. In such cases average peak positions are given. In general the uncertainties of peak positions and FWHMs are ±1 cm⁻¹ and ±2 cm⁻¹, except for the ν_B(OH/CH) feature in the 9:91% HCOOH:H₂O mixture, where the values are ±2 cm⁻¹ and ±4 cm⁻¹, respectively.

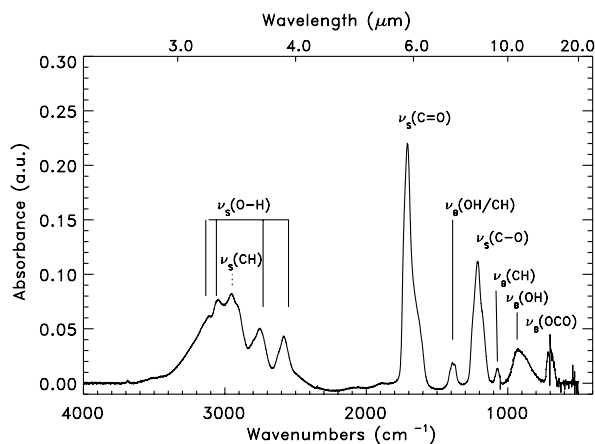


Fig. 1. Experimental infrared spectrum of pure HCOOH ice deposited at 15 K. The $\nu_s(\text{OH})$ band has several sub-maxima in the $\sim 3000\text{ cm}^{-1}$ range where $\nu_s(\text{CH})$ is also located. Below 2000 cm^{-1} the following absorption features are present: the C=O stretching mode $\nu_s(\text{C}=\text{O})$, the C–O stretch $\nu_s(\text{C}-\text{O})$, the CH bending mode $\nu_B(\text{CH})$, the OH bending mode $\nu_B(\text{OH})$ and OCO bending mode $\nu_B(\text{OCO})$. Peak positions and labels are given in Table 5. In this study the focus is on the $\nu_s(\text{C}=\text{O})$, $\nu_B(\text{OH/CH})$, and $\nu_s(\text{C}-\text{O})$ bands.

4. Results

The following section summarizes the results for the pure and mixed HCOOH ices at different temperatures. All spectra are available at <http://www.strw.leidenuniv.nl/~lab/databases/>

4.1. Pure HCOOH ices

Clearly, the pure HCOOH ice spectrum shown in Fig. 1 for 15 K is complex. For completeness we summarize the spectroscopic assignments for all bands here (see Table 5), but in this paper we mainly focus on the astrophysically relevant features $\nu_s(\text{C}=\text{O})$, $\nu_B(\text{OH/CH})$, as well as $\nu_s(\text{C}-\text{O})$. The measured peak positions and *FWHMs* in the astrophysically relevant $1800\text{--}1100\text{ cm}^{-1}$ range are given in Table A.1 of the online material. The features at 3115 cm^{-1} , 2754 cm^{-1} , and 2582 cm^{-1} ($3.21\text{--}3.87\text{ }\mu\text{m}$) are all due to the OH stretching mode, $\nu_s(\text{OH})$. The broad structure with different maxima of these modes are assigned to a combination of dimers and HCOOH organized in long chains (Cyriac & Pradeep 2005). The $\nu_s(\text{OH})$ bands overlap with the 2953 cm^{-1} CH stretch, $\nu_s(\text{CH})$. In observations these bands overlap with and are dominated by the $3\text{ }\mu\text{m}$ H₂O band. The astrophysically important $\nu_s(\text{C}=\text{O})$ band has its maximum around 1714 cm^{-1} and is broad and asymmetric. The C–O stretch, $\nu_s(\text{C}-\text{O})$, is located at 1211 cm^{-1} . Furthermore, bending modes are present of which $\nu_B(\text{CH})$ and $\nu_B(\text{OH})$ at 1073 cm^{-1} and 930 cm^{-1} (9.3 and $10.8\text{ }\mu\text{m}$), respectively, are out-of-plane.

The temperature evolution of the astrophysically relevant HCOOH features in the $1800\text{--}1100\text{ cm}^{-1}$ range is shown in Fig. 2a and the corresponding peak positions are given in Table A.1 of the online material. Most features are broad and consist of multiple components at 15 K. At 135 K the $\nu_s(\text{C}=\text{O})$, $\nu_s(\text{C}-\text{O})$, and $\nu_B(\text{OH/CH})$ modes at ~ 1700 , 1250 , and 1390 cm^{-1} split into two components. The splitting of the $\nu_s(\text{C}=\text{O})$ and $\nu_s(\text{C}-\text{O})$ bands is assigned to the out-of-phase and in-phase motions of different HCOOH molecules within the same chain (Cyriac & Pradeep 2005). The out-of-phase motions are located at higher wavenumbers than the in-phase

motions. The $\nu_s(\text{OH})$ and $\nu_s(\text{CH})$ in-plane bending modes overlap at 15 K, but are clearly observed as two separate peaks at higher temperatures. Our results are consistent with those of Cyriac & Pradeep (2005), who assigned the splitting to a conversion of HCOOH dimers to HCOOH organized in chains, following an experimental study by Millikan & Pitzer (1958) and calculations by Yokoyama et al. (1991).

In one experiment HCOOH is deposited at 145 K (see Fig. 2b). After deposition the spectrum is similar to that of HCOOH deposited at 15 K and subsequently heated to $\geq 135\text{ K}$. When the temperature is lowered to 15 K, the general features remain similar to the high temperature spectrum. Upon subsequent heating to 150 K an exact copy of the original (145 K) spectrum is obtained. Thus, the changes in ice matrix structure after heating are irreversible.

4.2. HCOOH:H₂O ices

Experiments with mixtures that range from 50:50% to 9:91% HCOOH:H₂O have been performed. In Fig. 2c, the temperature evolution of the mixture with 20% of HCOOH is shown. Peak positions and *FWHMs* for the $1800\text{--}1100\text{ cm}^{-1}$ range are given in Table A.2 of the online material. The dilution of HCOOH in H₂O clearly affects the band profiles of HCOOH. The $\nu_s(\text{C}=\text{O})$ band shifts 3 cm^{-1} to the red at 15 K and the $\nu_B(\text{OH/CH})$ feature is narrower than in pure HCOOH. Between 105 and 120 K both the peak intensity and *FWHM* of the $\nu_B(\text{OH/CH})$ band increase. At 150 K the H₂O features change due to a phase transition from amorphous to crystalline H₂O ice structure. At the same time the $\nu_s(\text{C}=\text{O})$, $\nu_B(\text{OH/CH})$, and $\nu_s(\text{C}-\text{O})$ modes of HCOOH split into two components. Finally, at 165 K all H₂O and most HCOOH have desorbed.

The spectra for different mixture concentrations with respect to H₂O are shown in Fig. 3 for 15 K (left) and 150 K (right). The $\nu_s(\text{C}=\text{O})$ band has an additional component at 15 K in HCOOH:H₂O mixtures with high HCOOH concentrations that is located at lower wavenumbers and is due to the in-phase motion between HCOOH molecules in the same chain. This implies that there is a small fraction of HCOOH ice that is in a bulk HCOOH environment. The pure HCOOH component decreases with increasing H₂O concentration. The $\nu_B(\text{OH/CH})$ and $\nu_s(\text{C}-\text{O})$ modes consist of one component and their *FWHMs* decrease for lower HCOOH concentrations.

At 150 K, the $\nu_s(\text{C}=\text{O})$, $\nu_B(\text{OH/CH})$, and $\nu_s(\text{C}-\text{O})$ modes are split into two peaks for the three mixtures with the highest HCOOH concentrations. The $\nu_B(\text{OH/CH})$ mode is broader compared to pure HCOOH. At the lowest HCOOH concentration, no splitting of the $\nu_s(\text{C}=\text{O})$ spectral feature of HCOOH is observed, but only a broad band with a *FWHM* of 67 cm^{-1} . Furthermore, the $\nu_s(\text{C}=\text{O})$ band is shifted to 1718 cm^{-1} and the $\nu_B(\text{OH/CH})$ feature is located at 1404 cm^{-1} .

