

Photochemistry of the PAH pyrene in water ice: the case for ion-mediated solid-state astrochemistry

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ABSTRACT

Context. Icy dust grains play an important role in the formation of complex molecules in the interstellar medium (ISM). Laboratory studies have mainly focused on the physical interactions and chemical pathways in ices containing rather simple molecules, such as H₂O, CO, CO₂, CH₄, and CH₃OH. Observational studies show that polycyclic aromatic hydrocarbons (PAHs) are also abundantly present in the ISM in the gas phase. It is likely that these non-volatile species also freeze-out onto dust grains and participate in the astrochemical solid-state network, but additional experimental PAH ice studies are largely lacking.

Aims. The study presented here focuses on a rather small PAH, pyrene (C₁₆H₁₀), and aims to understand and quantify photochemical reactions of PAHs in interstellar ices upon vacuum ultraviolet (VUV) irradiation as a function of astronomically relevant parameters.

Methods. Near UV/VIS spectroscopy is used to track the in situ VUV driven photochemistry of pyrene containing ices at temperatures ranging from 10 to 125 K.

Results. The main photoproducts of VUV photolyzed pyrene ices are spectroscopically identified and their band positions are listed for two host ices, H₂O and CO. Pyrene ionization is found to be most efficient in H₂O ices at low temperatures. The reaction products, triplet pyrene and the 1-hydro-1-pyrenyl radical are most efficiently formed in higher temperature water ices and in low temperature CO ice. Formation routes and band strength information of the identified species are discussed. Additionally, the oscillator strengths of Py, Py⁺, and PyH⁺ are derived and a quantitative kinetic analysis is performed by fitting a chemical reaction network to the experimental data.

Conclusions. Pyrene is efficiently ionized in water ice at temperatures below 50 K. Hydrogenation reactions dominate the chemistry in low temperature CO ice with trace amounts of water. The results are placed in an astrophysical context by determining the importance of PAH ionization in a molecular cloud. We conclude that the rate of pyrene ionization in water ice mantles is comparable to the rate of photodesorption of H₂O ice. The photoprocessing of a sample PAH in ice described in this manuscript indicates that PAH photoprocessing in the solid state should also be taken into account in astrochemical models.

Key words. astrochemistry – molecular processes – methods: laboratory – techniques: spectroscopic

1. Introduction

Strong infrared emission attributed to polycyclic aromatic hydrocarbons (PAHs) is characteristic of many galactic and extragalactic objects (Smith et al. 2007; Draine et al. 2007; Draine & Li 2007; Tielens 2008). While this emission generally originates in optically thin, diffuse regions, PAHs should also be common throughout the dense interstellar medium. There, as with most other interstellar species in molecular clouds, PAHs condense out of the gas onto cold icy grain mantles, where they are expected to influence or participate in the chemistry and physics of the ice. While laboratory studies of interstellar ice analogs have shown that complex organic molecules are produced upon extended vacuum ultraviolet (VUV) photolysis (e.g., Briggs et al. 1992; Bernstein et al. 1995), the photoinduced processes occurring during the irradiation of PAH containing interstellar ice analogues have not yet been studied in detail. In optical, in situ studies of the photochemistry of naphthalene, 4-methylpyrene, and quaterylene containing water ice at 20 K, Gudipati & Allamandola (2003, 2006a,b) and Gudipati (2004) showed that these PAHs are readily ionized and stabilized within the ice, suggesting that trapped ions may play important, but

overlooked roles in cosmic ice processes. Beyond this, there is little information about the VUV induced, in situ photochemistry and photophysics of PAH-containing water-rich ices.

Here, we describe a detailed study of the VUV-induced photochemistry that takes place within pyrene (Py or C₁₆H₁₀) containing water ices (Py:H₂O = 1:10 000–1:5000). The present study is an extension of the study of Bouwman et al. (2009) in which the focus was on the new experimental setup and where the use of PAH ice spectra was discussed to search for solid-state features of PAHs in space. In this work, the focus is on a detailed characterization of the chemical processes taking place upon VUV irradiation, particularly as a function of ice temperature ranging from 25 to 125 K. Additionally, measurements on Py:CO ices at 10 K were performed to elucidate the role of water in the reaction schemes and to clarify the formation routes of identified species. A similar study of three small PAHs is now underway to understand the general principles of PAH/ice photochemistry. This is part of an overall experimental program at the Sackler Laboratory for Astrophysics to study the fundamental processes of inter- and circumstellar ice analogues such as thermal (Acharyya et al. 2007) and photodesorption (Öberg et al. 2007b, 2009c), hydrogenation reactions

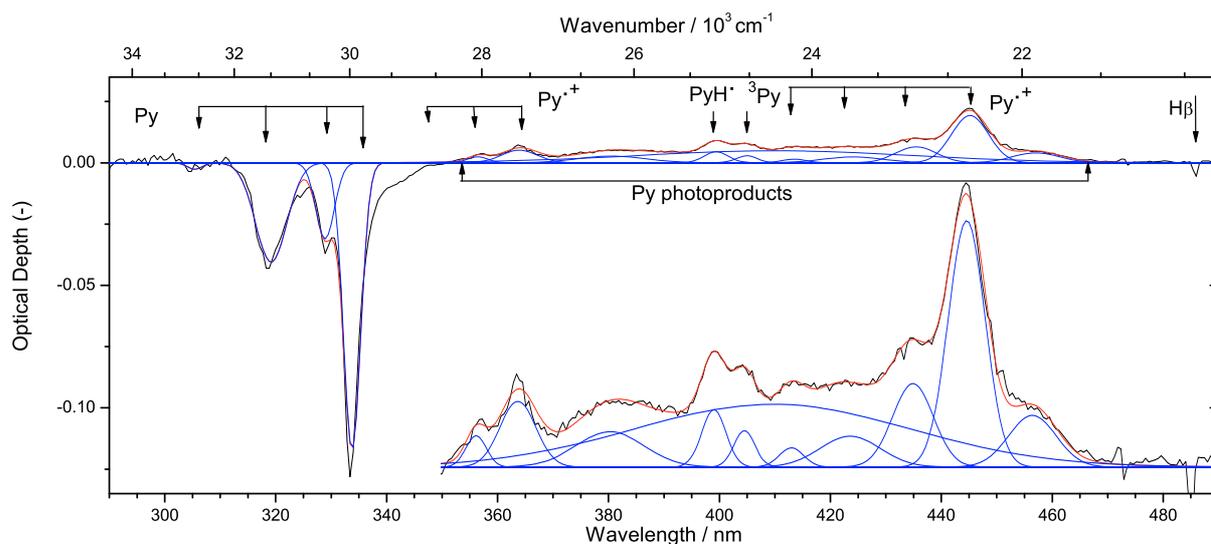


Fig. 1. The spectrum of a dilute pyrene:H₂O ice after 900 s of VUV irradiation at 125 K. The inset shows a blow-up of the pyrene photoproduct bands. Band assignments are discussed in Sect. 3. Note the broad feature ranging from about 350 to 470 nm which is indicated by a Gaussian fit. This is attributed to overlapping bands from individual pyrene photoproducts. Bands with negative optical depth indicate species destruction, those with positive optical depth show species formation. The blue bands are Gaussian profiles which co-add to the overall fit shown in red. Note the instrumental resolution indicated by the profile of the H β line at 486.1 nm.