4.3. HCOOH:CO, HCOOH:CO₂, and HCOOH:CH₃OH ices

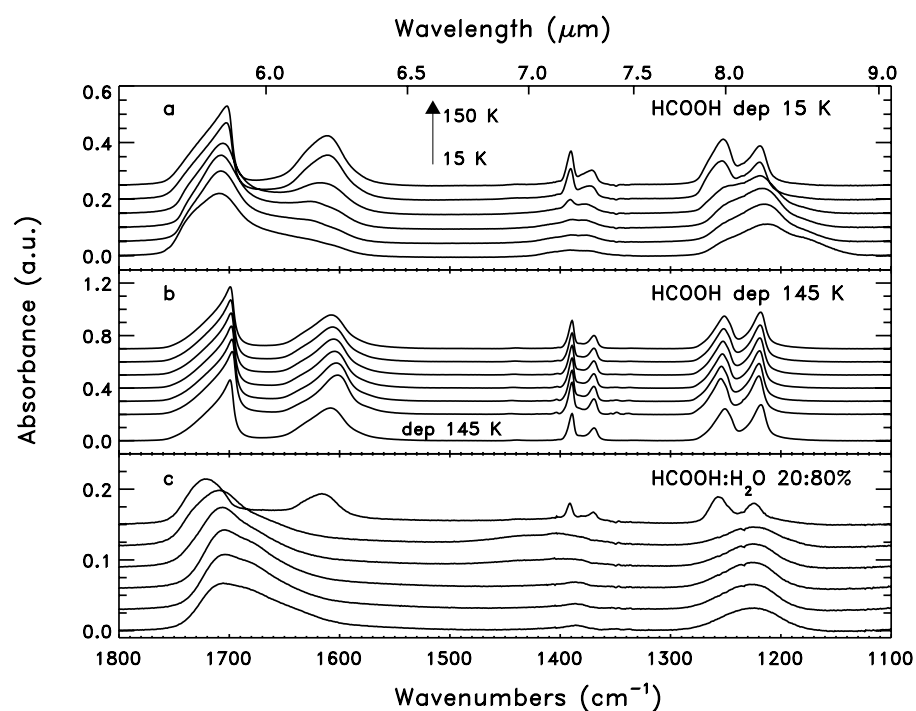
Figure 4 shows the effects different species have on the solid state HCOOH infrared spectrum. Tables A.3 and A.4 of the online material summarize the profile parameters. HCOOH mixed with H₂O results in a relatively simple spectrum, but mixtures with other species, especially CO or CO₂, show multiple components. In mixtures with CO or CO₂ the $\nu_s(\text{C}=\text{O})$ band is shifted to higher wavenumbers of which the two most prominent features are located at 1717 cm^{-1} and 1735 cm^{-1} in both mixtures. A feature is present at 1154 cm^{-1} for the $\nu_s(\text{C}-\text{O})$ band

Table 5. Solid state spectral features for HCOOH, H₂O, CO, CO₂, and CH₃OH at 15 K. The symbols ν_S and ν_B indicate a stretching or bending mode, respectively. The gas phase spectroscopic label is indicated as well. The astrophysically most relevant features of HCOOH are marked in bold face.

Wavenumber (cm ⁻¹)	Wavelength (μ m)	Mode	Label	Comment
HCOOH				
3115	3.21	$\nu_S(\text{OH})$	ν_1	has several sub-maxima
2953	3.39	$\nu_S(\text{CH})$	ν_2	
2754	3.63	$\nu_S(\text{OH})$	ν_1	has several sub-maxima
2582	3.87	$\nu_S(\text{OH})$	ν_1	has several sub-maxima
1714	5.83	$\nu_S(\text{C=O})$	ν_3	out-of-phase between adjacent chains
1650	6.06	$\nu_S(\text{C=O})$	ν_3	in-phase between adjacent chains
1387	7.21	$\nu_B(\text{OH})$ & $\nu_B(\text{CH})$	ν_5 and ν_4	
1211	8.26	$\nu_S(\text{C-O})$	ν_6	
1073	9.32	$\nu_B(\text{CH})$	ν_8	out-of-plane
930	10.75	$\nu_B(\text{OH})$	ν_5	out-of-plane
705	14.18	$\nu_B(\text{OCO})$	ν_7	
H₂O				
3280	3.05	$\nu_S(\text{OH})$	ν_1, ν_3^a	
1655	5.88	$\nu_B(\text{OH})$	ν_2	
750	13.33	$\nu(\text{lib})$		
CO				
2140	4.67	$\nu_S(\text{CO})$	ν_1	
C¹⁸O₂				
3671	2.72	overtone	$\nu_1 + \nu_3$	
3513	2.85	overtone	$2\nu_2 + \nu_3$	
2310	4.33	$\nu_S(\text{CO})$	ν_3	
645	15.50	$\nu_B(\text{OCO})$	ν_2	
CH₃OH				
3251	3.08	$\nu_S(\text{OH})$	ν_1	
2951	3.39	$\nu_S(\text{CH})$	ν_2, ν_9	
2827	3.54	$\nu_S(\text{CH})$	ν_3	
1460	6.85	CH ₃ deformation and $\nu_B(\text{OH})$	$\nu_4, \nu_{10}, \nu_5, \nu_6^b$	
1130	8.85	CH ₃ rock	ν_7, ν_{11}	
1026	9.75	$\nu_S(\text{CO})$	ν_8	
694	14.41	torsion	ν_{12}	

^a ν_1 refers to the symmetric stretching mode and ν_3 to the anti-symmetric stretching mode of H₂O.

^b ν_4 and ν_{10} are anti-symmetric deformations, ν_5 is a symmetric deformation of CH₃OH and ν_6 is the OH bending mode for CH₃OH.

**Fig. 2.** The temperature dependent spectra of **a)** pure HCOOH ice deposited at 15 K, **b)** pure HCOOH ice deposited at 145 K, subsequently cooled down to 15 K, and **c)** HCOOH:H₂O 20:80% mixed ice. The spectra of the ices are shown for 15, 60, 90, 105, 135, and 150 K. The $\nu_B(\text{OH})$ band of H₂O at 1655 cm⁻¹ has been removed from the spectra for the HCOOH:H₂O 20:80% mixture.

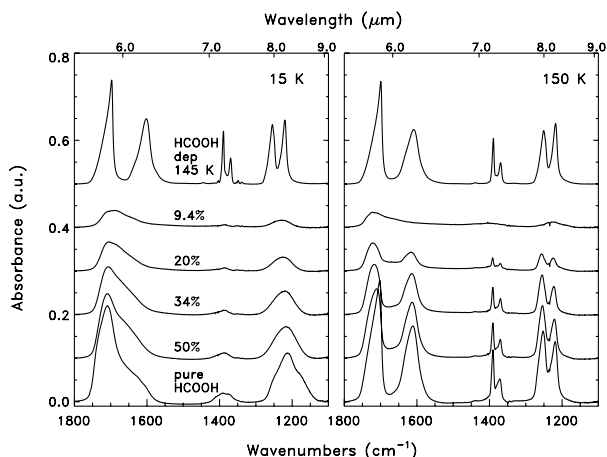


Fig. 3. Absorption spectra for temperatures of 15 K (*left*) and 150 K (*right*) in the 1800–1100 cm^{-1} range for HCOOH deposited at 15 K, the HCOOH:H₂O mixtures and HCOOH deposited at 145 K. The $\nu_{\text{B}}(\text{OH})$ band of H₂O has been removed from the spectra.

which is assigned to monomeric HCOOH. This band is especially strong in the HCOOH:CO₂ mixture. In CH₃OH, the absorption bands more closely resemble the overall features of mixtures with H₂O. However, the $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ modes are split into two components at 1721 cm^{-1} and 1691 cm^{-1} . The $\nu_{\text{B}}(\text{OH}/\text{CH})$ band cannot be compared to that in the other mixtures as it overlaps with and is dominated by the CH₃ deformation and $\nu_{\text{B}}(\text{OH})$ features of CH₃OH at 1460 cm^{-1} .

The temperature behavior of the spectral features of HCOOH ice depends strongly on the matrix species. In mixtures with H₂O, the HCOOH bands are similar for all temperatures. This is not surprising as the H₂O ice structure does not change until 135 K. HCOOH behaves differently in mixtures with CO and CO₂ at high temperatures, however. When CO has desorbed at 45 K, the $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ features are similar to pure HCOOH. Furthermore, its integrated intensity does not decrease after CO desorption. In the mixture of HCOOH with CO₂, the 1735 cm^{-1} and 1717 cm^{-1} bands of HCOOH only fully disappear between 75 K and 105 K at the same time that the features for pure HCOOH start to appear and after CO₂ has desorbed. Thus even when HCOOH is present as a minor component in CO or CO₂ it does not co-desorb at low temperatures. The HCOOH bands in CH₃OH have a similar lack of temperature variation as for mixtures with H₂O. Between 105 and 135 K CH₃OH has a phase transition that is observed around 1500 cm^{-1} where both the CH₃ deformation and $\nu_{\text{B}}(\text{OH})$ modes of CH₃OH are located. The HCOOH $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ bands do not resemble pure HCOOH features at 135 K, which indicates that both species are still mixed.

4.4. Tertiary mixtures

Spectra of tertiary mixtures ~7:67:26% of HCOOH:H₂O:CO, HCOOH:H₂O:CO₂ and HCOOH:H₂O:CH₃OH have been measured at 15 K. In Tables 6 and A.4 of the online material peak positions and *FWHMs* are given. The $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ features are similar to those in mixtures with H₂O, but are shifted by a few wavenumbers. The $\nu_{\text{B}}(\text{OH}/\text{CH})$ band is located at lower wavenumbers, especially for the HCOOH:H₂O:CH₃OH mixture and have *FWHMs* of ~14–19 cm^{-1} .

4.5. Comparison to previous laboratory data

Prior to this study, only Schutte et al. (1999) and Hudson & Moore (1999) reported infrared spectra of HCOOH mixtures with H₂O and CH₃OH at 1 and 4 cm^{-1} resolution, respectively. Their experiments only focused on the $\nu_{\text{B}}(\text{OH}/\text{CH})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ bands and consequently it is difficult to make a full comparison with our data which also includes other HCOOH bands. The peak positions of the $\nu_{\text{B}}(\text{OH}/\text{CH})$ features in pure HCOOH ice and 10:100% HCOOH:H₂O mixtures by Schutte et al. (1999) are shifted and the *FWHMs* are narrower compared to our data. However, the tertiary HCOOH:H₂O:CH₃OH 8:66:26% mixture results agree very well within the experimental uncertainties. These discrepancies are likely due to a different baseline subtraction. In pure HCOOH and mixtures with H₂O the $\nu_{\text{B}}(\text{OH}/\text{CH})$ mode has a wing on the blue side, which is not subtracted in the present work. The peak position and *FWHM* for the $\nu_{\text{S}}(\text{C}-\text{O})$ feature for a 7:93% HCOOH:H₂O mixture by Hudson & Moore (1999) are the same within the uncertainties.

5. Band strength changes in ice mixtures

As stated before in Sect. 3 absolute band strengths are difficult to measure in our experiment, whereas the relative band strengths can be accurately determined. The relative values discussed in this section can be used in combination with existing values in the literature given in Table 3 to derive absolute band strengths in different ice environments.

5.1. HCOOH

When HCOOH ice is mixed with H₂O the band strengths of HCOOH are strongly affected. In Fig. 5 the relative band strengths of the $\nu_{\text{S}}(\text{C}=\text{O})$, $\nu_{\text{S}}(\text{C}-\text{O})$ and $\nu_{\text{B}}(\text{OH}/\text{CH})$ spectral features are shown with error-bars of 12% superimposed. Clearly the uncertainties are relatively large, but still it is possible to note a number of effects with concentration and temperature. As discussed in Sect. 3 the error-bars in the same experiment at different temperatures are only 2%, and observed temperature trends are thus much more accurate than concentration trends.

All mixtures with H₂O show a similar temperature trend as pure HCOOH. Even though between 135 and 150 K the $\nu_{\text{S}}(\text{C}=\text{O})$ band splits, its band strength is constant with temperature. In contrast, that of the $\nu_{\text{S}}(\text{C}-\text{O})$ feature shows a pronounced dip around 135 K, which is partly recovered at 150 K. For the most diluted mixture where chain-formation does not occur, no such dip is present. The band strength of the $\nu_{\text{B}}(\text{OH}/\text{CH})$ mode increases with temperature and peaks at 135 K. A striking difference between the three absorption bands is that the band strength increases by a factor 2 to 5 with decreasing HCOOH ice concentration for the $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{B}}(\text{OH}/\text{CH})$ bands, but is the same within 30% for the $\nu_{\text{S}}(\text{C}-\text{O})$ feature in all mixtures.

A similar comparison for the relative band strength with respect to temperature is made for mixtures with CO, CO₂, and CH₃OH (see Fig. 6). The band strengths in mixtures with CO are higher compared to those in pure HCOOH, even when all CO has evaporated. However, the band strengths in CO₂ are close to or smaller than those for pure HCOOH. The $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ modes are largest around 75–90 K, decrease between 90–135 K and finally increase again up to 150 K. The $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ bands are relatively little affected in the mixtures with CH₃OH.

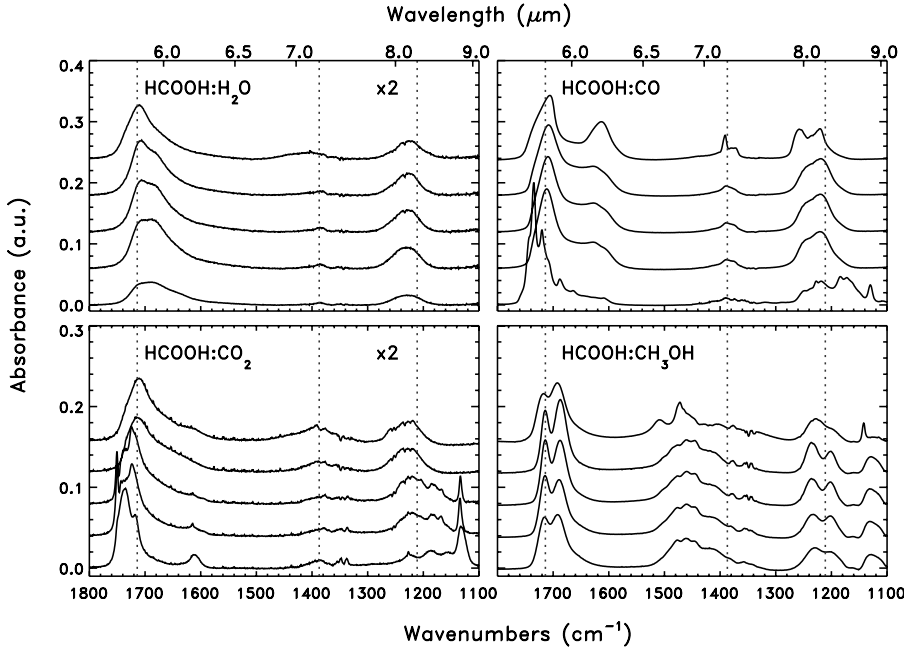


Fig. 4. Absorption spectra in the 1800–1100 cm^{-1} regime for mixtures with $\sim 90\%$ H_2O , CO , CO_2 and CH_3OH at 15 (*bottom curve*), 45, 75, 105, and 135 K (*top curve*). The $\nu_{\text{B}}(\text{OH})$ has been removed for the $\text{HCOOH}:\text{H}_2\text{O}$ mixture. The peak positions of pure HCOOH at 15 K are indicated by the dotted lines. The intensities of the bands are multiplied by a factor 2 for the $\text{HCOOH}:\text{H}_2\text{O}$ and $\text{HCOOH}:\text{CO}_2$ mixtures as indicated in the first and third panel, respectively.

Table 6. The peak positions, *FWHMs*, and relative band strengths of $\nu_{\text{B}}(\text{OH}/\text{CH})$ at 15 K of tertiary mixtures of $\text{HCOOH}:\text{H}_2\text{O}:\text{X}$, where X stands for CO , CO_2 , and CH_3OH . The uncertainties on the peak positions and *FWHMs* are $\pm 1 \text{ cm}^{-1}$ and $\pm 2 \text{ cm}^{-1}$, respectively.

Mixture	Peak position (cm^{-1})	<i>FWHM</i> (cm^{-1})	A/A_0
$\text{HCOOH}:\text{H}_2\text{O}:\text{CO}$	1386	19	4.2
$\text{HCOOH}:\text{H}_2\text{O}:\text{CO}_2$	1386	19	3.0
$\text{HCOOH}:\text{H}_2\text{O}:\text{CH}_3\text{OH}$	1382	13	8.3

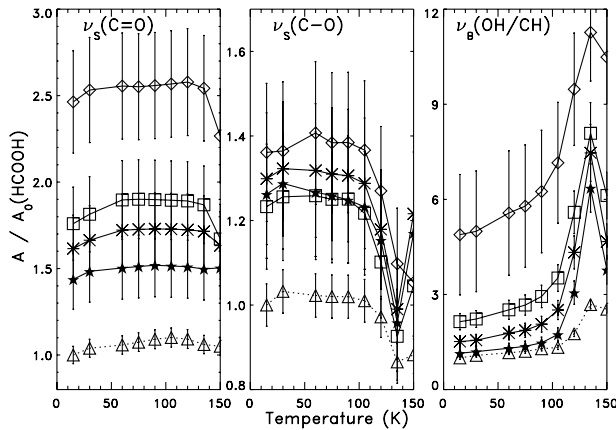


Fig. 5. $A/A_0(\text{HCOOH})$ with respect to temperature for the HCOOH $\nu_{\text{S}}(\text{C}=\text{O})$, $\nu_{\text{B}}(\text{OH}/\text{CH})$, and $\nu_{\text{S}}(\text{C}-\text{O})$ bands in $\text{HCOOH}:\text{H}_2\text{O}$. The \star symbols refer to the 50:50% mixture, $*$ to 34:66%, \square to 20:80%, \diamond to 9:91% $\text{HCOOH}:\text{H}_2\text{O}$, and \triangle to pure HCOOH . Note the different scales on the vertical axes.

The relative band strengths for the $\nu_{\text{B}}(\text{OH}/\text{CH})$ mode in the tertiary mixtures are shown in Table 6 and those of the other bands in Table A.4 of the online material. Clearly, the band strength in the tertiary mixtures with H_2O and CO , CO_2 or CH_3OH are enhanced at 15 K.