(Fuchs et al. 2009; Ioppolo et al. 2008), photochemistry (Öberg et al. 2009b), and physical interactions in interstellar ice analogues (Bouwman et al. 2007; Öberg et al. 2007a, 2009a).

The manuscript is organized as follows. The experimental technique is summarized in Sect. 2. Section 3 describes the Py:H₂O and Py:CO ice photochemistry, the resulting products and their formation routes. The temperature-dependent photochemistry and derived reaction dynamics are described in Sect. 4 and astrochemical implications are discussed in Sect. 5. The main conclusions are summarized in Sect. 6.

2. Experimental technique

We use a new apparatus as described in Bouwman et al. (2009), which follows the photochemistry in kinetic mode during VUV irradiation by measuring the near UV-visible absorption spectra of an ice, providing “real-time” tracking of the reactants and photoproducts. Dilute Py:H₂O ice samples (\sim 1:10000– \sim 1:5000) and a Py:CO ice sample of comparable concentration are prepared by depositing the vapor from a pyrene sample heated to 40°C together with H₂O vapor or CO gas onto a cold MgF₂ window. The window is cooled to 10 K in the case of CO deposition or 25 K in the case of H₂O deposition. The sample window is cooled by a closed cycle He refrigerator. Pyrene (Aldrich, 99%) and CO (Praxair 99.999%) are used as commercially available. Vapor from water, filtered through a milli-Q purification system and purified further by three freeze-pump-thaw cycles, is used. The sample window is mounted in a high-vacuum chamber ($P \approx 10^{-7}$ mbar). The ice growth rate and thickness are determined with a HeNe laser by monitoring the thin-film interference fringes generated during deposition. Simultaneously, the amount of pyrene is tracked by measuring the integrated strength of the $S_2 \leftarrow S_0$ neutral Py transition at 334 nm. Deposition is typically stopped when the optical depth (OD) of Py approaches \sim 0.15.

The ice samples are photolyzed with the 121.6 nm Ly α (10.6 eV) and the 160 nm molecular hydrogen emission bands (centered around 7.8 eV) generated by a microwave powered

discharge in a flowing H₂ gas with a vacuum ultraviolet flux of $\sim 10^{15}$ photons s⁻¹ (Muñoz Caro et al. 2002). This results in a photon flux of $\sim 10^{14}$ photons cm⁻² s⁻¹ at the sample surface (Öberg et al. 2009c).

Absorption spectra of VUV-photolyzed Py-containing ices are measured with a Xe-arc lamp serving as a white light source. Lenses and diaphragms direct the light through the ice sample along the optical axis determined by the HeNe laser beam after which it is focused onto the entrance slit of a 0.3 m spectrometer. A 150 lines mm⁻¹ grating, blazed at 300 nm, disperses the light onto a sensitive 1024 \times 256 pixel CCD camera with 16 bit digitization. The camera is read out in vertical binning mode by a data acquisition computer that converts the data to absorbance spectra ($OD = -\ln(I/I_0)$). This configuration spans the 270 to 830 nm spectral range, which permits simultaneously monitoring of the behavior of the neutral Py parent molecule and photoproduct bands without any adjustment of the elements along the optical path. This is critical to obtaining reliable and reproducible baselines in measuring the optical spectra of ices. The spectral resolution is of the order of 0.9 nm, which is more than sufficient to record broad solid-state absorption features.

The measurements described here were performed on various H₂O:Py ice samples at 25, 50, 75, 100, and 125 K. The CO ice experiments were carried out at 10 K to avoid matrix sublimation at higher temperatures. The sample temperature is maintained using a resistive heater with an accuracy of ± 2 K. The measured spectra are converted into units of optical depth by using the spectrum of the freshly deposited, unphotolyzed ice at the appropriate temperature as a reference spectrum (I_0). Recording a single spectrum typically takes about 5 ms, and 229 spectra are generally coadded to improve the S/N of a spectrum, producing one single spectrum every 10 seconds.

The optical configuration of the apparatus is such that spectra are recorded *simultaneously* with photolysis. Thus, the short spectral recording time permits us to monitor photoinduced changes on a roughly 10-second time scale. Figure 1 shows the 290 to 490 nm spectrum of a Py:H₂O ice at 125 K after 900 s of in situ VUV photolysis. Because the spectrum recorded before

VUV irradiation is taken as a reference (I_0), bands with positive OD values originate from species that are produced by photolysis, while the bands with negative OD correspond to the neutral pyrene that is lost upon photolysis. Comparing the Py and photoproduct absorption bands with the narrow H β lamp line at 486.1 nm shows that the instrumental resolution indeed far exceeds the ice band widths. The absolute wavelength calibration is accurate to within ± 0.5 nm.

More than 1400 individual spectra are recorded and are reduced in a typical 4 h experiment. Spectra are individually baseline-corrected by fitting a second order polynomial through data points where no absorption occurs and subsequently subtracting the fit from the measured spectrum. Integrated absorbances of absorption features are calculated numerically for all spectra. These are corrected for the contributions of atomic hydrogen lines originating in the H₂ discharge lamp. The data reduction software also allows us to plot correlation diagrams between integrated absorbances of different absorption features. All data handling and reduction is performed with LabView routines.

Integrated band areas are used, in conjunction with oscillator strengths (f), to derive molecular abundances. The oscillator strength is converted to integrated absorbance (cm molecule^{-1}) using the conversion factor 8.88×10^{-13} (Kjaergaard et al. 2000). The number of molecules per cm^2 (N) is given by

$$N = \frac{\int_{\nu_1}^{\nu_2} \tau d\nu}{8.88 \times 10^{-13} f}, \quad (1)$$

where τ is the optical depth and ν is the frequency in cm^{-1} .

3. Band assignments and band strength analysis

The typical photolysis duration of about 4 h is the time required for nearly complete loss of the neutral pyrene vibrational progression at 334.0, 329.2, and 319.2 nm. Irradiating the sample ices with VUV light produces a set of new absorption bands in the spectra, indicating active photochemistry. The band positions, $FWHM$, and assignments of the bands in the Py:H₂O ice at 25 K are listed in Table 1. The bands appearing in the Py:CO ice at 10 K are similar to those in the Py:H₂O ice at 25 K, although, with slightly altered band positions and $FWHM$ and with very different relative intensities (see also Table 1). Figure 1 presents a spectrum from the 125 K Py:H₂O series. This figure illustrates production of the pyrene radical cation (Py^+), triplet pyrene (^3Py), 1-hydro-1-pyrenyl radical (PyH^\cdot), and a broad underlying “residue” feature upon VUV irradiation. Additionally, a progression of distinct absorption features is found in the Py:CO experiment, which indicates the formation of the (reactive intermediate) HCO $^\cdot$ radical. The identifications of these species and their oscillator strengths are discussed below.

3.1. Neutral pyrene bands

As in Bouwman et al. (2009), the strong, negative bands peaking at 334.0 nm and weaker bands at 329.2 and 319.2 nm in the H₂O ice (see Fig. 1), and at slightly shifted positions in the CO ice, are assigned to the $^1B_{2u} \leftarrow ^1A_g$ electronic transition of neutral pyrene ($S_2 \leftarrow S_0$) based on previous studies of pyrene in rare gas matrices (Vala et al. 1994; Halasinski et al. 2005). To study the chemistry in absolute number densities, a value of $f = 0.33$ is adopted from the literature for the oscillator strength of pyrene (Bitto et al. 2000; Wang et al. 2003). This value is used throughout this paper both for the Py:H₂O and Py:CO experiments. Pure

Table 1. Band positions (λ_c) and $FWHM$ in nm for pure pyrene ice at 10 K, pyrene in H₂O ice at 25 K, pyrene in CO ice at 10 K, and photoproduct bands for the Py:H₂O and Py:CO UV processed ices.