In summary, the relative band strengths of the $\nu_{\text{S}}(\text{C}=\text{O})$ and $\nu_{\text{S}}(\text{C}-\text{O})$ bands can range from 0.8–3 $A_0(\text{HCOOH})$ and

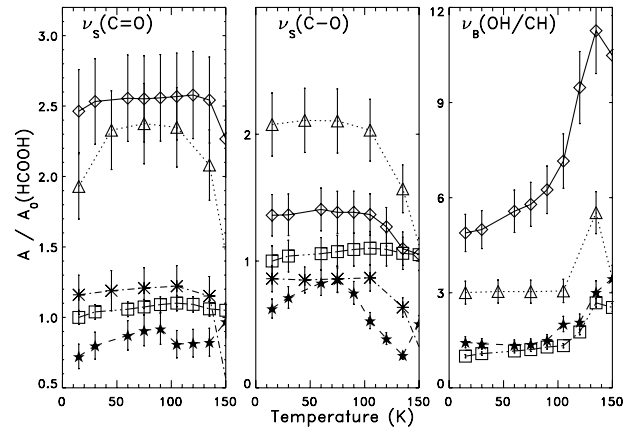


Fig. 6. $A/A_0(\text{HCOOH})$ with respect to temperature for the HCOOH $\nu_{\text{S}}(\text{C}=\text{O})$, $\nu_{\text{B}}(\text{OH}/\text{CH})$, and $\nu_{\text{S}}(\text{C}-\text{O})$ bands. The \square symbols refers to pure HCOOH , \diamond to 9:91% $\text{HCOOH}:\text{H}_2\text{O}$, \triangle to 11:89% $\text{HCOOH}:\text{CO}$, \star to 9:91% $\text{HCOOH}:\text{CO}_2$, and $*$ to 10:90% $\text{HCOOH}:\text{CH}_3\text{OH}$. Note the different scales on the vertical axes.

of the $\nu_{\text{B}}(\text{OH}/\text{CH})$ band from 1–12 $A_0(\text{HCOOH})$ in different ice environments.

5.2. H_2O

Water absorption bands are strongly affected by the presence of other ice components like CO_2 (Öberg et al. 2007). In Fig. 7 the relative band strengths compared to pure H_2O of the $\nu_{\text{S}}(\text{OH})$ and the $\nu(\text{lib})$ modes in the mixtures with HCOOH are shown versus temperature. The $\nu_{\text{S}}(\text{OH})$ band strength is significantly smaller in all mixtures compared to that of pure H_2O , but does not depend on the mixing ratio. The $\nu(\text{lib})$ mode behaves differently. It decreases as a function of HCOOH concentration, with the smallest band strength for the highest HCOOH concentrations. As the temperature increases the relative band strength of the $\nu(\text{lib})$ band increases, until a temperature of 135 K is reached. The different behavior of $\nu_{\text{S}}(\text{OH})$ and the $\nu(\text{lib})$ mode are further illustrated in the right panel of Fig. 7. For pure H_2O the

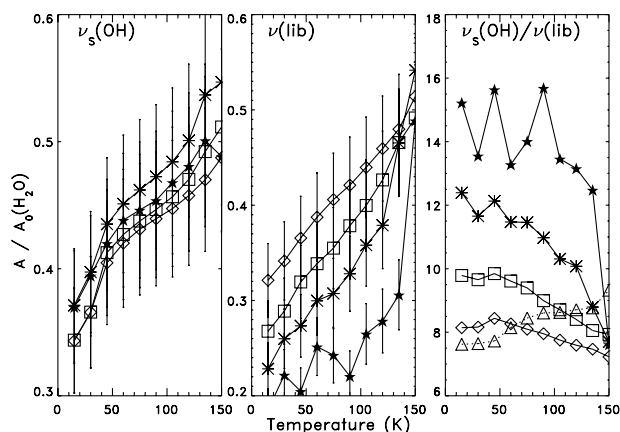


Fig. 7. $A/A_0(\text{H}_2\text{O})$ of the H_2O $\nu_S(\text{OH})$, $\nu(\text{lib})$, and the band strength ratio $A_{\nu_S(\text{OH})}/A_{\text{lib}}$ with respect to temperature in HCOOH: H_2O mixtures and pure H_2O . The dotted line with Δ in the $A_{\nu_S(\text{OH})}/A_{\text{lib}}$ panel indicates the trend for pure H_2O , the \star for the 50:50% mixture, $*$ for 34:66%, \square for 20:80%, and \diamond for 9:91% HCOOH: H_2O . Note the different scales on the vertical axes.

absorbance ratio $A_{\nu_S(\text{OH})}/A_{\nu(\text{lib})}$ (the 3.0/13 μm features) starts at ~ 7.5 and increases with temperature to 9.5. The mixtures have a higher ratio at 15 K, but decrease with temperature. The $A_{\nu_S(\text{OH})}/A_{\nu(\text{lib})}$ ratio furthermore increases with HCOOH concentration at 15 K. In conclusion, the ratio of the relative band strengths of the 3.0/13 μm features with a value larger than 10 implies that H_2O is present in ice with HCOOH concentrations $>20\%$, and a lower value implies that the HCOOH concentration is $<20\%$.

6. Discussion

In this section, our results are discussed in the light of recent chemical physics experiments of HCOOH-containing ices. The behavior of HCOOH in different ices may lead to important clues regarding the behavior of the species under interstellar conditions. Of particular relevance for interstellar chemistry is the mobility in the ice and desorption behavior of the ice constituents with temperature.

The infrared spectrum of pure HCOOH distinguishes itself from HCOOH mixtures by the presence of dimers to a chain-like structure of HCOOH molecules at 120–135 K for HCOOH deposited on a KBr substrate at 15 K (see e.g., Cyriac & Pradeep 2005). The chain-structure at higher temperatures results in a spectrum in which the $\nu_S(\text{C}=\text{O})$ and $\nu_S(\text{C}-\text{O})$ bands are split. The transition from dimers to HCOOH-chains is irreversible as evidenced by the splitting of the $\nu_S(\text{C}=\text{O})$ mode at 15 K for HCOOH deposited at 145 K (see Fig. 2 and Sect. 4.1).

In binary mixtures with H_2O , the HCOOH ice spectral features remain relatively similar throughout a large temperature regime. Only from 120 K onwards, changes are observed for our experiments (see Fig. 3). Initially, only the $\nu_B(\text{OH}/\text{CH})$ mode increases in intensity and broadens, but between 135 K and 150 K the $\nu_S(\text{C}=\text{O})$ band splits into the out-of-phase and in-phase modes for $>10\%$ HCOOH, similar to what is observed for pure HCOOH ice. This implies that the molecules first reorganize locally at 120 K and subsequently become mobile throughout the ice at higher temperatures. This is corroborated by studies of layered binary HCOOH- H_2O ices (Cyriac & Pradeep 2005; Borodin et al. 2005; Bahr et al. 2005; Souda 2006). In addition to the infrared experiments by Cyriac & Pradeep (2005), Metastable Impact Electron Spectroscopy (MIES) and

Temperature Programmed Desorption (TPD) have been used to probe the layered HCOOH/ H_2O ices. No mixing is observed in any of these experiments up to 120 K. Furthermore, the splitting of the $\nu_S(\text{C}=\text{O})$ mode between 135 K and 150 K in our experiments implies that H_2O and HCOOH segregate at high temperatures. This is consistent with TPD data from Bahr et al. (2005), where HCOOH desorbs after H_2O , even in ices where H_2O has been deposited on top of HCOOH.

The interactions between HCOOH with CO , CO_2 and CH_3OH are different compared to H_2O . Very little has been published about either dimers or ice mixtures of these species. Park et al. (2002) performed experiments with HCOOH dissolved in liquid CO_2 and C_2H_6 . The dimer/monomer ratio for HCOOH was higher in C_2H_6 than CO_2 . This is due to the stronger electrostatic interactions of HCOOH with CO_2 compared to C_2H_6 . A similar effect may be responsible for the difference between CO and CO_2 as the monomeric HCOOH feature at 1154 cm^{-1} is stronger in the mixture with CO_2 compared to CO . When HCOOH is diluted in an ice dominated by other molecules with no or very small dipole-moments like CH_4 , N_2 and O_2 , a similar HCOOH ice spectrum is likely obtained. The monomer/dimer ratio will depend on the HCOOH concentration as well as the strength of the electrostatic interaction between HCOOH and the other ice species. A similar behavior is observed for H_2O ice diluted in CO_2 , CO , N_2 and O_2 (e.g., Öberg et al. 2007; Bouwman et al., in prep.; Awad et al., in prep.).