Species	Pyrene		Pyrene:H ₂ O		Pyrene:CO	
	λ_c	$FWHM$	λ_c	$FWHM$	λ_c	$FWHM$
Py $^1B_{2u}$	312.7	7.1	319.2	6.5	319.4	7.5
	325.3	10.0	329.2	3.2	329.2	2.3
	341.5	14.0	334.0	4.4	334.3	4.1
Py $^+$ $^2B_{1u}$	363.2	3.6	... ^a	... ^a
	354.0	6.5	... ^a	... ^a
	344.9	6.2	... ^a	... ^a
Py $^+$ 2A_u	445.6	6.6	445.3	7.8
	435.5	10.2	... ^a	... ^a
	423.0	12.2	... ^a	... ^a
	413.8	5.3	... ^a	... ^a
Py $^+$ $^2B_{1u}$	490.1	10.0	... ^a	... ^a
PyH $^\cdot$	399.4 ^b	5.2 ^b	400.5	4.2
 ^a	... ^a	392.5	6.7
 ^a	... ^a	378.4	15.7
^3Py $^3A_g^-$	405.0 ^b	4.5 ^b	406.2	4.8
HCO $^\cdot$ $^2A''$	513.4	17.5
	535.3	12.5
	556.3	14.5
	583.0	16.8
	604.9	10.0
	639.2	15.1

Notes. (a) Absorption feature was too weak to perform an accurate fit; (b) features are too weak at 25 K; the 125 K values are indicated.

pyrene ice measured at 10 K shows broader absorptions located at 341.5, 325.3, and 312.7 nm (see Table 1). We did not perform VUV experiments on the pure pyrene sample.

3.2. Pyrene cation bands

Positive bands at 363.2, 354.0, and 344.9 nm appear upon photolysis in the Py:H₂O experiments. This progression is assigned to the $^2B_{1u} \leftarrow ^2B_{3g}$ vibronic transition of the pyrene cation (Py^+) in accordance with the proximity to the band positions reported by Vala et al. (1994) and Halasinski et al. (2005). This transition for Py^+ in H₂O ice was reported in Bouwman et al. (2009). The $^2B_{1u} \leftarrow ^2B_{3g}$ transition is too weak to be detected in the Py:CO experiment. A stronger Py^+ progression occurs at 445.6, 435.5, 423.0, and 413.8 nm in water ice. Of these bands, only the strongest at 445.3 nm is detectable in the irradiated Py:CO ice. This progression is assigned to the $^2A_u \leftarrow ^2B_{3g}$ transition of Py^+ . The much weaker absorption caused by the $^2B_{1u} \leftarrow ^2B_{3g}$ Py^+ transition at 490.1 nm in H₂O is again undetectable in CO.

In these H₂O and CO ice experiments, Py^+ formation is the result of direct single photon ionization of the neutral species, following:



We emphasize that ionization in Py:H₂O ices is far more efficient than in Py:CO ices. Additional measurements on Py:CO:H₂O mixtures indicate that the presence of H₂O indeed enhances the ionization. Hence, it is possible that water contamination in the CO ice is responsible for the formation of some, if not all, of the cation species in the Py:CO experiment. The role of water contamination in CO ice will be discussed in more detail in Sect. 3.3.

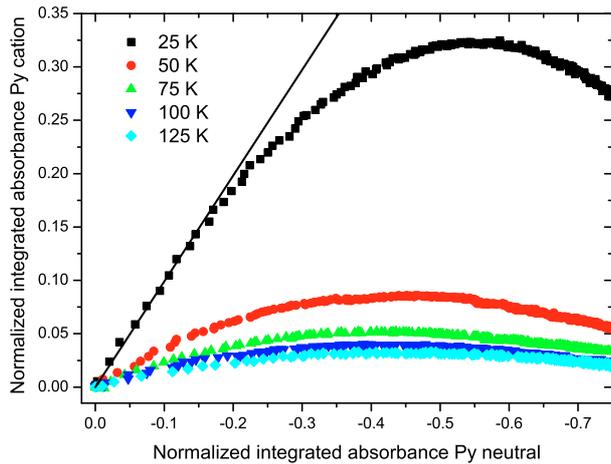


Fig. 2. Integrated absorbance of the 445.6 nm Py^+ band growth plotted against the loss of the 334.0 nm Py band in the 25, 50, 75, 100, and 125 K $\text{Py}:\text{H}_2\text{O}$ ices. Ten seconds elapse between subsequent data points. Values are normalized to the maximum integrated absorbance of neutral pyrene. The straight line portion of these plots is used to determine the oscillator strength of Py^+ as described in Sect. 3.2.

Using baseline-corrected spectra as shown in Fig. 1, the photochemical evolution is tracked by integrating areas of bands produced by each species in every spectrum and plotting them as a function of photolysis time. The strongest Py^+ band at ~ 445 nm is selected to track the number density evolution of this species. To put the kinetic analysis (Sect. 4) on a quantitative footing, we determine the oscillator strength of the 445 nm Py^+ band as follows. First, the integrated absorbance of the 445 nm Py^+ band is plotted versus that of the 334 nm Py band during the course of VUV photolysis at different ice temperatures. These graphs are shown in Fig. 2. It should be noted that there is a tight, linear behavior between the loss of neutral pyrene and growth of the pyrene cation during early photolysis times up to 100 s (the first 10 successive datapoints). Inspection of Fig. 2 shows that the slope is steepest and the ratio of the integrated absorbance of the Py^+ band to the Py band is optimum in the 25 K ice. Since no other photoproduct bands are evident during the linear correlation stage, we conclude that during this phase, neutral pyrene is converted solely into the cation as described previously for naphthalene and quaterrylene (Gudipati & Allamandola 2006a). The straight-line portion, fitted through the first 10 data points of irradiation at 25 K, is used to determine the oscillator strength of Py^+ . Given that the ratio of the Py^+ to the 334 nm Py band is 0.99 and the oscillator strength of this Py transition is 0.33, the oscillator strength of the 443 nm Py^+ band in water ice is also taken to be 0.33. This conclusion is consistent with ab initio calculations on pyrene by Weisman et al. (2005). They calculated that the oscillator strength of the cation is only $\sim 2\%$ stronger than that of the neutral species.

As described below, the photolysis of Py in water ices at higher temperatures produces other species in addition to the cation. This explains the different curves in Fig. 2.

3.3. HCO bands in $\text{Py}:\text{CO}$

VUV irradiation of a $\text{Py}:\text{CO}$ ice also produces a vibrational progression ranging from ~ 500 to 650 nm. As shown in Fig. 3, these absorption bands, located at 513.4, 535.3, 556.3, 583.0, 604.9, and 639.2 nm, are assigned to the ${}^2A''(0, \nu'', 0) \leftarrow X^2A'(0, 0, 0)$ HCO^{\cdot} ($\nu'' = 8\text{--}13$) transitions based on band positions reported

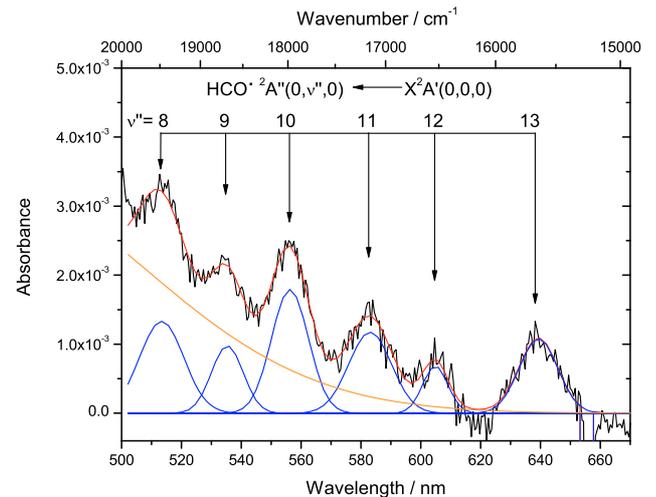


Fig. 3. Vibrational progression of HCO^{\cdot} generated in a $\text{Py}:\text{CO}$ ice at 10 K after 600 s of VUV irradiation plotted together with a Gaussian fit (red) to the absorption spectrum (black). The individual Gaussians are shown in blue. The orange line indicates the red wing of the underlying broad absorption feature (Sect. 3.6).

by van Ijzendoorn et al. (1983). The clear HCO^{\cdot} progression indicates a photolytic source of free H atoms in the ice. In addition, it confirms the ability of this setup to record small reactive intermediates in the ice.

A possible explanation of the source of H atoms is related to the experimental conditions. The experiments reported here are performed under high vacuum (10^{-7} mbar) conditions. Therefore, background H_2O vapor has ample time to condense onto the sample window while cooling down and growing the ice sample. Water is well known to photodissociate upon VUV irradiation (e.g., Öberg et al. 2009c; Andersson & van Dishoeck 2008) according to



An experiment on VUV irradiation of a “pure” CO ice indicated that HCO^{\cdot} is also efficiently produced in the absence of pyrene. Therefore, it is likely that water contamination is responsible for the production of HCO^{\cdot} via



Another possible formation route could be by means of VUV-induced hydrogen abstraction from pyrene. This pyrene photodissociation reaction, however, is unlikely to occur, since PAHs are generally highly photostable molecules.

3.4. The 400 nm band carrier

Another vibrational progression appears at 400.5, 392.5, and 378.4 nm in the CO ice experiments. As shown in Fig. 4, the 400.5 nm band dominates this progression. In contrast, a single band appears at 399.4 nm in the $\text{Py}:\text{H}_2\text{O}$ ice upon VUV irradiation of the samples. The relative intensity of these bands varies with respect to the Py^+ bands. The 400 nm bands are more pronounced than the cation bands in the H_2O ice only at high temperatures, whereas they are more pronounced in the low temperature CO ice.

Two additional measurements were performed to identify the carrier responsible for these transitions. A kinetic experiment was performed on non-VUV-irradiated $\text{Py}:\text{CO}$ ice. This

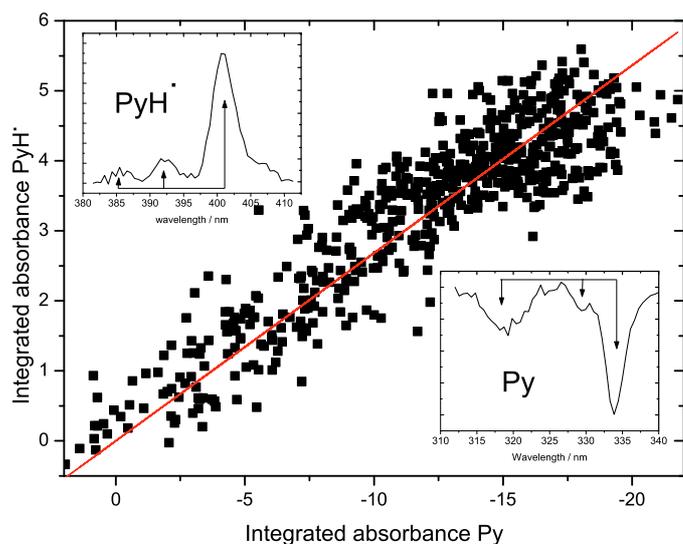


Fig. 4. Integrated absorbance of the 400 nm PyH' band growth plotted against the loss of the 334 nm Py band in a CO ice at 10 K *after* VUV irradiation is stopped. The straight line directly reflects the relative oscillator strength of both bands as described in Sect. 3.2. The two insets show the PyH' and Py vibrational progressions in a CO ice 90 min after photolysis is stopped.

ice showed no sign of pyrene ionization by the Xe-lamp, which is used as a spectroscopic light source. The production of HCO[•] and the formation of the 400 nm band were not observed either. Subsequently, the ice was irradiated by the VUV source for 10 min. The steady growth of the 400 nm band with VUV photolysis indicates that the species responsible for the 400 nm band is a product of the VUV processing of the ice. Moreover, when the VUV irradiation is stopped, the 400 nm band carrier *continues* to grow at the expense of the remaining neutral pyrene. This indicates that the chemical reaction leading to the formation of this species is not directly photon-dependent, but rather depends on the diffusion of a photoproduct. A similar experiment on a Py:H₂O ice at 25 K indicates that the same process also takes place in H₂O ice. The detection of HCO[•] radicals in the ice and the inherent presence of free photolytic H atoms, implies that the growth of the vibrational progression starting at ~400 nm could be the result of the reaction of pyrene with diffusing H atoms



This assignment to the 1-hydro-1-pyrenyl radical (PyH[•]) is supported by other experimental studies (Okada et al. 1976, 1980), where progressions at similar band positions are observed upon (laser) flash photolysis.

In contrast to the Py:H₂O experiments where pyrene is also efficiently ionized, the experiment on PyH[•] formation in CO shows no sign of other reaction products. The integrated absorbance of the growing PyH[•] transition is plotted versus the integrated absorbance of the diminishing neutral in Fig. 4. Growth is tracked over a duration of more than 1.5 h. Since there is a one-to-one conversion of Py to PyH[•] in the Py:CO ice (Eq. (5)), as described in Sect. 3.2, we derive an oscillator strength of 0.089 for this species by fitting a line through the correlating absorbances in Fig. 4.

3.5. The 405 nm band carrier

Besides the Py⁺ and the PyH bands, another distinct absorption is found in the spectra of VUV irradiated ices. This feature is located at 405.0 nm in the Py:H₂O and at 406.2 nm in the Py:CO experiment. In our previously published paper on low temperature Py:H₂O ice, we tentatively assigned this absorption to a negative ion, Py⁻ or PyO⁻ (Bouwman et al. 2009). The experiments on Py:CO ices presented here enable us to exclude this assignment because of the nearly absent Py⁺ transitions. Firstly, Py⁻ is ruled out because a much stronger second Py⁻ absorption band, expected at 490 nm (Montejano et al. 1995), is absent in our Py:CO experiment. Secondly, PyO⁻ is also ruled out, because it should exhibit absorption bands down to 350 nm (Milosavljevic & Thomas 2002), bands that are also absent in the Py:CO experiment. Additionally, in our previous paper we assumed that PyO⁻ was a product of PyOH. The formation of PyO⁻ is also unlikely in the absence of PyOH absorption in these experiments, as discussed below.