The CH_3OH spectrum is very little affected by HCOOH. This is probably due to a comparable strength of hydrogen bonds between HCOOH and CH_3OH . The infrared spectra of HCOOH: CH_3OH 10:90% mixtures shown in Fig. 4 clearly illustrate that both species do not segregate at temperatures below 135 K, even when CH_3OH has a phase transition. It is likely that HCOOH and CH_3OH only become mobile between 135 K and 150 K, as is found for mixtures of HCOOH with H_2O . Since CH_3OH has a smaller dipole-moment and fewer possibilities for making hydrogen-bonds compared to H_2O , it is also expected to segregate from, and desorb prior to, HCOOH. Other polar or hydrogen-bonding species such as NH_3 and H_2CO may interact with HCOOH and affect its infrared spectrum in the same way.

7. Astrophysical implications

7.1. HCOOH ice identification

Formic acid ice has been observed in many different astrophysical objects, such as high mass and low mass star-forming regions (e.g., Schutte et al. 1997, 1999; Boogert et al., in prep.). In this section we compare the laboratory spectra with the observed ISO spectrum toward the low mass YSO HH 46 and the high mass YSO W 33A (Boogert et al. 2004; Gibb et al. 2000). The 5.9 μm ($\nu_S(\text{C}=\text{O})$) 7.2 μm ($\nu_B(\text{OH}/\text{CH})$) and 8.1 μm ($\nu_S(\text{C}-\text{O})$) bands are used to determine the ice environment in which HCOOH is located. The 5.9 μm HCOOH feature is blended with other interstellar bands and the 8.1 μm band overlaps with the wing of a strong silicate feature. Nonetheless they can be used to exclude some astrophysical HCOOH ice environments. In Fig. 8 the 5–9 μm range is shown toward the low mass YSO HH 46 with the experimental spectra overplotted. In Fig. 8a the spectra for the $\sim 10:90\%$ HCOOH: CO and HCOOH: CO_2 mixtures are shown. These mixtures can clearly be ruled out for HH 46 because the $\nu_S(\text{C}=\text{O})$ feature is shifted too far to the blue. Similarly, pure high temperature HCOOH (Fig. 8b) can be discarded as the out-of-phase $\nu_S(\text{C}-\text{O})$ mode at 8.1 μm band is not detected. Furthermore, for all pure

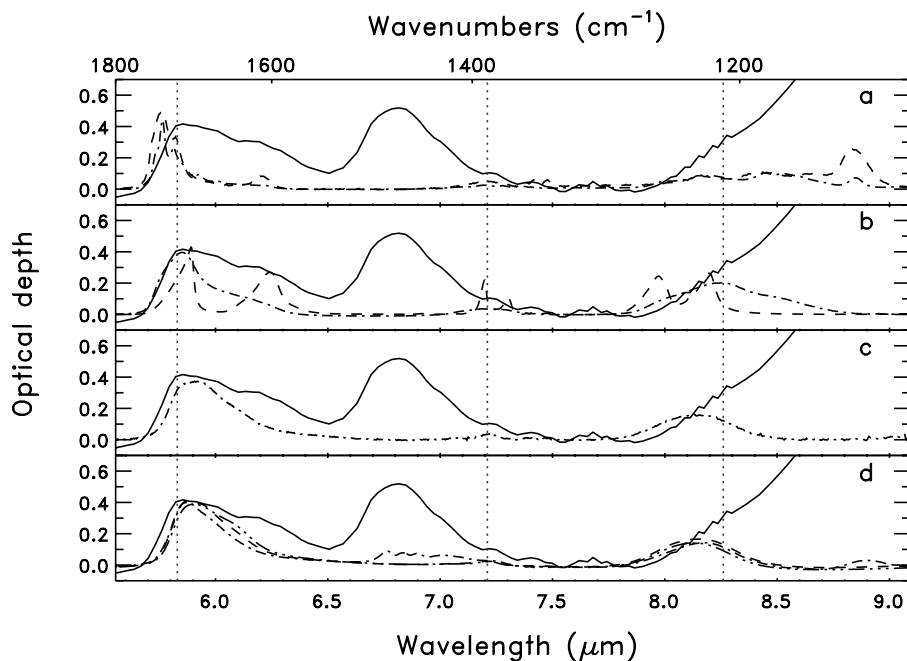


Fig. 8. The observed ice spectrum toward HH 46 in the 5–9 μm range (Boogert et al. 2004). Also included are in **a**) the $\sim 10:90\%$ mixtures of HCOOH:CO (dot-dash) and HCOOH:CO₂ (dashed), in **b**) pure HCOOH spectra at 15 K for ice deposited at 15 K (dot-dash) and at 145 K (dashed), in **c**) the $\sim 10:90\%$ mixtures of HCOOH:H₂O (dash-dot) and CH₃OH (dash-dot-dot-dot) and in **d**) the tertiary mixtures 7:67:26% of HCOOH:H₂O:CO (dash-dot-dot-dot), HCOOH:H₂O:CO₂ (dashed), and HCOOH:H₂O:CH₃OH (dot-dash) at 15 K. The relevant peak positions of pure HCOOH ice at 15 K are indicated with dotted lines to guide the eye.

ices and mixtures except the binary HCOOH:CH₃OH ice, the $\nu_{\text{S}}(\text{C}-\text{O})$ band causes a small wing on the silicate band at 8.1 μm , which is not observed due to the much larger intensity of the silicate absorption.

The integrated absorbance and peak position is determined most accurately from the $\nu_{\text{B}}(\text{OH}/\text{CH})$ band, because it is the least blended feature of HCOOH. In Fig. 9 the spectrum observed with ISO toward W 33A is shown for the $\nu_{\text{B}}(\text{OH}/\text{CH})$ feature. The ISO data are chosen for comparison because they have higher spectral resolution ($\lambda/\Delta\lambda \approx 800$, or 2 cm^{-1} at 7.2 μm) than the Spitzer data ($\lambda/\Delta\lambda \approx 100$, or 14 cm^{-1} at 7.2 μm). In Fig. 9a the 15 K spectra of pure HCOOH and 20:80% HCOOH:H₂O are shown. Clearly, the pure HCOOH ice spectrum is too broad to fit the astronomical feature. The $\nu_{\text{B}}(\text{OH}/\text{CH})$ mode is also too far shifted to the blue for the HCOOH:H₂O mixture. The tertiary mixtures depicted in Fig. 9b match the observations better. The HCOOH spectral features in the HCOOH:H₂O:CO and HCOOH:H₂O:CO₂ mixtures are almost identical, but have a blue wing that is not observed for W 33A. However, the HCOOH:H₂O:CH₃OH 6:68:26% mixture reproduces the observed spectrum very well.

This sensitivity of peak position and *FWHM* of the $\nu_{\text{B}}(\text{OH}/\text{CH})$ band at 7.2 μm is further illustrated in Fig. 10, together with the observed values toward the high mass YSO W 33A. A Gaussian fit is made to the observations in the same way as to the experimental data. The black horizontal line corresponds to the fitted peak position in the observations and the black vertical line to the fitted *FWHM*. The dotted and dashed lines represent 2 cm^{-1} and 1 cm^{-1} error-bars for the *FWHM* and peak position, respectively. Clearly, the binary mixtures with H₂O as well as pure HCOOH are too broad and shifted too far to the blue side compared to the observed feature in W 33A. The tertiary mixtures provide better fits. The 6:27:67% HCOOH:CO₂:H₂O and 8:62:30% HCOOH:H₂O:CO mixtures have *FWHM*s that are too broad compared to the observations within their uncertainties. Furthermore, their peak positions are shifted to 1386 cm^{-1} , a few wavenumbers too high. Potentially, CO₂/CO:H₂O ratios higher than 0.40–0.52:1 could shift the $\nu_{\text{B}}(\text{OH}/\text{CH})$ feature to lower wavenumbers

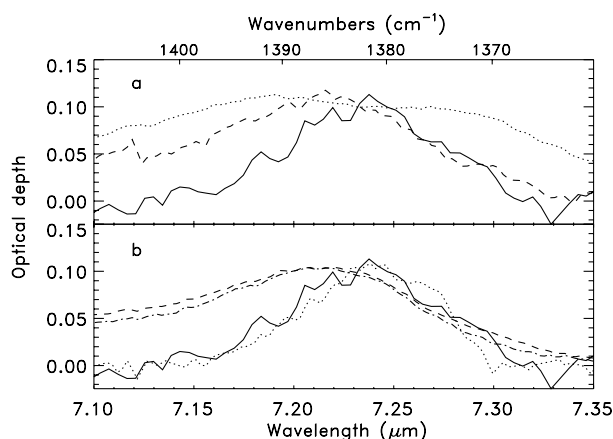


Fig. 9. The observed ice spectrum in the 7.10–7.35 μm range toward W 33A indicated by the solid line (Gibb et al. 2000). Also included are in **a**) the spectra at 15 K of pure HCOOH (dotted) and 20:80% HCOOH:H₂O (dashed) and in **b**) the tertiary mixtures HCOOH:H₂O:CO 8:62:30% (dashed), HCOOH:H₂O:CO₂ 6:67:27% (dash-dot) and HCOOH:H₂O:CH₃OH 6:68:26% (dotted).

and match the observations toward W 33A better. This is not unlikely as Öberg et al. (2007) find that they can fit the observed Spitzer spectrum toward HH 46 well if also a 0.5:1 CO₂:H₂O mixture is included. The 7.2 μm band for the HCOOH:CH₃OH:H₂O mixture fits the observed spectrum very well.