The absorption at 405 nm does not correlate with that of the cation, nor with the PyH[•] band. The band only appears during photolysis and hence is characterized as a VUV-photon-related product. From the literature, it is known that a pyrene triplet-triplet (³A_g⁻ ← ³B_{2u}⁺) transition is expected at this wavelength upon laser excitation of pyrene in solution which populates the lowest member of the triplet manifold (e.g., Hsiao & Webber 1992; Langelaar et al. 1970). For the 405 nm band to originate from this triplet-triplet transition, the lowest level must be populated and remain so with a long enough lifetime to allow absorption to the ³A_g⁻ level. In the ice experiments reported here, there are a number of possible routes for pumping the ³Py state. The most obvious route is by means of photoexcitation followed by intersystem crossing



Triplet formation is found to decrease with decreasing temperature in ethanol ice (Stevens et al. 1967). This translates to our experiment in a nearly absent 405 nm band in the low temperature Py:H₂O experiment, because of the high Py⁺ formation efficiency. In the high temperature H₂O ice experiments, on the other hand, the 405 nm absorption is much stronger because pyrene is available.

In the CO ice, on the other hand, where Py⁺ production is low, formation of the 405 nm band carrier appears to be very efficient at low temperatures. The production of the 405 nm band carrier requires VUV photons to be initiated. The pumping of the ³Py state can again occur by means of Eq. (6). Moreover, CO has a dipole allowed electronic transition in the VUV. Hence, speculating, pumping of the ³Py state by collisional de-excitation of CO molecules excited by the VUV radiation provides a reaction path of



followed by



In summary, while we cannot identify the carrier of the 405 nm band, the ³A_g⁻ ← ³B_{2u}⁺ transition seems a plausible explanation.

3.6. Broad absorption feature

Finally, besides the narrower bands reported in the previous sections, we discuss a broad underlying feature extending from

about 350 to 570 nm, which grows upon photolysis in all cases. This band probably comprises overlapping bands caused by a number of Py/H₂O or Py/CO photoproducts. Part of this Py-residue feature remains even after warming up the sample window to room temperature, whereas all other features disappear at the water desorption temperature.

As discussed above, the very broad feature must be produced by a variety of similar but distinct photoproducts, all containing the pyrene chromophore. Mass spectral analysis of the species produced by the VUV photolysis of a few other PAHs in water ice show that the parent PAH is not destroyed but that OH, O, and H are added to some of the edge carbon atoms (Bernstein et al. 1999). Given the multiplicity of the side sites on pyrene that can undergo substitution, it is likely that the photoproducts produced in the experiments reported here are multiply substituted, rather than singly substituted. Thus, it is possible that a mixture of related but distinct Py-X_n species, where X may be H, OH, or O, produce the broad band.

In our previous work, we reported the production of a clear and reproducible PyOH band at 344.9 nm in a low temperature H₂O ice (Bouwman et al. 2009). The results presented here do not show evidence of this absorption feature. However, in some instances the absorption was detected upon irradiation or warm-up of the ice. The irregular appearance of the PyOH absorption feature in these experiments indicates that the formation of this species is highly sensitive to the sample's physical parameters, i.e., structure of the ice, temperature, and concentration. One possible explanation is that in the previously reported experiments, the Py concentration was not controlled and those experiments sampled a very different ice concentration and, by implication, physical ice structure. While we do not have a solution for this discrepancy, we emphasize that both measurement series have been fully reproducible over many independent experiments for periods of months. An experimental program to investigate the role of the PAH:H₂O concentration on ice photochemistry is underway.

4. Py:H₂O ice photochemistry at different temperatures

Figure 5 shows the spectral evolution of two different Py:H₂O samples at different temperatures. The top frame presents the 280 to 540 nm spectra of the 100 K Py:H₂O ice after 0, 20, 40, 80, and 160 s of in situ photolysis and the bottom frame the corresponding spectra for the 25 K ice. These spectra are snapshots of the more than one thousand spectra collected during 4 h of photolysis. They illustrate the rapid changes that occur during the early stages in the photochemistry of these ices and the major differences in reaction products at different temperatures.

To probe the VUV-driven photophysics and reaction dynamics for a set of selected temperatures, the production and depletion of species was tracked as a function of irradiation time. To this end, the Py 334 nm, Py⁺ 445 nm, and PyH⁺ 400 nm bands were integrated for every spectrum. The spectra in Figs. 1 and 5 show that it is rather straightforward to determine the boundaries needed to integrate these bands. We estimate that the uncertainty in most of these band areas is of the order of 10%.

The integrated absorbances of the neutral Py, strongest Py⁺, and PyH⁺ bands in H₂O ice at temperatures of 25, 50, 75, 100, and 125 K are plotted versus photolysis time (VUV fluence) in Fig. 6. The spectra in Fig. 5 and photochemical behavior in Fig. 6 show that, upon photolysis, neutral pyrene loss is immediate and rapid. The initial growth of Py⁺ mirrors the rapid, initial loss of Py. However, while Py steadily decreases, and several other Py

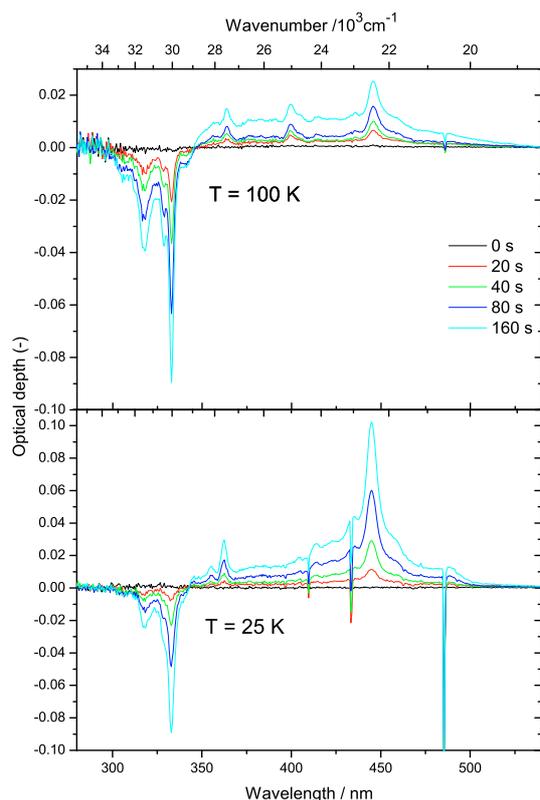


Fig. 5. The VUV-induced spectroscopic changes in Py:H₂O ice for two different temperatures as a function of photolysis time. Comparing the spectra from the 25 K ice (bottom) with those of the 100 K ice (top) shows the critical role that temperature plays in determining photochemical pathways in a PAH-containing ice. In the 25 K ice, cation formation is favored over production of the pyrene residue and the 400 and 405 nm band carriers. The opposite holds for the 100 K ice.

photoproduct bands increase during some 4 h of photolysis, the production of Py⁺ reaches a maximum and then slowly diminishes. From Fig. 6, one can clearly see that ionization of pyrene is most efficient in the low temperature ice. Formation of PyH⁺, on the other hand, is far more efficient at higher temperatures.

For comparison, the integrated absorbances for the irradiated Py:CO ice are plotted as a function of time in the right bottom frame of Fig. 6. It should be noted that the PyH⁺ band is multiplied by a factor of 10 in the Py:CO experiment, compared to a factor of 20 in the Py:H₂O experiment. The PyH⁺ band is clearly more prominent in the CO ice experiment than in the H₂O ice experiments. The Py⁺ signal on the other hand is negligible. This indicates that the H₂O ice plays a role in ion formation and stabilization.

To place this behavior on a quantitative footing, the integrated areas for the Py and Py⁺ bands are converted to number densities using Eq. (1). Here, an oscillator strength of 0.33 is used for the 334 nm Py bands. The values used for the oscillator strengths of the Py⁺ and PyH⁺ bands are 0.33 and 0.089, respectively, as determined in Sects. 3.2 and 3.4. Perusal of Fig. 6 shows that Py behaves similarly in all of the ices considered here. Regardless of temperature, its signal drops quickly with the onset of irradiation and continues to diminish with ongoing photolysis. Likewise, Py⁺ grows rapidly with initial photolysis but peaks after a relatively short time interval corresponding to a fluence of roughly 8×10^{16} photons and then drops continuously. While the Py⁺ growth and loss curves resemble each other, cation production efficiency is strongest in the 25 K ice.

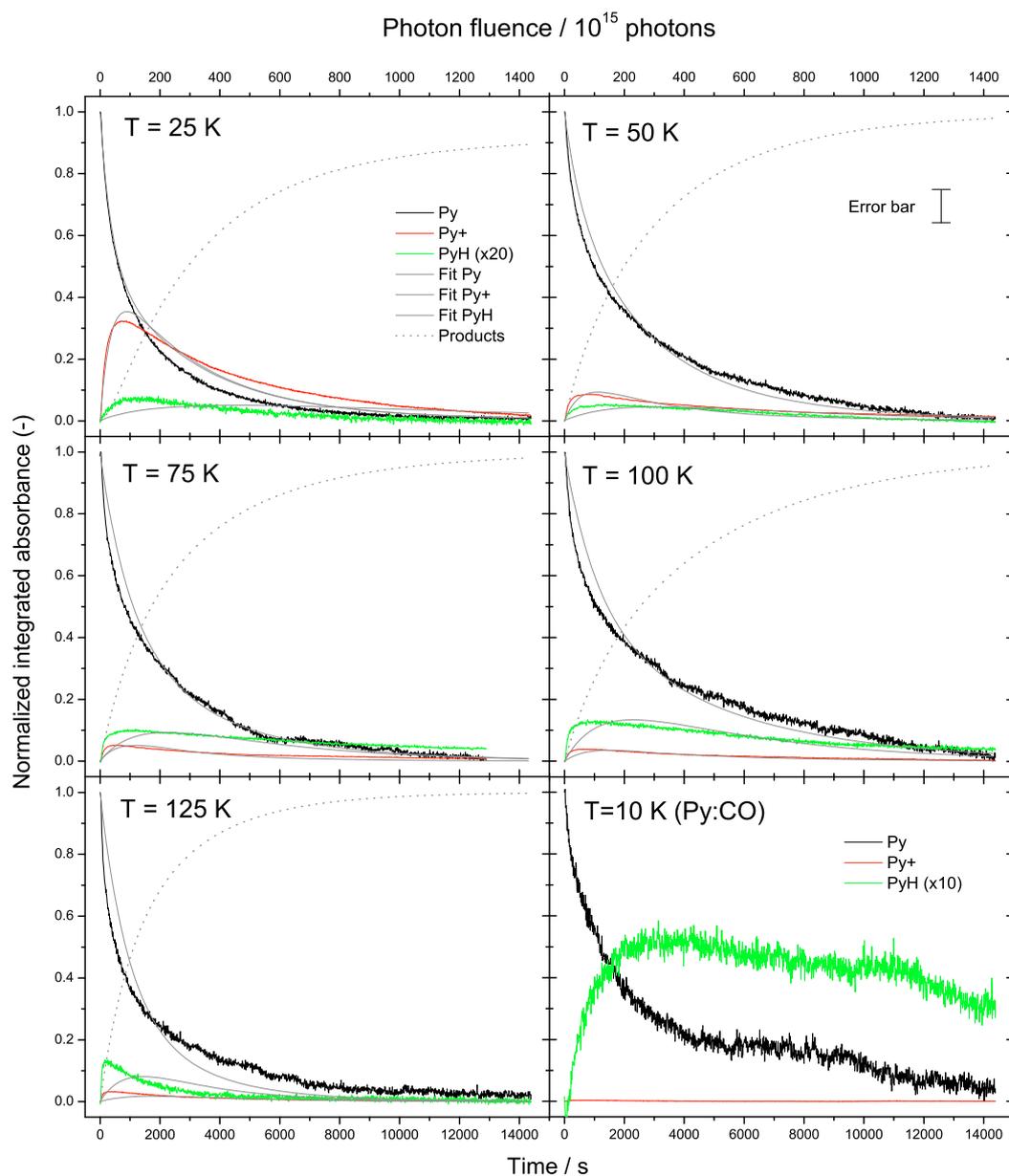


Fig. 6. The integrated absorbance of the Py 334 nm, Py^+ 445 nm, and PyH 400 nm bands as a function of VUV irradiation in $\text{Py}:\text{H}_2\text{O}$ ices at 25, 50, 75, 100, and 125 K and a $\text{Py}:\text{CO}$ ice at 10 K plotted together with the fits (grey lines) described in Sect. 4. Integrated absorbance values are scaled and normalized to the initial value for the Py signal. The PyH feature is multiplied by a factor of 20 for the $\text{Py}:\text{H}_2\text{O}$ experiments and by a factor of 10 for the $\text{Py}:\text{CO}$ experiment. The approximate, overall growth of the total Py photoproduct band ($\text{P}_1+\text{P}_2+\text{P}_3$) is also shown (dotted line).

This efficiency remains of the same order at even lower temperatures (not shown here). The photolysis time required for the cation to reach a maximum shortens with increasing temperature. The PyH band contribution is minor with respect to the Py^+ band for ices below 50 K. This reverses between 50 and 75 K, suggesting that there is a change in the dominant $\text{Py}:\text{H}_2\text{O}$ ice photochemical channel in this temperature range.

A kinetic analysis of the plots in Fig. 6 is carried out using the reaction scheme indicated in Fig. 7. Here, k_{11} is the photoionization rate of Py to Py^+ , k_{12} the electron-ion recombination rate of Py^+ , k_{21} the production rate of the PyH feature, and k_{22} the rate of the reverse reaction of PyH to Py. The rates designated k_1 , k_2 , and k_3 are the production rates for the different products that comprise the Py-residue band. The oscillator strengths for the Py^+ and PyH bands are also fitted, but are restricted to

remain within $\pm 10\%$ of the experimentally determined values of 0.33 and 0.089. All reactions are assumed to be first order in the reactant. The relative abundances of “free or solvated electrons” and O, H, and OH radicals in the ice are not considered.

The fits to the growth and decay curves are included in Fig. 6 and the temperature dependence of the derived rate constants is presented in Fig. 8. The agreement between the fit and the experimental data in terms of curve shape and absolute intensity is good. The fitted oscillator strengths of the Py^+ and PyH bands amount to 0.31 and 0.082, respectively, and hence do not deviate much from the experimentally determined values.