7.2. HCOOH ice formation

The presence of HCOOH in a complex mixture dominated by H₂O but also including CO, CO₂ or CH₃OH or a combination of these three species is not unlikely when potential formation mechanism are evaluated. CO or CO₂ are both potential precursors to HCOOH and CH₃OH is known to form from CO through H-atom bombardment (Watanabe et al. 2004; Fuchs et al., in prep.). H-bombardment of CO or CO₂ in mixtures with H₂O (Fuchs et al., in prep.; Bisschop et al., in prep.) and

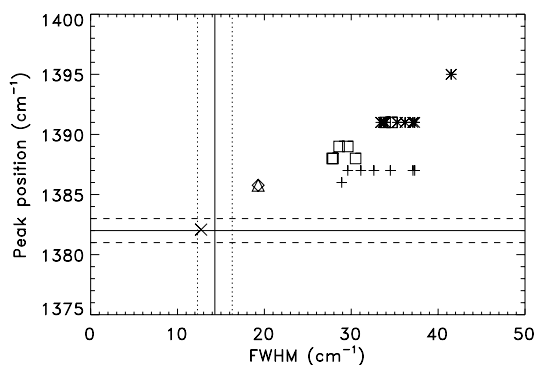
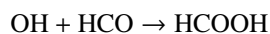
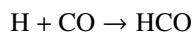
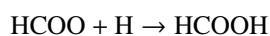
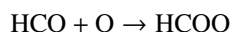


Fig. 10. Comparison between the observed 7.2 μm band of HCOOH in W 33A with laboratory spectra of HCOOH ice. The black lines indicate the peak position and $FWHM$ of the observed feature and the dotted and dashed lines 2 cm^{-1} uncertainties on the $FWHM$ and 1 cm^{-1} on the peak position, respectively. Mixture ratios for all temperatures have the same symbol. The + symbols refer to pure HCOOH, * to 50:50% HCOOH:H₂O, \square to 9:91% HCOOH:H₂O, Δ to HCOOH:H₂O:CO 8:62:30%, \diamond to 6:67:27% HCOOH:H₂O:CO₂ and \times to 6.5:26:68% HCOOH:CH₃OH:H₂O.

UV-irradiation of CO:H₂O and CO₂:H₂O mixtures (Watanabe et al. 2004; Watanabe & Kouchi 2002; Gerakines et al. 2000) do not give HCOOH as one of the main products. Since the UV experiments were performed under HV conditions, they need to be repeated under ultra high vacuum conditions to fully exclude this route. In H₂O:CO mixtures that are bombarded by 0.8 MeV protons, to simulate cosmic rays, HCOOH is formed (Hudson & Moore 1999). Its formation is explained by:



where the H and OH result from H₂O dissociation in the experiment. In space, H-atoms from the gas phase can react with the adsorbed O to form OH. HCO is formed by H addition to CO, which will further react to H₂CO and finally to CH₃OH, unless HCO encounters an OH radical first. Garrod & Herbst (2006) use this formation mechanism for HCOOH in the solid state. However, in their model HCOOH can only form at higher temperatures because the OH and HCO radicals are not mobile at low temperatures. This would be inconsistent with the detection of HCOOH ice in cold sources such as HH 46 and background stars unless OH is produced with excess energy. This is for example found in theoretical calculations by Andersson et al. (2006) for the photo-dissociation of H₂O ice. The newly formed OH radicals are not very mobile in bulk ice, but can move more than 80 Å over the surface. Thus HCOOH formation through this reaction mechanism is expected to be strongly localized at the ice surface. Another potential formation mechanism is given by Keane (2001).



In this case successive H- and O-addition leads to HCOOH formation, and HCOOH may be produced at low ice temperatures at the same time as H₂O and CH₃OH are formed.

In summary, both reaction pathways make it likely that HCOOH is formed in an environment where H₂O and CH₃OH are located. Potentially CO, CO₂, and H₂CO are present in the

ice as well, if conversion to CH₃OH and HCOOH is incomplete. As discussed in Sect. 7.1 it is probable that H₂CO will have a similar effect on the ice structure as CH₃OH and H₂O and thus also on the HCOOH infrared bands. In the future both reaction paths can be tested experimentally and thus resolve whether HCOOH is formed during ice formation, ice processing at low temperatures or potentially even both.

8. Summary and conclusions

The main conclusions derived from this work are:

- The peak positions, $FWHM$ and profiles of the $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH}/\text{CH})$ and $\nu_S(\text{C}-\text{O})$ bands are strongly affected by the ice matrix. At high temperatures and when HCOOH is deposited at high temperatures, the $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH}/\text{CH})$ and $\nu_S(\text{C}-\text{O})$ modes split into two bands. In binary mixtures with H₂O, the HCOOH features have very simple profiles, whereas multiple bands are found in binary ice mixtures with CO, CO₂ and CH₃OH. The peak positions of the HCOOH $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH}/\text{CH})$, and $\nu_S(\text{C}-\text{O})$ bands are shifted up to $\sim 40 \text{ cm}^{-1}$. The spectra for tertiary mixtures resemble the spectra of binary mixtures with H₂O, but small shifts exist.
- The band strengths of the $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH}/\text{CH})$, and $\nu_S(\text{C}-\text{O})$ features are strongly affected by the ice structure. They are enhanced in mixtures of HCOOH with H₂O and CO by factors of 2–10, but are similar to those of pure HCOOH ice for binary mixtures with CO₂ and CH₃OH. Strong temperature effects are also observed.
- Comparison of our data with the Spitzer spectrum of HH 46 and the ISO spectrum toward the high mass YSO W 33A excludes the presence of the high temperature HCOOH chain-structure, as well as binary mixtures of HCOOH with any of the other species H₂O, CO, CO₂ and CH₃OH. The 7.2 μm feature toward W 33A is very similar to that found in tertiary mixtures of $\sim 7:67:26\%$ HCOOH:H₂O:X, where X stands for CO, CO₂ or CH₃OH. Especially the HCOOH:H₂O:CH₃OH mixture reproduces the observed W 33A spectrum very well. Potentially, such tertiary or even more complex ices consisting of HCOOH, H₂O, CH₃OH, CO and CO₂ may be responsible for the observed spectrum. The presence of HCOOH in such complex ice mixtures is consistent with its possible formation mechanisms.

Acknowledgements. We thank the referee for useful comments on the paper. Funding was provided by NOVA, the Netherlands Research School for Astronomy, and by a Spinoza grant from the Netherlands Organization for Scientific Research, NWO.

References

- Andersson, S., Al-Halabi, A., Kroes, G.-J., & van Dishoeck, E. F. 2006, *J. Chem. Phys.*, 124, 4715
- Bahr, S., Borodin, A., Höfft, O., Kempter, V., & Allouche, A. 2005, *J. Chem. Phys.*, 122, 234704
- Bisschop, S. E., Jørgensen, J. K., van Dishoeck, E. F., & de Wachter, E. B. M. 2007, *A&A*, 465, 913
- Boogert, A. C. A., Pontoppidan, K. M., Lahuis, F., et al. 2004, *ApJS*, 154, 359
- Borodin, A., Höfft, O., Bahr, S., Kempter, V., & Allouche, A. 2005, *Nucl. Instrum. Meth. B*, 232, 79
- Bottinelli, S., Ceccarelli, C., Williams, J. P., & Lefloch, B. 2007, *A&A*, 463, 601
- Cyriac, J., & Pradeep, T. 2005, *Chem. Phys. Lett.*, 402, 116
- Dartois, E., Schutte, W., Geballe, T. R., et al. 1999, *A&A*, 342, L32