The graph in Fig. 8 indicates that the Py photoionization rate (k_{11}) drops rapidly between 25 and 50 K. The electron recombination rate (k_{12}) decreases only slightly, if at all, within the errors over the entire temperature range. As mentioned above,

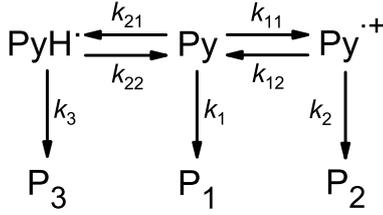


Fig. 7. Reaction scheme used to fit the experimental data.

the production of the PyH^{\cdot} becomes more important at higher temperatures. Its formation rate (k_{21}) is low in all ices up to 50 K ($<4.4 \times 10^{-5}$), but jumps to $>1 \times 10^{-4}$ in the ices with temperatures of 75 K and higher. The back channel from PyH^{\cdot} to Py , k_{22} , also shows a temperature dependence. It increases almost linearly in going from cold to warm ices. The formation rate of a photoproduct produced directly from Py (k_1) also seems to jump at 50 K. The formation rate of products originating in the $\text{Py}^{\cdot+}$ species, on the other hand, seems to lower with increasing temperature. Finally, the rate of product formation from the PyH^{\cdot} channel is low throughout the entire temperature range. The jump in rate of the formation of P_1 and PyH^{\cdot} with temperature probably reflects the diffusion barrier of radical species (H and OH^{\cdot}) in the ice.

Since published studies of the processes induced by the photolysis of other PAH: H_2O ices are limited, not much information is available with which to compare these results. While, to the best of our knowledge, there are no reports of the photochemistry that takes place as a function of ice temperature or of long-term fluence, the VUV photochemistry of the PAHs naphthalene, 4-methylpyrene (4MP), and quaterrylene in water ice at 10 K has been studied (Gudipati & Allamandola 2003; Gudipati 2004; Gudipati & Allamandola 2006a,b). The results obtained are in good agreement with the low temperature (25 K) case reported here. Namely, the parent PAH is easily and efficiently ionized, by quantitative conversion of the neutral species to the cation form. The focus of the earlier studies was on cation production and stabilization and not on long-duration photolysis experiments. In their study of 4MP: H_2O (1:>500) ice at 15 K, Gudipati & Allamandola (2003) utilized a reaction scheme similar to that on the right half of that presented in Fig. 7. Table 2 compares the reaction rates that they determined with those of the 25 K ice reported here. Except for the production of P_2 , which differs by one order of magnitude, there is very good agreement between the rate constants for each step in the two experiments.

The growth and decay curves in Fig. 6, taken together with the temperature dependence of the reaction rates in Fig. 8, show that the VUV-driven PAH photochemistry depends strongly on ice temperature. The influence of the ice morphology on this chemistry was also investigated, to understand the origin of the temperature dependence. An experiment on an ice deposited at 25 K, annealed to 125 K, and subsequently cooled to 25 K before photolysis, showed that the ionization rate and efficiency are similar to that of an unannealed ice. Apparently, it is not the morphology but the temperature of the ice that primarily determines which process dominates. We discriminate between two temperature regimes. One governed by ion-mediated processes that dominate at 25 K and slightly higher temperatures, and a second, presumably radical-driven regime, that becomes increasingly more important at higher temperatures.

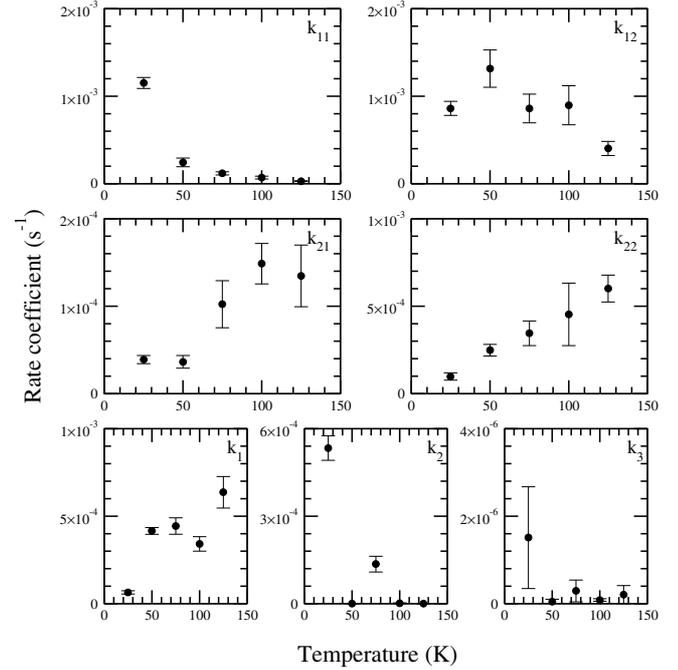


Fig. 8. Parameters (k_{nm}) as a function of temperature resulting from fitting the reaction scheme (Fig. 7) to the kinetic experiments (Fig. 6). All rates are indicated in s^{-1} .

Table 2. The reaction rates for the VUV photolysis of $\text{Py}:\text{H}_2\text{O}$ (~1:5000) ice at 25 K compared to those for 4-methylpyrene: H_2O (1:>500) ice at 15 K (Gudipati & Allamandola 2003).

Rate	This work (s^{-1})	Photon rate ^a ($\text{cm}^2 \text{photon}^{-1}$)	Gudipati (2003) (s^{-1})
$\text{Py} \xrightarrow{\text{VUV}} \text{Py}^{\cdot+}$ (k_{11})	$(1.2 \pm 0.1) \times 10^{-3}$	1.2×10^{-17}	1.3×10^{-3}
$\text{Py}^{\cdot+} + e^- \rightarrow \text{Py}$ (k_{12})	$(9 \pm 2) \times 10^{-4}$	9×10^{-18}	8×10^{-4}
$\text{Py} \rightarrow \text{P}_1$ (k_1)	$(5 \pm 1) \times 10^{-5}$...	4×10^{-5}
$\text{Py}^{\cdot+} \rightarrow \text{P}_2$ (k_2)	$(5 \pm 1) \times 10^{-4}$...	5×10^{-5}

Notes. ^(a) Photon rates are indicated only for reaction channels which are dominated by photon processes.

5. Astrochemical implications

As shown in the previous sections, ionization and chemistry of a rather small PAH species, pyrene, trapped in H_2O ice turns out to be very efficient in a laboratory setting. Here, we extend these findings to interstellar conditions, with the aim of including the calculated rates in astrochemical models. For this, it is crucial to distinguish pure photochemical processes from diffusion, since the latter will be highly dependent on the number density of radicals and electrons in the ice. As mentioned in the previous section, the photoionization of Py is probably a single-photon process, whereas protonation of Py and the electron recombination of $\text{Py}^{\cdot+}$ are the results of both VUV photolysis and diffusion. The mechanism for PyH^{\cdot} deprotonation is unclear, since it can proceed by means of either VUV processing or through hydrogen abstraction by diffusing species. Diffusion of radicals through the ice is a thermally activated process and will therefore increase with temperature. Recombination, however, is largely temperature-independent in our experiments, indicating that the rate of $\text{Py}^{\cdot+}$ recombination is not dominated by the diffusion of electrons in the ice. If $\text{Py}^{\cdot+}$ loss occurs by means of electron recombination and not $\text{Py}^{\cdot+}$ reaction with H_2O or one of its

photoproducts, the electron most likely originates from the initial photoionization event after which electrons remain in the vicinity of the recombining Py^+ species. Hence, this local process can be, although indirectly, regarded as a single-photon process.

The rates of protonation of Py and deprotonation of PyH^+ show a temperature dependence and the importance of diffusion can therefore not be excluded. This makes it harder to directly translate the rates (s^{-1}) into photon rates ($\text{cm}^2\text{photon}^{-1}$). However, we can determine astrochemical photon rates for both ionization and recombination of pyrene in interstellar H_2O ice (see Table 2).

Now, to translate this to the astrochemical situation and with other processes, we assume that PAHs generally have an ionization rate similar to that of pyrene. How do ionization and chemistry compare with other processes such as the photodesorption of the icy grain mantle, in which the PAHs are embedded? To exemplify this, the rate of ionization of a PAH in water ice at 25 K (in photon^{-1}) is calculated anywhere in a dense cloud where $A_V = 3$ and compared with the VUV photodesorption rate of H_2O derived by Öberg et al. (2009c). It is well established that the onset of ice formation occurs in clouds with an edge-to-edge (through the cloud) magnitude of $A_V = 3$ (e.g. Whittet et al. 2001). Thus, inside our hypothetical dense cloud at $A_V = 3$ (from cloud edge to within the cloud), ices are present.

The experimentally determined PAH ionization rate in H_2O at 25 K, normalized to the total amount of deposited PAH is given by

$$k_{11} = \frac{d[\text{PAH}^+]}{d[\text{PAH}]_0} = 10^{-3} \text{ s}^{-1}. \quad (9)$$

Consider a typical interstellar grain, covered by a 100 monolayer (ML) thick ice. The number of sites on a grain is 10^{15} cm^{-2} . If we assume that one in every 10^4 particles on the grain is a PAH, the total number of PAH molecules on the grain is $[\text{PAH}]_0 = 100 \times 10^{15} \times 10^{-4} = 10^{13} \text{ cm}^{-2}$. Furthermore, the VUV photon flux in our laboratory, Φ , is $10^{14} \text{ photons cm}^{-2} \text{ s}^{-1}$. The production rate of PAH cations on an interstellar grain is now given by $[\text{PAH}]_0 \cdot k_{11}/\Phi = 10^{-4} \text{ photon}^{-1}$. This ionization rate is an order of magnitude lower than the rate of photodesorption ($\sim 10^{-3} \text{ photon}^{-1}$) (Öberg et al. 2009c).

However, in our dense cloud the number of photons available for PAH photoionization is larger than the number of photons available for photodesorption of H_2O ice. This is because H_2O photodesorption is primarily caused by VUV photons, whereas PAH ionization can occur for much lower energy photons. To quantify the radiation field in a dense cloud at $A_V = 3$ as a function of wavelength (λ), we take the average UV interstellar radiation field (I_ν) from Sternberg (1988) and rewrite the expression to I_λ with units $\text{photons cm}^{-2} \text{ s}^{-1} \text{ nm}^{-1}$

$$I_\lambda = \frac{1.068 \times 10^{-4} c}{\lambda^3} - \frac{1.719 \times 10^{-2} c}{\lambda^4} + \frac{6.853 \times 10^{-1} c}{\lambda^5}, \quad (10)$$

where c is the speed of light in nm s^{-1} . The attenuation of the radiation field by dust as a function of wavelength is given by

$$D_\lambda = \exp\left(\frac{-11.6 A_V}{R_V} \frac{A_\lambda}{A_{1000 \text{ \AA}}}\right), \quad (11)$$

from Draine & Bertoldi (1996), where we assume that $R_V = 3.1$ and $A_V/A_{1000 \text{ \AA}} = 0.21$ (Whittet 2003). This results in

$$D_\lambda = \exp\left(-0.8 \frac{A_\lambda}{A_V} A_V\right), \quad (12)$$

where the table of A_λ/A_V values is taken from Mathis (1990). The photon flux per second per wavelength interval is given by

$$P_\lambda = I_\lambda D_\lambda. \quad (13)$$

Water ice absorbs photons with wavelengths ranging from 130 to 150 nm (Kobayashi 1983; Andersson & van Dishoeck 2008). The ionization energy of PAHs on the other hand, is lowered by about 2 eV when in H_2O ice (Gudipati & Allamandola 2004; Woon & Park 2004). For the wavelength range available for ionization of PAHs, assuming that H_2O blocks all photons below 150 nm, we take 150 to 250 nm (Li & Draine 2001). By integrating the photon flux in a cloud of $A_V = 3$ over both wavelength intervals a number of photons available for PAH ionization is found that is 6 times larger than the number of photons available for photodesorption of H_2O . Additionally, at $A_V = 3$, the cosmic-ray-induced UV field is negligible compared to the interstellar UV field (Shen et al. 2004). Therefore, the occurrence of photoionization is of similar order as photodesorption of the main component in the grain mantle in a dense cloud. The ionization rates from Table 2 can be directly included in astrochemical models in the form

$$\frac{d[\text{PAH}^+]}{dt} = k_{11} \Psi[\text{PAH}], \quad (14)$$

where $[\text{PAH}^+]$ is the concentration of the PAH (pyrene) cation in the ice, k_{11} is the photon rate in $\text{cm}^2 \text{ photon}^{-1}$, Ψ is the photon flux in $\text{photon s}^{-1} \text{ cm}^2$, and $[\text{PAH}]$ is the PAH (pyrene) concentration in the ice.

In the above calculation, we assume that all PAHs exhibit the ionization behavior of the pyrene chromophore. Of course, more PAHs need to be investigated experimentally before drawing conclusions on their general photochemical behavior in interstellar ices. However, if all PAHs have an ionization rate similar to that of pyrene, photoionization and subsequent chemical reactions of PAHs trapped in ices are important processes in dense clouds. When frozen-out in ices, PAHs have an important impact on the radical and electron budget in solid state chemistry. Hence, the processes described here may be more important than previously assumed in modeling complex interstellar grain chemistry.

6. Conclusions

A recently constructed setup has been used to track, on a sub-second timescale, the photochemistry of a PAH in H_2O and CO ices as a function of temperature. The setup used here clearly has advantages compared to relatively slow infrared photochemical ice studies. The conclusions from this work on a PAH, pyrene, trapped in an interstellar ice analogue are summarized below:

1. A set of photochemical reaction products has been identified in both irradiated $\text{Py}:\text{H}_2\text{O}$ and $\text{Py}:\text{CO}$ ice experiments. The reaction products result from direct photoionization of pyrene, or from a reaction of the parent, pyrene, with free H atoms produced in the matrix. Additionally, an absorption band is tentatively assigned to a triplet-triplet transition of pyrene. A vibrational progression assigned to HCO^+ is found in spectra of the VUV-irradiated $\text{Py}:\text{CO}$ ice.
2. Pyrene is easily and efficiently ionized when trapped in H_2O ice. Photoionization is a non-diffusion-related reaction and hence a photon rate of $1.2 \times 10^{-17} \text{ cm}^2 \text{ photon}^{-1}$, which can serve as input for astrochemical models, is derived.
3. When trapped in CO ice, pyrene ionization is inefficient compared to that in water ice.

4. Electron-ion recombination is independent of ice temperature and is characterized as a non-diffusion-dominated reaction. For this process, a photon rate of $9 \times 10^{-18} \text{ cm}^2 \text{ photon}^{-1}$ is derived, which can be directly used in astrochemical models.
5. There are two distinct reaction paths in the photochemistry of pyrene trapped in H₂O ice. At low temperatures (<50 K), the chemistry is dominated by ion-molecule interactions and processes. At temperatures above 50 K, reactions are dominated by diffusing radical species.
6. A simple model indicates that, in dense clouds where $A_V = 3$, the rate of pyrene ionization is comparable to the rate of photodesorption in water-rich ices. Hence, chemical reactions involving pyrene and its cation, and other PAHs in general, may be important and should be taken into account in modeling grain chemistry in these environments.

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