- Ehrenfreund, P., Dartois, E., Demyk, K., & D'Hendecourt, L. 1998, *A&A*, 339, L17
- Ehrenfreund, P., Kerkhof, O., Schutte, W. A., et al. 1999, *A&A*, 350, 240
- Garrod, R. T., & Herbst, E. 2006, *A&A*, 457, 927
- Gerakines, P. A., Moore, M. H., & Hudson, R. L. 2000, *A&A*, 357, 793
- Gerakines, P. A., Schutte, W. A., Greenberg, J. M., & van Dishoeck, E. F. 1995, *A&A*, 296, 810
- Gibb, E. L., Whittet, D. C. B., Boogert, A. C. A., & Tielens, A. G. G. M. 2004, *ApJS*, 151, 35
- Gibb, E. L., Whittet, D. C. B., Schutte, W. A., et al. 2000, *ApJ*, 536, 347
- Hollis, J. M., Pedelty, J. A., Snyder, L. E., et al. 2003, *ApJ*, 588, 353
- Hudson, R. L., & Moore, M. H. 1999, *Icarus*, 140, 451
- Ikeda, M., Ohishi, M., Nummelin, A., et al. 2001, *ApJ*, 560, 792
- Keane, J. V. 2001, Ph.D. Thesis, Rijks Universiteit Groningen
- Knez, C., Boogert, A. C. A., Pontoppidan, K. M., et al. 2005, *ApJ*, 635, L145
- Liu, S.-Y., Girart, J. M., Remijan, A., & Snyder, L. E. 2002, *ApJ*, 576, 255
- Marechal, Y. 1987, *J. Chem. Phys.*, 87, 6344
- Millikan, R. C., & Pitzer, K. S. 1958, *J. Am. Chem. Soc.*, 80, 3515
- Öberg, K. I., Fraser, H. J., Boogert, A. C. A., et al. 2007, *A&A*, 462, 1187
- Park, J.-Y., & Woon, D. E. 2006, *ApJ*, 648, 1285
- Park, Y., Gupta, R. B., Curtis, C. W., & Roberts, C. B. 2002, *J. Phys. Chem.*, 106, 9696
- Pontoppidan, K. M., Fraser, H. J., Dartois, E., et al. 2003, *A&A*, 408, 981
- Remijan, A. J., & Hollis, J. M. 2006, *ApJ*, 640, 842
- Requena-Torres, M. A., Martín-Pintado, J., Rodríguez-Franco, A., et al. 2006, *A&A*, 455, 971
- Schutte, W. A., Boogert, A. C. A., Tielens, A. G. G. M., et al. 1999, *A&A*, 343, 966
- Schutte, W. A., Greenberg, J. M., van Dishoeck, E. F., et al. 1997, *Ap&SS*, 255, 61
- Souda, R. 2006, *Surface Science*, 600, 3135
- Tielens, A. G. G. M., & Charnley, S. B. 1997, *Origins Life Evol. B.*, 27, 23
- Tielens, A. G. G. M., Tokunaga, A. T., Geballe, T. R., & Baas, F. 1991, *ApJ*, 381, 181
- van Dishoeck, E. F., Blake, G. A., Jansen, D. J., & Groesbeck, T. D. 1995, *ApJ*, 447, 760
- Watanabe, N., & Kouchi, A. 2002, *ApJ*, 567, 651
- Watanabe, N., Nagaoka, A., Shiraki, T., & Kouchi, A. 2004, *ApJ*, 616, 638
- Yokoyama, I., Miwa, Y., & Machida, K. 1991, *J. Am. Chem. Soc.*, 113, 6458

Online Material

Appendix A: Line parameters of HCOOH in the 1800–1000 cm⁻¹ range**Table A.1.** Peak positions, *FWHMs* and A/A_0 for the HCOOH $\nu_S(\text{C=O})$, $\nu_B(\text{OH/CH})$, and $\nu_S(\text{C-O})$ bands for all temperatures for pure HCOOH. The peak positions are given in both wavenumbers and wavelengths for the ice deposited at 15 K and 145 K. The uncertainties on the peak positions and *FWHMs* are ± 1 and ± 2 cm⁻¹, respectively.

<i>T</i> (K)	$\nu_S(\text{C=O})$					$\nu_B(\text{OH/CH})$			$\nu_S(\text{C-O})$		
	out-of-phase		in-phase		A/A_0^a	ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	A/A_{0a}	ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	A/A_0^a
ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	ν_S (cm ⁻¹ (μm))							
HCOOH deposited at 15 K											
15	1714(5.84)	45	1650(6.06)	69	1.0	1387(7.21)	37	1.0	1211(8.26)	64	1.0
30	1713(5.84)	43	1651(6.06)	71	1.0	1387(7.21)	37	1.1	1213(8.24)	61	1.0
60	1710(5.85)	41	1644(6.08)	65	1.1	1387(7.21)	35	1.2	1219(8.20)	53	1.0
75	1710(5.85)	40	1641(6.09)	65	1.1	1387(7.21)	33	1.2	1221(8.19)	51	1.0
90	1709(5.85)	39	1637(6.11)	64	1.1	1387(7.21)	31	1.3	1223(8.18)	50	1.0
105	1709(5.85)	39	1634(6.12)	63	1.1	1387(7.21)	30	1.3	1224(8.17)	50	1.0
120	1709(5.85)	41	1625(6.15)	56	1.1	1386(7.21)	29	1.8	1228(8.15)	51	1.0
135	1703(5.87)	10	1615(6.19)	43	1.1	1391(7.19)	8	2.7	1255(7.97)	22	0.9
	1716(5.83)	39				1375(7.27)	14		1221(8.19)	21	
150	1703(5.87)	11	1615(6.19)	40	1.1	1391(7.19)	7	2.5	1254(7.97)	22	0.9
	1719(5.82)	34				1374(7.28)	13		1221(8.19)	17	
165	1704(5.87)	11	1617(6.18)	36	—	1391(7.19)	7	—	1253(7.98)	20	—
	1722(5.81)	32				1373(7.28)	12		1220(8.20)	15	
HCOOH deposited at 145 K											
15	1699(5.89)	10	1606(6.23)	36	1.5	1390(7.19)	6	3.7	1256(7.96)	18	1.3
	1712(5.84)	32				1371(7.29)	9		1222(8.18)	15	
30	1699(5.89)	9	1606(6.23)	36	1.5	1390(7.19)	6	3.7	1256(7.96)	18	1.3
	1712(5.84)	32				1371(7.29)	9		1222(8.18)	15	
45	1699(5.89)	9	1606(6.22)	36	1.5	1390(7.19)	6	3.7	1255(7.97)	18	1.3
	1713(5.84)	32				1371(7.29)	9		1222(8.18)	15	
60	1699(5.89)	8	1607(6.22)	36	1.5	1390(7.19)	6	3.6	1255(7.97)	18	1.3
	1712(5.84)	32				1371(7.29)	9		1222(8.19)	15	
75	1699(5.88)	8	1608(6.22)	36	1.5	1390(7.19)	6	3.6	1255(7.97)	18	1.3
	1713(5.84)	32				1371(7.29)	9		1222(8.19)	15	
90	1700(5.88)	8	1608(6.22)	37	1.5	1390(7.19)	6	3.5	1254(7.97)	18	1.3
	1713(5.84)	32				1371(7.29)	10		1221(8.19)	15	
105	1700(5.88)	9	1609(6.21)	37	1.5	1390(7.19)	6	3.5	1254(7.98)	18	1.3
	1713(5.84)	32				1371(7.29)	10		1221(8.19)	15	
120	1700(5.88)	9	1610(6.21)	38	1.4	1390(7.19)	6	3.4	1253(7.98)	18	1.3
	1713(5.84)	32				1371(7.29)	10		1220(8.19)	15	
135	1700(5.88)	9	1611(6.21)	39	1.4	1390(7.20)	6	3.3	1252(7.99)	18	1.3
	1714(5.84)	32				1371(7.29)	11		1220(8.10)	15	
150	1701(5.88)	9	1612(6.20)	39	1.4	1390(7.20)	6	3.2	1252(7.99)	19	1.3
	1714(5.83)	31				1371(7.29)	11		1219(8.20)	14	
165	1701(5.88)	9	1614(6.20)	38	1.3	1390(7.20)	6	2.8	1250(8.00)	19	1.2
	1715(5.83)	31				1371(7.30)	11		1219(8.20)	14	

^a A_0 refers to the band strength in pure HCOOH at 15 K.

Table A.2. Peak positions, *FWHMs* and A/A_0 for the HCOOH $\nu_S(\text{C}=\text{O})$, $\nu_B(\text{OH}/\text{CH})$, and $\nu_S(\text{C}-\text{O})$ bands for all temperatures for mixtures of HCOOH with H₂O. The uncertainties on the peak positions and *FWHMs* are ± 1 and ± 2 cm⁻¹, respectively.

T (K)	$\nu_S(\text{C}=\text{O})$					$\nu_B(\text{OH}/\text{CH})$			$\nu_S(\text{C}-\text{O})$		
	out-of-phase		in-phase			ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	A/A_0^a	ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	A/A_0^a
ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	ν_S (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	A/A_0^a							
50:50% HCOOH:H ₂ O											
15	1708(5.85)	48	1658(6.03)	103	1.4	1391(7.19)	37	1.1	1215(8.23)	50	1.3
30	1709(5.85)	39	1675(5.97)	75	1.5	1391(7.19)	37	1.2	1220(8.20)	58	1.3
45	1711(5.85)	36	1665(6.01)	78	1.5	1391(7.19)	36	1.3	1223(8.18)	55	1.3
60	1710(5.85)	33	1672(5.98)	83	1.5	1391(7.19)	35	1.3	1223(8.17)	54	1.3
75	1709(5.85)	35	1668(6.00)	77	1.5	1391(7.19)	34	1.4	1224(8.17)	52	1.3
90	1709(5.85)	34	1670(6.00)	78	1.5	1391(7.19)	33	1.5	1225(8.17)	51	1.2
105	1709(5.85)	35	1669(5.99)	75	1.5	1392(7.19)	34	1.7	1225(8.16)	51	1.2
120	1710(5.85)	37	1671(5.98)	75	1.5	1395(7.17)	42	3.1	1224(8.17)	50	1.2
135	1713(5.84)	45	1670(5.99)	69	1.5	1403(7.13)	59	6.4	1223(8.17)	51	0.95
150	1707(5.86)	10	1616(6.19)	40	1.5	1391(7.19)	7	3.8	1255(7.97)	19	1.2
	1720(5.82)	16				1372(7.29)	13		1222(8.18)	17	
165	1706(5.86)	12	1619(6.18)	34	—	1391(7.19)	6	—	1253(7.98)	18	—
	1723(5.80)	32				1373(7.29)	12		1222(8.19)	14	
34:66% HCOOH:H ₂ O											
15	1707(5.86)	44	1652(6.05)	95	1.6	1391(7.19)	35	1.5	1216(8.23)	46	1.3
30	1706(5.86)	38	1688(5.92)	94	1.7	1391(7.19)	35	1.6	1224(8.17)	56	1.3
45	1707(5.86)	36	1667(6.00)	83	1.7	1391(7.19)	33	1.7	1226(8.16)	51	1.3
60	1706(5.86)	37	1668(6.00)	77	1.7	1391(7.19)	32	1.8	1226(8.15)	50	1.3
75	1706(5.86)	36	1669(5.99)	75	1.7	1391(7.19)	32	1.9	1227(8.15)	49	1.3
90	1706(5.86)	36	1670(5.99)	74	1.7	1391(7.19)	32	2.1	1227(8.15)	49	1.3
105	1706(5.86)	36	1671(5.98)	73	1.7	1393(7.18)	34	2.5	1227(8.15)	48	1.3
120	1708(5.85)	39	1674(5.98)	73	1.7	1399(7.15)	46	4.3	1226(8.16)	48	1.2
135	1714(5.84)	47	1670(5.99)	69	1.7	1406(7.11)	62	7.5	1224(8.17)	51	1.0
150	1711(5.84)	4	1617(6.18)	42	1.6	1391(7.19)	7	4.7	1256(7.96)	18	1.2
	1719(5.82)	34				1371(7.29)	13		1223(8.17)	16	
165	—	—	1621(6.17)	32	—	1391(7.19)	5	—	1255(7.97)	16	—
	1726(5.79)	28				1372(7.29)	11		1224(8.17)	13	
20:80% HCOOH:H ₂ O											
15	1707(5.86)	40	1643(6.09)	81	1.8	1391(7.19)	35	2.1	1215(8.23)	44	1.2
30	1704(5.87)	43	1667(6.00)	89	1.8	1390(7.19)	33	2.2	1227(8.15)	55	1.3
45	1703(5.87)	44	1660(6.02)	77	1.9	1390(7.19)	31	2.3	1230(8.13)	47	1.3
60	1703(5.87)	40	1669(5.99)	83	1.9	1390(7.19)	31	2.5	1229(8.13)	48	1.3
75	1703(5.87)	42	1664(6.01)	76	1.9	1391(7.19)	31	2.7	1229(8.13)	47	1.3
90	1703(5.87)	39	1670(5.99)	79	1.9	1391(7.19)	32	2.9	1229(8.14)	46	1.3
105	1705(5.87)	40	1668(5.99)	75	1.9	1393(7.18)	35	3.5	1228(8.14)	46	1.2
120	1709(5.85)	39	1675(5.97)	76	1.9	1403(7.13)	51	5.6	1227(8.15)	47	1.1
135	1714(5.83)	45	1673(5.98)	72	1.9	1413(7.08)	61	8.2	1225(8.16)	49	0.93
150	—	—	1619(6.17)	51	1.7	1392(7.19)	8	6.1	1256(7.96)	19	1.0
	1717(5.82)	41				1372(7.29)	17		1223(8.17)	19	
165	—	—	-1622(6.16)	31	—	1392(7.18)	4	—	1255(7.97)	15	—
	1730(5.78)	27				1375(7.27)	11		1226(8.16)	11	
9.4:91% HCOOH:H ₂ O											
15	1708(5.86)	38	1636(6.11)	70	2.5	1391(7.19)	35	4.9	1214(8.24)	42	1.4
30	1706(5.86)	40	1680(5.95)	128	2.5	1388(7.21)	31	4.9	1233(8.11)	59	1.4
45	1698(5.89)	45	1662(6.02)	94	2.6	1389(7.20)	29	5.2	1231(8.12)	42	1.4
60	1699(5.89)	55	1649(6.06)	90	2.6	1388(7.20)	28	5.6	1231(8.12)	46	1.4
75	1698(5.89)	49	1655(6.04)	100	2.6	1388(7.20)	28	5.8	1231(8.12)	45	1.4
90	1700(5.88)	52	1641(6.10)	88	2.6	1389(7.20)	30	6.2	1230(8.13)	44	1.4
105	1701(5.88)	50	1645(6.08)	89	2.6	1391(7.19)	34	7.2	1230(8.13)	44	1.4
120	1706(5.86)	47	1664(6.01)	94	2.6	1402(7.13)	52	9.5	1228(8.14)	45	1.3
135	1713(5.84)	46	1658(6.03)	76	2.5	1415(7.06)	59	11.3	1226(8.15)	47	1.1
150	1718(5.82)	43	1674(5.97)	93	2.3	1404(7.12)	67	10.5	1223(8.18)	51	1.1

^a A_0 refers to the band strength in pure HCOOH at 15 K.

Table A.3. Peak positions, *FWHMs* and A/A_0 for the HCOOH $\nu_S(\text{C}=\text{O})$ and $\nu_S(\text{C}-\text{O})$ bands for all temperatures for $\sim 10:90\%$ HCOOH:CH₃OH. The uncertainties on the peak positions and *FWHMs* are ± 1 and ± 2 cm⁻¹, respectively.

T (K)	$\nu_S(\text{C}=\text{O})$					$\nu_S(\text{C}-\text{O})$				
	out-of-phase		in-phase			out-of-phase		in-phase		
	ν (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	ν (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)		ν (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	ν (cm ⁻¹ (μm))	<i>FWHM</i> (cm ⁻¹)	
15	1721(5.81)	15	1691(5.91)	31	1.2	1231(8.12)	29	1198(8.35)	26	0.86
45	1717(5.82)	18	1690(5.92)	34	1.2	1235(8.10)	24	1203(8.31)	21	0.85
75	1716(5.83)	11	1689(5.92)	33	1.2	1234(8.10)	24	1201(8.33)	18	0.85
105	1716(5.83)	13	1686(5.93)	28	1.2	1234(8.10)	21	1202(8.32)	18	0.86
135	1725(5.80)	13	1696(5.90)	44	1.2	1230(8.13)	27	1201(8.33)	22	0.63

^a A_0 refers to the band strength in pure HCOOH at 15 K.**Table A.4.** Peak positions and band strengths for the $\nu_S(\text{C}=\text{O})$ and $\nu_S(\text{C}-\text{O})$ bands for the binary $\sim 10:90\%$ HCOOH:CO and HCOOH:CO₂ mixtures and tertiary $\sim 7:67:26\%$ HCOOH:H₂O:CO, HCOOH:H₂O:CO₂ and HCOOH:H₂O:CH₃OH mixtures at 15 K. The uncertainties on the peak positions are ± 1 cm⁻¹.

Mode	ν (cm ⁻¹)	A/A_0^a
11:89% HCOOH:CO		
$\nu_S(\text{C}=\text{O})$	1743, 1735, 1721, 1702	1.9
$\nu_B(\text{OH}/\text{CH})$	1390	3.0
$\nu_S(\text{C}-\text{O})$	1250, 1235, 1227, 1218, 1185, 1172, 1130	2.1
9:91% HCOOH:C ¹⁸ O ₂		
$\nu_S(\text{C}=\text{O})$	1744, 1734, 1717, 1612	0.72
$\nu_B(\text{OH}/\text{CH})$	1387	1.4
$\nu_S(\text{C}-\text{O})$	1227, 1187, 1157, 1132	0.63
8:62:30% HCOOH:H ₂ O:CO		
$\nu_S(\text{C}=\text{O})$	1705, 1685	4.7
$\nu_S(\text{C}-\text{O})$	1227	2.4
6:67:27% HCOOH:H ₂ O:C ¹⁸ O ₂		
$\nu_S(\text{C}=\text{O})$	1695	4.2
$\nu_S(\text{C}-\text{O})$	1229	1.5
6:68:26% HCOOH:H ₂ O:CH ₃ OH		
$\nu_S(\text{C}=\text{O})$	1703	3.7
$\nu_S(\text{C}-\text{O})$	1223	1.5

^a A_0 refers to the band strength in pure HCOOH at 15 K.